

3.8: References

1. Hall NF, Conant JB (1927). "A Study of Superacid Solutions". *Journal of the American Chemical Society* **49** (12): 3062–70. doi:[10.1021/ja01411a010](https://doi.org/10.1021/ja01411a010).
2. George A. Olah, Schlosberg RH (1968). "Chemistry in Super Acids. I. Hydrogen Exchange and Polycondensation of Methane and Alkanes in FSO₃H–SbF₅ ("Magic Acid") Solution. Protonation of Alkanes and the Intermediacy of CH₅⁺ and Related Hydrocarbon Ions. The High Chemical Reactivity of "Paraffins" in Ionic Solution Reactions". *Journal of the American Chemical Society* **90** (10): 2726–7. doi:[10.1021/ja01012a066](https://doi.org/10.1021/ja01012a066).
3. Herlem, Michel (1977). "Are reactions in superacid media due to protons or to powerful oxidising species such as SO₃ or SbF₅?" *Pure & Applied Chemistry* **49**: 107–113. doi:[10.1351/pac197749010107](https://doi.org/10.1351/pac197749010107).
4. *Equilibrium acidities in dimethyl sulfoxide solution* Frederick G. Bordwell Acc. Chem. Res.; **1988**; 21(12) pp 456 - 463; DOI:[10.1021/ar00156a004](https://doi.org/10.1021/ar00156a004)
5. R. J. Cava, et al. (1988). "Superconductivity near 30 K without copper: the Ba_{0.6}K_{0.4}BiO₃ perovskite". *Nature* **332**: 814–6. doi:[10.1038/332814a0](https://doi.org/10.1038/332814a0).
6. Christe, Karl O. (1986). "Chemical synthesis of elemental fluorine". *Inorganic Chemistry* **25** (21): 3721. doi:[10.1021/ic00241a001](https://doi.org/10.1021/ic00241a001).
7. Cramer RE, Bopp TT (1977). "Great E and C plot. Graphical display of the enthalpies of adduct formation for Lewis acids and bases". *Journal of Chemical Education* **54** (10): 612–613. doi:[10.1021/ed054p612](https://doi.org/10.1021/ed054p612).
8. Stephan, D. W. (2008). "Frustrated Lewis pairs: a concept for new reactivity and catalysis". *Org. Biomol. Chem.* **6**: 1535–1539. doi:[10.1039/b802575b](https://doi.org/10.1039/b802575b).
9. Stephan, D. W.; Erker, G. (2010). "Frustrated Lewis Pairs: Metal-free Hydrogen Activation and More". *Angewandte Chemie International Edition* **49** (1): 46–76. doi:[10.1002/anie.200903708](https://doi.org/10.1002/anie.200903708). ISSN [1433-7851](https://doi.org/10.1002/anie.200903708).
10. Stephan, D. W.; Erker, G. (2017). "Frustrated Lewis pair chemistry". *Philosophical Transactions of the Royal Society A: Mathematical, Physical and Engineering Sciences* (Royal Society) **375** (2101): 20170239. doi:[10.1098/rsta.2017.0239](https://doi.org/10.1098/rsta.2017.0239). ISSN [1364-503X](https://doi.org/10.1098/rsta.2017.0239). PMID [28739971](https://pubmed.ncbi.nlm.nih.gov/28739971/). Bibcode: [2017RSPTA.37570239S](https://pubmed.ncbi.nlm.nih.gov/28739971/).
11. Welch, G. C.; San Juan, R. R.; Masuda, J. D.; Stephan, D. W. (2006). "Reversible, Metal-Free Hydrogen Activation". *Science* **314** (5802): 1124–1126. doi:[10.1126/science.1134230](https://doi.org/10.1126/science.1134230). ISSN [0036-8075](https://doi.org/10.1126/science.1134230). PMID [17110572](https://pubmed.ncbi.nlm.nih.gov/17110572/). Bibcode: [2006Sci...314.1124W](https://pubmed.ncbi.nlm.nih.gov/17110572/).
12. Birkmann, B., et al. (2010). "Frustrated Lewis Pairs and Ring-Opening of THF, Dioxane, and Thioxane". *Organometallics* **29**: 5310–5319. doi:[10.1021/om1003896](https://doi.org/10.1021/om1003896).

This page titled [3.8: References](#) is shared under a [CC BY-SA 4.0](#) license and was authored, remixed, and/or curated by [Chemistry 310 \(Wikibook\)](#) via [source content](#) that was edited to the style and standards of the LibreTexts platform.