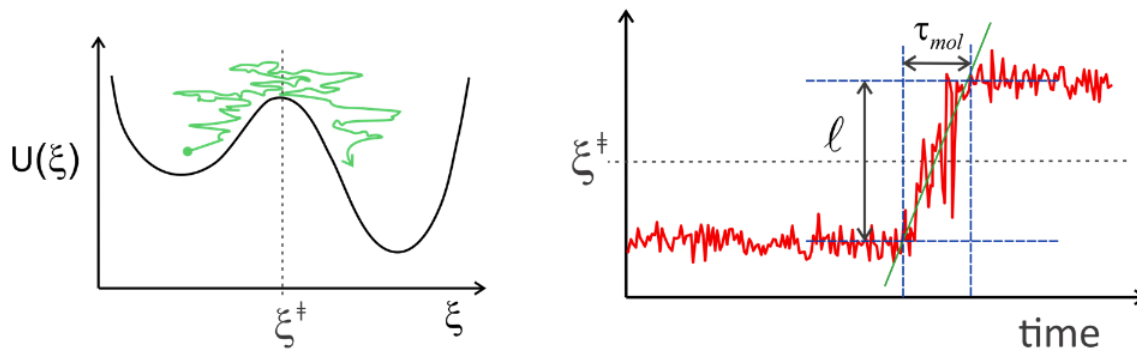


CHAPTER OVERVIEW

23: Barrier Crossing and Activated Processes

"Rare but important events"



The rates of chemical reaction are obtained by calculating the forward flux of reactant molecules passing over the transition state, i.e. the time rate of change of concentration, population, or probability for reactants passing over the transition state.

$$\langle J_f^\ddagger \rangle = dP_R^\ddagger / dt \quad (23.1)$$

‡

[23.1: Transition State Theory](#)

[23.2: Kramers' Theory](#)

This page titled [23: Barrier Crossing and Activated Processes](#) is shared under a [CC BY-NC-SA 4.0](#) license and was authored, remixed, and/or curated by [Andrei Tokmakoff](#) via [source content](#) that was edited to the style and standards of the LibreTexts platform.