

Chapter 9.5 Assigning 4 Quantum Numbers to Electrons in Subshells (Video)

This project was preformed to supply **Libretext Authors** with videos on General Chemistry topics which can be used to enhance their projects. Also, these videos are meant to act as a learning resource for **all General Chemistry students**.

Video Topics

When assigning quantum number to electrons in a given subshell the value of n and l are given in the orbital designation. When assigning values for m_l and m_s it is important to remember their values are also affected by other rules. Pauli Exclusion Principle: When two electrons occupy the same orbital (m_l) their spins must be paired. $+1/2$ and $-1/2$ Hund's rule: Fill each orbital (m_l) in a subshell with one electron first before putting two electrons in the same orbital. (Keep spins parallel) This video contains multiple examples of assigned quantum number to electrons in a given orbital designation.

Link to Video

Assigning 4 Quantum Numbers to Electrons in Subshells: <https://youtu.be/4InLA131aE0>



Attribution

- Prof. Steven Farmer ([Sonoma State University](#))

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