

14.3.2 Vapor Pressure and Boiling Point (Video)

This project was preformed to supply **Libretext Authors** with videos on General Chemistry topics which can be used to enhance their projects. Also, these videos are meant to act as a learning resource for **all General Chemistry students**.

Video Topics

Vapor Pressure: The pressure exerted by a vapor in dynamic equilibrium with its liquid. At room temperature some of the molecules of a liquid will have enough energy to enter a gaseous state. This represents the process of evaporation. If the liquid is not in a container, eventually all of liquid molecules will enter a gaseous state. If the liquid is in a container, eventually a point will be reached where the gaseous molecules start to condense back to the liquid state. High vapor pressure at room temperature: Volatile Low vapor pressure at room temperature: Nonvolatile As the intermolecular forces of a liquid increase, the vapor pressure tends to decrease. IMF \uparrow Vapor Pressure \downarrow Boiling point: When the vapor pressure greater than the applied pressure. As the vapor pressure of a liquid decreases the boiling point tends to increase. Vapor Pressure \downarrow Boiling Point \uparrow As the intermolecular forces of a liquid increases the boiling point tends to increase. IMF \uparrow Boiling Point \uparrow

Link to Video

Vapor Pressure & Boiling Point: <https://youtu.be/4QtcdpfRO1M>



Attribution

- Prof. Steven Farmer ([Sonoma State University](#))

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