

15.11.1 Finding the Vapor Pressure of a Solution (Nonionic-Nonvolatile Solute) (Video)

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Video Topics

A nonionic-nonvolatile solute will cause the vapor pressure of the solvent to decrease. Because the solute is nonvolatile it will not create a vapor pressure of its own. Examples are: sugar, caffeine, fats, and proteins.

Raoult's Law $P_a = X_a P_{ao}$

P_a = Vapor pressure of solution

X_a = Mol fraction of the solvent

P_{ao} = Vapor pressure of the pure solvent

Raoult's Law says that the vapor pressure of a solvent tends to decrease when it is part of a solution. P_a less than P_{ao} .

Link to Video

Finding the Vapor Pressure of a Solution (Nonionic-Nonvolatile Solute): <https://youtu.be/WLceQuRlsPU>



Attribution

- Prof. Steven Farmer ([Sonoma State University](#))

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