

4.1: Aldehydes and Ketones

Learning Objectives

- Identify the general structure for an aldehyde and a ketone.
- Use common names to name aldehydes and ketones.
- Use the IUPAC system to name aldehydes and ketones.

The next group we consider, the **carbonyl group**, represents a carbon-oxygen double bond. The carbon of the carbonyl group also has two single bonds to other atoms/groups. One of the bonds is usually to another R group and the other bond distinguishes the functional group.



Figure 4.1.1: Structure of the carbonyl group.

The carbonyl group is ubiquitous in biological compounds. It is found in carbohydrates, fats, proteins, nucleic acids, hormones, and vitamins—organic compounds critical to living systems.

Carbonyl groups define two related families of organic compounds: the aldehydes and the ketones. In an **aldehyde**, at least one of the groups attached to the carbonyl carbon must be a hydrogen atom. The following general formulas, in which R represents an alkyl group and Ar stands for an aryl (aromatic) group, are aldehydes:

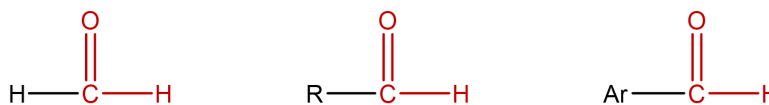


Figure 4.1.2: General structures of aldehydes.

In a **ketone**, two carbon groups are attached to the carbonyl carbon atom. The carbon groups can be identical or different. The following general formulas represent ketones.

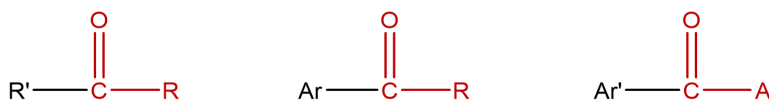


Figure 4.1.3: General structures of ketones.

In condensed formulas, we use CHO to identify an aldehyde rather than COH, which might be confused with an alcohol. This follows the general rule that in condensed structural formulas H comes after the atom it is attached to (usually C, N, or O). The carbon-oxygen double bond is not shown but understood to be present. Because they contain the same functional group, aldehydes and ketones share many common properties, but they still differ enough to warrant their classification into two families.



Figure 4.1.4: Condensed structures of aldehydes and ketones.

Naming Aldehydes and Ketones

Both common and International Union of Pure and Applied Chemistry (IUPAC) names are frequently used for aldehydes and ketones, with common names predominating for the lower homologs. The common names of aldehydes are taken from the names of the acids into which the aldehydes can be converted by *oxidation*.

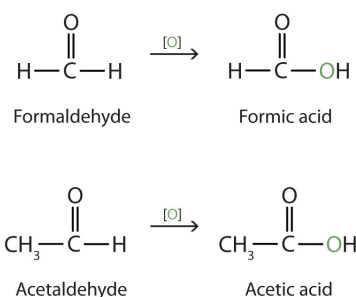
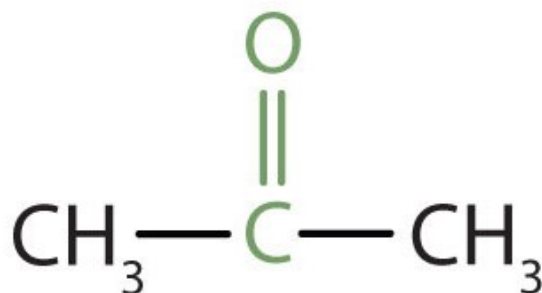


Figure 4.1.5: Aldehydes with common names and the organic acid they can form.

The stems for the common names of the first four aldehydes are as follows:

- 1 carbon atom: *form-*
- 2 carbon atoms: *acet-*
- 3 carbon atoms: *propion-*
- 4 carbon atoms: *butyr-*

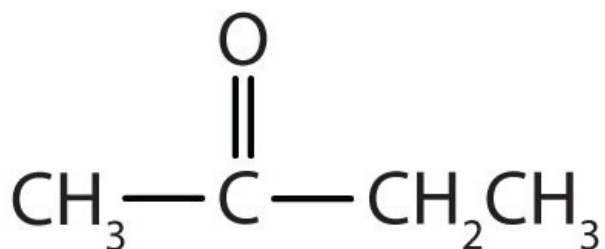
Because the carbonyl group in a ketone must be attached to two carbon groups, the simplest ketone has three carbon atoms. It is widely known as **acetone**, a unique name unrelated to other common names for ketones.



