15.11.2 Finding the Vapor Pressure of a Solution (Ionic-Nonvolatile Solute) (Video)

This project was preformed to supply **Libretext Authors** with videos on General Chemistry topics which can be used to enhance their projects. Also, these videos are meant to act as a learning resource for **all General Chemistry students**.

Video Topics

An ionic-nonvolatile solute will also cause the vapor pressure of the solvent to decrease.
Because the solute is nonvolatile it will not create a vapor pressure of its own.
However, we will need to adjust our calculations to account for the Van't Hoff factor of the solute (i). Examples are salts such as NaCl or NaOH.

Raoult's Law Pa = XaPao

Pa = Vapor pressure of solution
Xa = Mol fraction of the solvent
Pao = Vapor pressure of the pure solvent

Link to Video

Finding the Vapor Pressure of a Solution (Ionic-Nonvolatile Solute): https://youtu.be/sRBaRXsql9s

Attribution

- Prof. Steven Farmer (Sonoma State University)