

## 9.1 Introduction to Quantum Numbers (Video)

This project was preformed to supply **Libretext Authors** with videos on General Chemistry topics which can be used to enhance their projects. Also, these videos are meant to act as a learning resource for **all General Chemistry students**.

### Video Topics

In 1927 Schrödinger put forth the idea that the wavelike properties of an electron could be described by mathematical equations called wave functions. These wave functions were used to create orbitals which are probability distribution map showing where the electron is likely to be found. Quantum numbers are a way of describing the orbitals and electrons contained in an atom. There are 4 unique quantum numbers for every electron in an atom. 1) Principal quantum number ( $n$ ) which defines the electrons distance from the nucleus. 2) Orbital angular momentum ( $l$ ) which defines the type of orbital subshell 3)  $m_l$  is the magnetic quantum number which defines a specific orbital 4) The fourth quantum number ( $m_s$ ) refers to the electron spin of each electron.

### Link to Video

**Introduction to Quantum Numbers:** <https://youtu.be/07JpBeaPxL8>



### Attribution

- Prof. Steven Farmer ([Sonoma State University](#))

9.1 Introduction to Quantum Numbers (Video) is shared under a [not declared](#) license and was authored, remixed, and/or curated by LibreTexts.