

## 19.7.6 Summary of the pH Curve for a Strong Acid/Strong Base Titration (Video)

This project was preformed to supply **Libretext authors** with videos on General Chemistry topics which can be used to enhance their projects. Also, these videos are meant to act as a learning resource for **all General Chemistry students**.

### Video Topics

This video contains the summary of the different parts for a titration of a strong acid with a strong base.

At the start, you can find the initial pH by using:

$$\text{pH} = -\text{Log}\{\text{H}_3\text{O}^+\}$$

Before the equivalence point, subtract mol  $\text{OH}^-$  added from mol  $\text{H}_3\text{O}^+$  initial to find # mol  $\text{H}_3\text{O}^+$  remaining. Then divide by  $V_{\text{total}}$  to find  $\{\text{H}_3\text{O}^+\}$ .

**At the equivalence point pH = 7**

After the equivalence point, subtract mol  $\text{H}_3\text{O}^+$  initial from

mol  $\text{OH}^-$  added to find # mol  $\text{OH}^-$  remaining. Then divide by  $V_{\text{total}}$  to find  $\{\text{OH}^-\}$ .

### Link to Video

Summary of the pH Curve for a Strong Acid/Strong Base Titration: <https://youtu.be/hBo5w9uJt-c>



### Attribution

- Prof. Steven Farmer (Sonoma State University)

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