

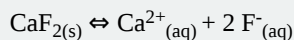
20.4 The Common Ion Effect in Solubility Products (Video)

This project was preformed to supply **Libretext Authors** with videos on General Chemistry topics which can be used to enhance their projects. Also, these videos are meant to act as a learning resource for **all General Chemistry students**.

Video Topics

The presence of a common ion greatly decreases the solubility of a salt.

For the solubility of insoluble salt $\text{CaF}_{2(s)}$.



The common ions are Ca^{2+} and F^{-} . These can be introduced as a soluble salt such as $\text{Ca}(\text{NO}_3)_{2(s)}$ or $\text{NaF}_{(s)}$.

If we add a $\text{NaF}_{(aq)}$ solution to a saturated solution of $\text{CaF}_{2(aq)}$ we are adding $\text{F}^{-}_{(aq)}$ to the equilibrium.

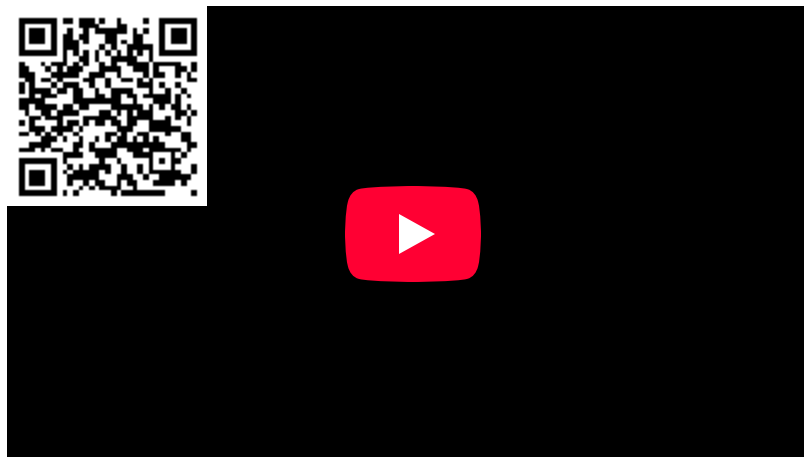
Le Chatlier's principal says the equilibrium will shift left.

This means that solid $\text{CaF}_{2(s)}$ will form (Precipitate) and the $\{\text{Ca}^{2+}\}$ will decreases.

This video contains examples of determining the solubility of a salt (s) if given the solubility product (K_{sp}) and the concentration of a common ion of the salt.

Link to Video

The Common Ion Effect in Solubility Products: https://youtu.be/_P3wozLs0Tc



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- Prof. Steven Farmer (Sonoma State University)

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