

21.8 Electroplating (Video)

This project was preformed to supply **Libretext Authors** with videos on General Chemistry topics which can be used to enhance their projects. Also, these videos are meant to act as a learning resource for **all General Chemistry students**.

Video Topics

Electroplating is the process of adding voltage to an electrochemical cell to cause a solid metal to form. Like in the half reaction: $\text{Ag}^+_{(\text{aq})} + \text{e}^- \rightarrow \text{Ag}_{(\text{s})}$

In order find the amount of solid metal deposited, we must be able to measure the number of electrons added.

Using the equation: Current = amperes = charge / time = Coulombs / sec

Coulombs are related to the mols of e^- added.

F = Faraday constant = 96,485 C/mol e^- this is the charge on a mol of electrons

If a known current is applied to a half reaction for a given amount of time, the number of grams of metal deposited can be calculated.

Link to Video

Electroplating: <https://youtu.be/yrNuWNleYAg>



Attribution

- Prof. Steven Farmer ([Sonoma State University](#))

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