

19.8.6 Summary of the pH Curve of a Weak Acid/Strong Base Titration (Video)

This project was preformed to supply **Libretext authors** with videos on General Chemistry topics which can be used to enhance their projects. Also, these videos are meant to act as a learning resource for **all General Chemistry students**.

Video Topics

The video contains a summary of the important points of Weak Acid / Strong Base Titration Curve

Remember of the mol of strong base added = 0 then the pH is determined by the K_a equilibrium of the weak acid.

If the mol of strong base added < mol of weak acid initial then convert to mols and used the Henderson Hasselbalch equation.

When mol of strong base = mol weak acid you are at the equivalence point. The pH is determined by the concentration of the conjugate base.

If the mol of strong base added > mol of weak acid initial then you are past the equivalence point. The pH is determined by the concentration of the remaining ^-OH .

Link to Video

Summary of the pH Curve of a Weak Acid/Strong Base Titration: <https://youtu.be/hY8X727AiG0>



Attribution

- Prof. Steven Farmer ([Sonoma State University](#))

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