

10.3 Electron Configuration of Transition Metals (Video)

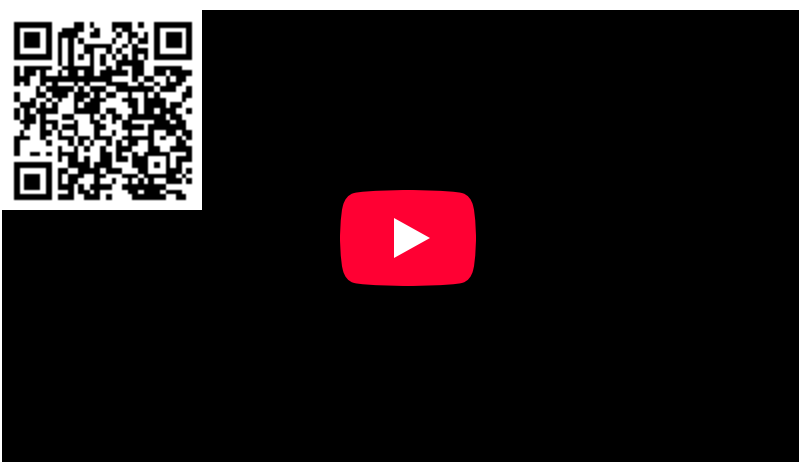
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Video Topics

Because of the relative energy levels atoms with $3d^4$ and $3d^9$ will steal an s electron to become $3d^5$ and $3d^{10}$ respectively. This is to gain a $\frac{1}{2}$ filled or whole filled d subshell. This means that the electron configuration of Cr is $[Ar]4s^13d^5$ and not $[Ar]4s^23d^4$. Also, the electron configuration of Cu is $[Ar]4s^13d^{10}$ and not $[Ar]4s^23d^9$.

Link to Video

Electron Configuration of Transition Metals: https://youtu.be/HzpfE0fk_E0



Attribution

- Prof. Steven Farmer ([Sonoma State University](#))

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