

Chapter 9.3 Magnetic Quantum Number (m_l) and Spin Quantum Number (m_s) (Video)

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Video Topics

The value m_l is the magnetic quantum number. The possible m_l values follow the equation: $-l, +1, +2, +3, +l$. The m_l 's values represent the orbitals in a subshell which actually contain electrons. Because each orbital (m_l) value can contain 2 electrons we can see how many electrons can be contained in a particular orbital subshell. The value m_s is called the spin quantum number. m_s refers to the electron spin of each electron. This can be $+1/2$ or $-1/2$.

Link to Video

Magnetic Quantum Number (m_l) & Spin Quantum Number (m_s):
<https://youtu.be/gbmGVUXBOBk>



Attribution

- Prof. Steven Farmer (Sonoma State University)

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