

14.2.2 Dipole Intermolecular Force (Video)

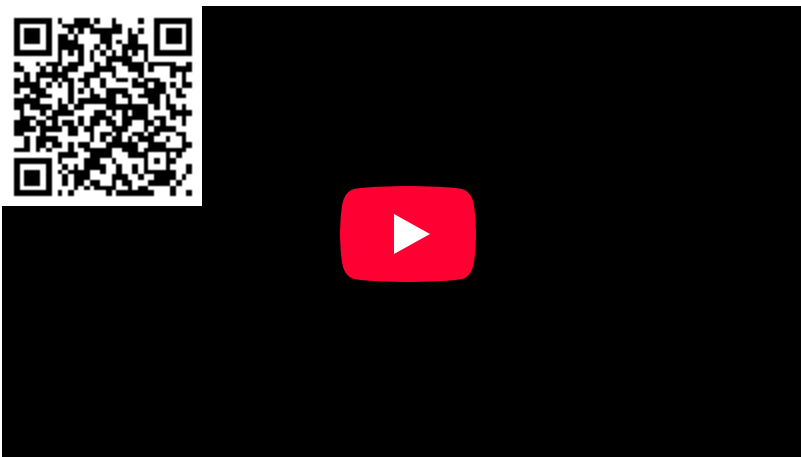
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Video Topics

Bond dipole moment. Electronegativity (EN) is a measure of an atoms ability to hold the electrons in a bond. EN increases as we go up and to the right other periodic table. The electron pair in a polar covalent bond is not shared equally. $\Delta EN \neq 0$. If A is more EN than B then $A^{\delta-}-B^{\delta+}$. A molecule's dipole moment is approximately the vector sum of its bond dipole moments. A nonpolar molecule can have polar bonds oriented so that the vector sum of the bond dipole moments is 0. Polar molecules have a stronger intermolecular force than non-polar molecules because of the charge separation.

Link to Video

Dipole Intermolecular Force: https://youtu.be/ACq_95SIBck



Attribution

- Prof. Steven Farmer ([Sonoma State University](#))

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