

16.12 Transition State Theory (Video)

This project was preformed to supply **Libretext authors** with videos on General Chemistry topics which can be used to enhance their projects. Also, these videos are meant to act as a learning resource for **all General Chemistry students**.

Video Topics

Why do some reactions proceed faster than others? Between the reactants and the products lies a transition state. The species in the transition state is called an activated complex. The activated complex is a mixture of the reactant and the product. The activated complex is unstable and requires the input of energy to create. This energy is called the energy of activation (E_a) and typically has the units of kJ/mol. E_a represents an energy barrier that must be overcome. The relative ease of formation of the activated complex determines the rate law constant (k) for a reaction.

Link to Video

Transition State Theory: <https://youtu.be/xmWDhq89b38>



Attribution

- Prof. Steven Farmer ([Sonoma State University](#))

16.12 Transition State Theory (Video) is shared under a [not declared](#) license and was authored, remixed, and/or curated by LibreTexts.