

## 17.10 Converting $K_c$ to $K_p$ (Video)

This project was preformed to supply **Libretext authors** with videos on General Chemistry topics which can be used to enhance their projects. Also, these videos are meant to act as a learning resource for **all General Chemistry students**.

### Video Topics

Equilibrium constants can be expressed as a partial pressure of reactants and products ( $K_p$ ). Concentration and partial pressure can be related using the ideal gas law  $PV = nRT$   $KP = KC(RT)^n$  gas  $R = 0.08205 \text{ L atm/mol K}$   $T = \text{Temp in K}$   $n_{\text{gas}} = \text{Stoic. Coefficients of the gaseous products} - \text{Stoic. Coefficients of the gaseous reactant}$ .

### Link to Video

Converting  $K_c$  to  $K_p$ : [https://youtu.be/\\_2WVnlqXrV4](https://youtu.be/_2WVnlqXrV4)



### Attribution

- Prof. Steven Farmer ([Sonoma State University](#))

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