Acetone

Figure 4.1.6: Structure of acetone.

Generally, the common names of ketones consist of the names of the groups attached to the carbonyl group, followed by the word *ketone*. (Note the similarity to the naming of ethers.) Another name for acetone, then, is *dimethyl ketone*. The ketone with four carbon atoms is ethyl methyl ketone (Figure 4.1.7).

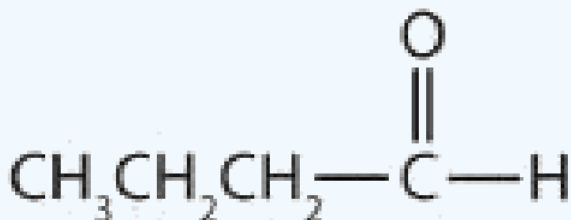


Ethyl methyl ketone

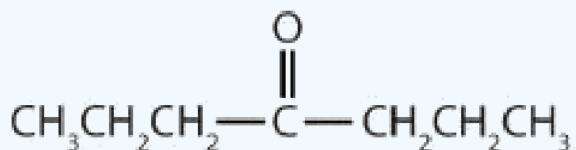
Figure 4.1.7: Structure of ethyl methyl ketone.

✓ Example 4.1.1

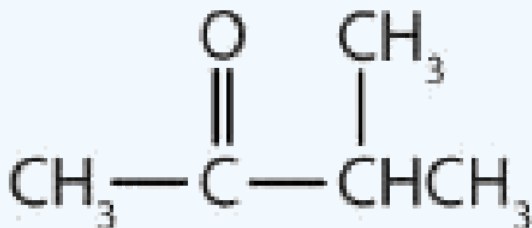
Classify each compound as an aldehyde or a ketone. Give the common name for each ketone.



a.



b.



c.

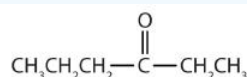
Solution

1. This compound has the carbonyl group on an end carbon atom, so it is an aldehyde.
2. This compound has the carbonyl group on an interior carbon atom, so it is a ketone. Both alkyl groups are propyl groups. The name is therefore dipropyl ketone.
3. This compound has the carbonyl group between two alkyl groups, so it is a ketone. One alkyl group has three carbon atoms and is attached by the middle carbon atom; it is an isopropyl group. A group with one carbon atom is a methyl group. The name is therefore isopropyl methyl ketone.

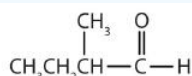
? Exercise 4.1.1

Classify each compound as an aldehyde or a ketone. Give the common name for each ketone.

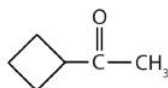
1.



2.



3.



Here are some simple IUPAC rules for naming aldehydes and ketones:

- The stem names of aldehydes and ketones are derived from those of the parent alkanes, defined by the **longest continuous chain (LCC) of carbon atoms that contains the functional group**.
 - For an **aldehyde**, drop the *-e* from the alkane name and add the ending **-al**. Methanal is the IUPAC name for formaldehyde, and ethanal is the name for acetaldehyde.
 - For a **ketone**, drop the *-e* from the alkane name and add the ending **-one**. Propanone is the IUPAC name for acetone, and butanone is the name for ethyl methyl ketone.
- Number the parent chain** to determine the location of atoms present in the molecule.
 - To indicate the position of a substituent on an aldehyde, the carbonyl carbon atom is always considered to be C1; it is unnecessary to designate this group by number.
 - To indicate the position of a substituent on a ketone, number the chain in the manner that gives the carbonyl carbon atom the lowest possible number. In cyclic ketones, it is understood that the carbonyl carbon atom is C1.
- Substituents are named and numbered** as in alkanes. These groups are listed in alphabetical order prior to the parent name. According to these rules, the IUPAC name for the molecule shown below in Figure 4.1.8 is 5-methylhexanal.

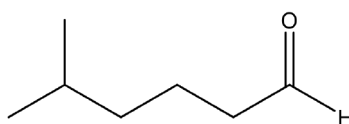
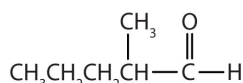


Figure 4.1.8: Skeletal structure of an aldehyde named using IUPAC rules.

