

19.8.5 pH After the Equivalence Point of a Weak Acid/Strong Base Titration (Video)

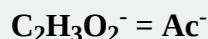
This project was preformed to supply **Libretext authors** with videos on General Chemistry topics which can be used to enhance their projects. Also, these videos are meant to act as a learning resource for **all General Chemistry students**.

Video Topics

This video contains an example problem where pH after the equivalence of a weak acid / strong base titration is calculated.

If we have 50. mL of a 0.100 M $\text{HC}_2\text{H}_3\text{O}_2$ solution:

Calculate the pH after the addition of 60 mL of a 0.100 NaOH solution:



$$K_a = 1.80 \times 10^{-5}$$

$$\text{p}K_a = 4.74$$

Link to Video

pH After the Equivalence Point of a Weak Acid/Strong Base Titration: <https://youtu.be/KHXPJIsxoLE>



Attribution

- Prof. Steven Farmer (Sonoma State University)

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