

7.7: Polyprotic Acids

The name "polyprotic" literally means many protons. Therefore, in this section we will be observing some specific acids and bases which either lose or accept *more than one* proton. Then, we will be talking about the equations used in finding the degree of dissociation. Finally, with given examples, we will be able to approach problems dealing with polyprotic acids and bases.

Introduction

Polyprotic acids are specific acids that are capable of *losing* more than a single proton per molecule in acid-base reactions. (In other words, acids that have more than one ionizable H^+ atom per molecule). Protons are lost through several stages (one at each stage), with the first proton being the fastest and most easily lost. Contrast with *monoprotic acids* in section [Monoprotic Versus Polyprotic Acids And Bases](#).

Common Polyprotic Acids	Formula	Strong/Weak Acid	Number of Ionizable Hydrogens	K_{a1}	K_{a2}	K_{a3}
Sulfuric acid	H_2SO_4	Strong	2 (diprotic)	Very Large	$1.1E-2$	
Sulfurous acid	H_2SO_3	Weak	2 (diprotic)	$1.3E-2$	$6.2E-8$	
Phosphoric acid	H_3PO_4	Weak	3 (triprotic)	$7.1E-3$	$6.3E-8$	$4.2E-13$
Carbonic acid	H_2CO_3	Weak	2 (diprotic)	$4.4E-7$	$4.7E-11$	
Hydrosulfuric acid or Hydrogen sulfide	H_2S	Weak	2 (diprotic)	$1.0E-7$	$1E-19$	
Oxalic acid	$H_2C_2O_4$	Weak	2 (diprotic)	$5.4E-2$	$5.3E-5$	
Malonic acid	$H_2C_3H_2O_4$	Medium Strong	2 (diprotic)	$1.5E-3$	$2.0E-6$	

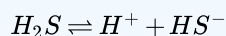
From the table above, we see that sulfuric acid is the strongest.

Ionization Constant

It is important to know that $K_1 > K_2 > K_3$, where K stands for the acidity constant or **acid ionization constant** (first, second, and third, respectively). These constants are used to measure the degree of dissociation of hydrogens in the acid. For a more in depth discussion on this, go to [Ionization Constants](#).

✓ Example 1: Hydrosulfuric acid

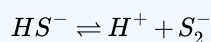
To find K_{a1} of Hydrosulfuric acid (H_2S), you must first write the reaction:



Dividing the products by the reactants, we then have:

$$K_{a1} = \frac{[H^+][HS^-]}{[H_2S]}$$

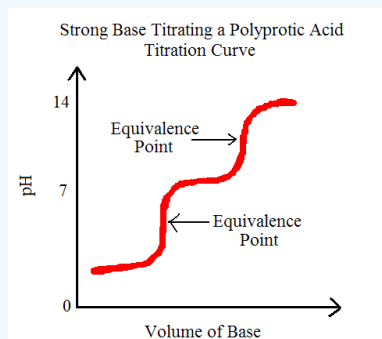
To find K_{a2} , we start with the reaction:



Then, like when finding K_{a1} , write the products over the reactants:

$$K_{a2} = \frac{[H^+][S^{2-}]}{[HS^-]}$$

From these reactions we can observe that it takes two steps to fully remove the H^+ ion. This also means that this reaction will produce two **equivalence points** or **stoichiometric points**. The equivalence point, by definition, is the point during an acid-base titration in which there has been equal amounts of acid and base reacted. If we were to graph this, we would be able to see exactly just what two equivalence points looks like. Let's check it out:



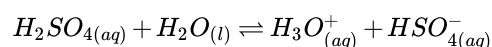
Note the multiple equivalence points and notice that they are almost straight lines at that point, indicating equal added quantities of acid and base.

Titration

In strong acid + strong base titrations, the pH changes slowly at first, rapidly through the equivalence point of $pH=7$, and then slows down again. If it is being titrated in a strong acid, the pH will go up as the base is added to it. Conversely, if it is in a strong base, the pH will fall down as acid is added.

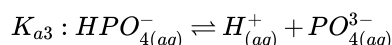
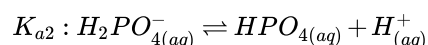
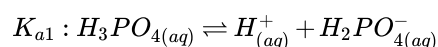
- In strong acid + weak base titrations, the pH changes slowly at the equivalence point and the pH equals the pK_a of the acid. The pH is below 7.
- For the weak acid + strong base, the pH is above 7 at the equivalence point.
- If there is strong acid or strong base left over after the equivalence point, this can be used to find the pH of the solution.

Next, let's take a look at sulfuric acid. This unique polyprotic acid is the only one to be completely deprotonated after the first step:



Now let's try something a little harder. The ionization of phosphoric acid (three dissociation reactions this time) can be written like this:

Start with H_3PO_4 :



So from *these* above reactions we can see that it takes three steps to fully remove the H^+ ion. This also means that this reaction will produce three equivalence points. **Polyprotic Bases** are bases that can *accept* at least one H^+ ion, or proton, in acid-base reactions.

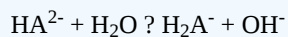
Common Polyprotic Bases	Formula	Strong/Weak Base	Diprotic/Triprotic Base
Phosphate ion	PO_4^{3-}	Weak	Triprotic
Sulfate ion	SO_4^{2-}	Very Weak	Diprotic
Carbonate ion	CO_3^{2-}	Strong	Diprotic

✓ Example 2: Some examples for calculating the constant, K_b

First, start with the reaction $A^{3-} + H_2O \rightleftharpoons HA^{2-} + OH^-$

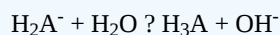
$$K_{b1} = [OH^-][HA^{2-}]/[A^{3-}] = K_w/K_{a3}$$

Then, we plug in the products over the reactants:



$$K_{b2} = [OH^-][H_2A^-]/[HA^{2-}] = K_w/K_{a2}$$

Finally, we are left with the third dissociation, or K_{b3} :



$$K_{b3} = [OH^-][H_3A]/[H_2A^-] = K_w/K_{a1}$$

References

1. Petrucci, et al. General Chemistry: Principles & Modern Applications: AIE (Hardcover). Upper Saddle River: Pearson/Prentice Hall, 2007.

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