

19.8.3 pH at the Halfway Point of a Weak Acid/Strong Base Titration (Video)

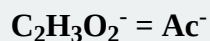
This project was preformed to supply **Libretext authors** with videos on General Chemistry topics which can be used to enhance their projects. Also, these videos are meant to act as a learning resource for **all General Chemistry students**.

Video Topics

This video contains an example problem where pH at the $\frac{1}{2}$ way point of a weak acid / strong base titration is calculated.

If we have 50. mL of a 0.100 M $\text{HC}_2\text{H}_3\text{O}_2$ solution:

Calculate the pH after the addition of 25 mL of a 0.100 NaOH solution:

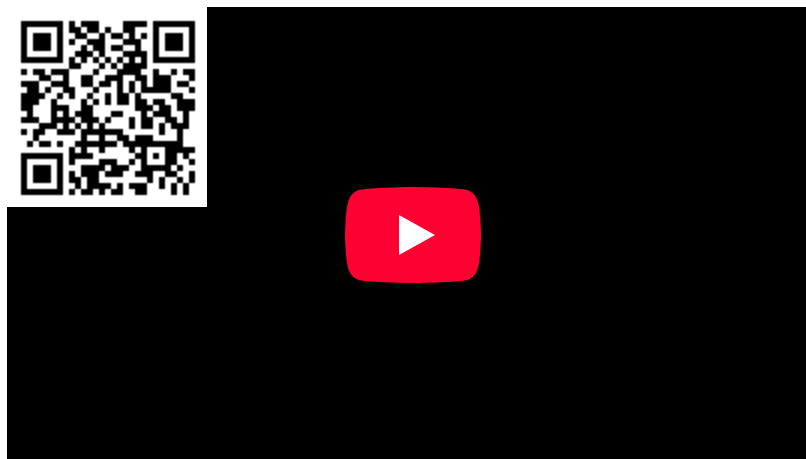


$$K_a = 1.80 \times 10^{-5}$$

$$\text{p}K_a = 4.74$$

Link to Video

pH at the Halfway Point of a Weak Acid/Strong Base Titration: <https://youtu.be/EYCj9TYMB4I>



Attribution

- Prof. Steven Farmer ([Sonoma State University](#))

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