

19.8.4 pH at the Equivalence Point of a Weak Acid/Strong Base Titration (Video)

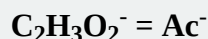
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Video Topics

This video contains an example problem where pH at the equivalence of a weak acid / strong base titration is calculated.

If we have 50. mL of a 0.100 M HC₂H₃O₂ solution:

Calculate the pH after the addition of 50 mL of a 0.100 NaOH solution:



$$K_a = 1.80 \times 10^{-5}$$

$$\text{p}K_a = 4.74$$

Link to Video

pH at the Equivalence Point of a Weak Acid/Strong Base Titration: <https://youtu.be/RuSr1z6F0To>



Attribution

- Prof. Steven Farmer (Sonoma State University)

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