17.5 Using the Reaction Quotient (Q) (Video)

This project was preformed to supply **Libretext authors** with videos on General Chemistry topics which can be used to enhance their projects. Also, these videos are meant to act as a learning resource for **all General Chemistry students**.

Video Topics

The reaction quotient Q determines the direction a reaction must go to get to equilibrium. $N_{2(g)} + O_{2(g)} - 2\ NO_{(g)}$ For the above reaction if we start with only N_2 and O_2 the reaction must go the right of make NO. If we start with only NO the reaction must go to the left to make N_2 and O_2. To find the direction of change first find the reaction quotient Q. Find Q by plugging the initial reaction concentrations into the equilibrium expression. If Q less than K the reaction goes to the right If Q more than K the reaction goes to the left If Q = K the system is at equilibrium

Link to Video

Using the Reaction Quotient (Q): https://youtu.be/_J04fgRs7QU

Attribution

- Prof. Steven Farmer (Sonoma State University)