19.8.5 pH After the Equivalence Point of a Weak Acid/Strong Base Titration (Video)

This project was preformed to supply **Libretext authors** with videos on General Chemistry topics which can be used to enhance their projects. Also, these videos are meant to act as a learning resource for **all General Chemistry students**.

Video Topics

This video contains an example problem where pH after the equivalence of a weak acid / strong base titration is calculated.

If we have 50. mL of a 0.100 M $HC_2H_3O_2$ solution:

Calculate the pH after the addition of 60 mL of a 0.100 NaOH solution:

$HC_2H_3O_2$ = HAc

$C_2H_3O_2^-$ = Ac^-

Ka = 1.80 x 10^{-5}

pKa = 4.74

Link to Video

pH After the Equivalence Point of a Weak Acid/Strong Base Titration: https://youtu.be/KHXPJIsxoLE

Attribution

- Prof. Steven Farmer (Sonoma State University)