

16.2 Average Reaction Rates (Video)

This project was preformed to supply **Libretext authors** with videos on General Chemistry topics which can be used to enhance their projects. Also, these videos are meant to act as a learning resource for **all General Chemistry students**.

Video Topics

Because the rate changes as the reaction proceeds what we are actually calculating is the average rate for the given time period.

Rate of appearance of product = $\frac{[P]_t2 - [P]_t1}{t2 - t1}$

Rate of disappearance of reactant = $\frac{[R]_t2 - [R]_t1}{t2 - t1}$

The rate of reaction is independent of stoichiometry.

Average rates cannot be used to find concentrations at a different time.

Average rates can be used to find information about other species in the reaction at the given time. Stoichiometry from the reaction is key.

Link to Video

Average Reaction Rates: <https://youtu.be/jc6jntB7GHk>



Attribution

- Prof. Steven Farmer ([Sonoma State University](#))

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