

17.3.3 Le Châtelier's Principle (Changes in Temperature) (Video)

This project was preformed to supply **Libretext authors** with videos on General Chemistry topics which can be used to enhance their projects. Also, these videos are meant to act as a learning resource for **all General Chemistry students**.

Video Topics

Treat an exothermic reactions as having heat as a product and endothermic reactions as having heat as a reactant. Adding or removing heat is like changing a reactant or product. $\text{Heat} + \text{N}_{2(g)} + 3 \text{H}_{2(g)} \rightarrow 2 \text{NH}_{3(g)}$ $\Delta H = + 100 \text{ kJ/mol}$ ΔH is positive so the reaction is endothermic. Treat the reaction as if heat is a reactant. Which way will the reaction shift to reach equilibrium if the reaction temperature was increased? The reaction shifts right. Decreasing the reaction temp? Reaction goes left.

Link to Video

Le Châtelier's Principle (Changes in Temperature): <https://youtu.be/-P5uGuJZ-r8>



Attribution

- Prof. Steven Farmer ([Sonoma State University](#))

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