

## 7.9 Calculating $\Delta H^\circ$ using $\Delta H_f^\circ$ (Video)

This project was preformed to supply **Libretext Authors** with videos on General Chemistry topics which can be used to enhance their projects. Also, these videos are meant to act as a learning resource for **all General Chemistry students**.

### Video Topics

The enthalpy of an unknown reaction can be found using the equation:  $\Delta H^\circ = \text{Sum of } V\Delta H_f^\circ(\text{product}) - \text{Sum of } V\Delta H_f^\circ(\text{reactants})$ .  $\Delta H_f$  is the enthalpy of formation for a given species.  $V$  is the stoichiometric coefficient of the species from the balanced reaction. Remember that  $\Delta H_f^\circ = 0$  for elements in their reference states. This video contains a sample problem, which involves these concepts.

### Link to Video

Calculating  $H^\circ$  using  $H_f^\circ$ : <https://youtu.be/Y3aJJno9W2c>



### Attribution

- Prof. Steven Farmer ([Sonoma State University](#))

7.9 Calculating  $\Delta H^\circ$  using  $\Delta H_f^\circ$  (Video) is shared under a [not declared](#) license and was authored, remixed, and/or curated by LibreTexts.