16.10 Zero-Order Reactions (Video)

This project was preformed to supply **Libretext authors** with videos on General Chemistry topics which can be used to enhance their projects. Also, these videos are meant to act as a learning resource for **all General Chemistry students**.

Video Topics

The zero order rate law expression is: $[A]_t = -kt + [A]_0$
t = time in seconds
$[A]_0$ = initial concentration of A
$[A]_t$ = concentration of A at time t

This video contains the solution to the following question:

For a zero order reaction. If the initial $[N_2O_2]$ is 0.10 M what would be the final $[N_2O_2]$ after 160 seconds?

Link to Video

Zero-Order Reactions: https://youtu.be/64i7uYsVsSs

Attribution

- Prof. Steven Farmer (Sonoma State University)