

17.3.2 Le Châtelier's Principle (Changes in Pressure or Volume) (Video)

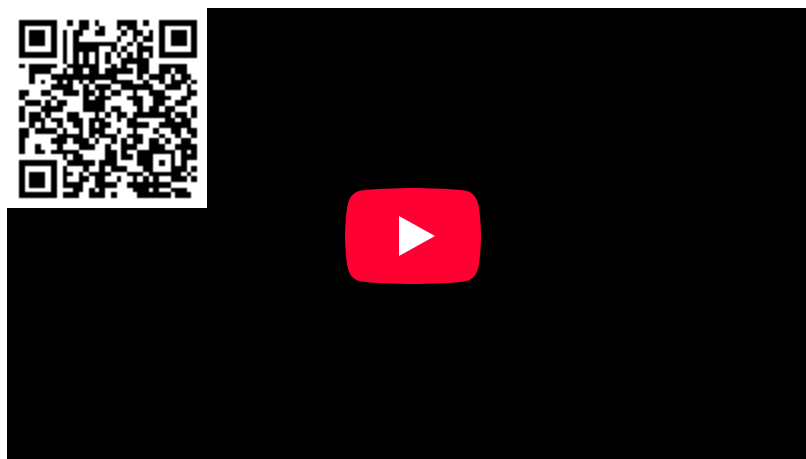
This project was preformed to supply **Libretext authors** with videos on General Chemistry topics which can be used to enhance their projects. Also, these videos are meant to act as a learning resource for **all General Chemistry students**.

Video Topics

When the volume is reduced or the pressure is increased, the reaction shifts in the direction that produces the fewest mols of gas. If the volume is increased or the pressure is decreased, the reaction shifts in the direction that produces the most mols of gas. $\text{N}_{2(g)} + 3 \text{H}_{2(g)} \rightleftharpoons 2 \text{NH}_{3(g)}$ Which way will the reaction shift to reach equilibrium if the volume is decreased or pressure is increased? The side with the least mols of gas is favored. 4 VS 2 so the reaction goes right Which way will the reaction shift to reach equilibrium if the volume is increased or the pressure is decreased? 4 VS 2 so the reaction goes left

Link to Video

Le Châtelier's Principle (Changes in Pressure or Volume): <https://youtu.be/fHIFTm2gxU>



Attribution

- Prof. Steven Farmer ([Sonoma State University](#))

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