

14.5 Phase Diagrams (Video)

This project was preformed to supply **Libretext Authors** with videos on General Chemistry topics which can be used to enhance their projects. Also, these videos are meant to act as a learning resource for **all General Chemistry students**.

Video Topics

Major Phases = Solid, Liquid and Gas. A phase diagram is a graphical representation of the different phases of a substance which exist at different temperatures and pressures. Temperature is on the x axis and pressure is on the y axis of the graph. The different regions of the diagram correspond to the different phases. Low pressure and High temperature = Gas High pressure and Low temperature = Solid High pressure and High temperature = Liquid The lines in a phase diagram correspond to the different phase changes. These are called the Sublimation, Fusion, and Vaporization lines. Supercritical Fluid: An ambiguous phase with properties of liquids and gases. The critical point is at the end of the vaporization line. Triple point: The convergence of the sublimation, fusion, and vaporization lines. All three phases exist at the same time under at a specific triple point pressure (P-TP) and triple point temperature (TTP). Normal Boiling Point: The temperature at which the Liquid/Gas phase transition occurs at 1 atm. Normal Freezing Point: The temperature at which the Solid/Liquid phase transition occurs at 1 atm. Normal Sublimation Point: The temperature at which the Solid/Gas phase transition occurs at 1 atm.

Link to Video

Phase Diagrams: <https://youtu.be/op1v7PaMsmU>



Attribution

- Prof. Steven Farmer ([Sonoma State University](#))

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