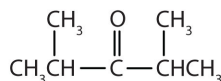
✓ Example 4.1.2

Give the IUPAC name for each compound.

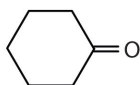
a.



b.



c.

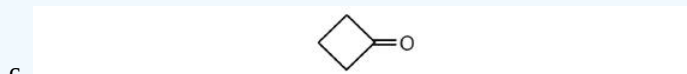
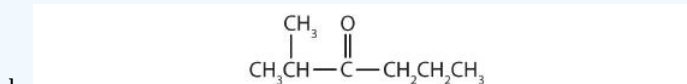
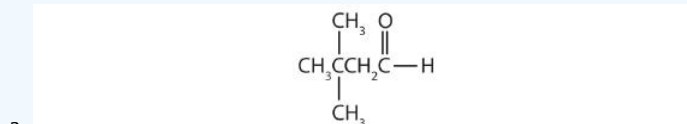


Solution

- There are five carbon atoms in the LCC. The methyl group (CH_3) is a substituent on the second carbon atom of the chain; the aldehyde carbon atom is always C1. The name is derived from pentane. Dropping the *-e* and adding the ending *-al* gives pentanal. The methyl group on the second carbon atom makes the name 2-methylpentanal.
- There are five carbon atoms in the LCC. The carbonyl carbon atom is C3, and there are methyl groups on C2 and C4. The IUPAC name is 2,4-dimethyl-3-pentanone.
- There are six carbon atoms in the ring. The compound is cyclohexanone. No number is needed to indicate the position of the carbonyl group because all six carbon atoms are equivalent.

? Exercise 4.1.2

Give the IUPAC name for each compound.



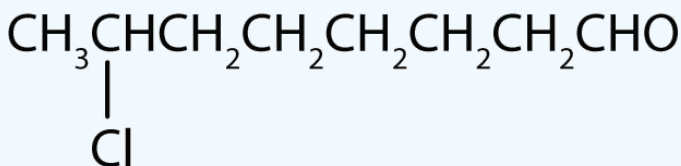
✓ Example 4.1.3

Draw the structure for each compound.

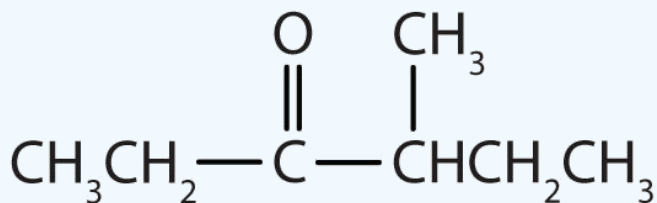
- 7-chlorooctanal
- 4-methyl-3-hexanone

Solution

- The *octan-* part of the name tells us that the LCC has eight carbon atoms. There is a chlorine (Cl) atom on the seventh carbon atom; numbering from the carbonyl group and counting the carbonyl carbon atom as C1, we place the Cl atom on the seventh carbon atom.



- The *hexan-* part of the name tells us that the LCC has six carbon atoms. The 3 means that the carbonyl carbon atom is C3 in this chain, and the 4 tells us that there is a methyl (CH_3) group at C4:



? Exercise 4.1.3

Draw the structure for each compound.

- 5-bromo-3-iodoheptanal
- 5-bromo-4-ethyl-2-heptanone

Summary

- IUPAC names of aldehydes and ketones are derived from those of the parent alkanes, using an *-al* ending for an aldehydes and an *-one* ending for a ketone.
 - aldehydes: location and identity of substituents + parent stem name + al suffix
 - ketones: location and identity of substituents + parent stem name (with location of functional group) + one suffix
- The common names of aldehydes are taken from the names of the corresponding carboxylic acids: formaldehyde, acetaldehyde, and so on. Common name of aldehydes: name based on number of carbon in alkyl group + aldehyde.
 - 1 carbon atom: *form-*
 - 2 carbon atoms: *acet-*
 - 3 carbon atoms: *propion-*
 - 4 carbon atoms: *butyr-*
- The common names of ketones, like those of ethers, consist of the names of the groups attached to the carbonyl group, followed by the word *ketone*. Common name of ketones: identities of alkyl groups (in alphabetical order) + aldehyde.

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