

18.1 Definitions of Acids and Bases (Video)

This project was preformed to supply **Libretext authors** with videos on General Chemistry topics which can be used to enhance their projects. Also, these videos are meant to act as a learning resource for **all General Chemistry students**.

Video Topics

This video discusses the different definitions of acids and bases used in chemistry.

Arrhenius Theory of acids and bases

Acids increases $\{H^+\}$ when dissolved in water. H^+ is called a hydrogen ion or proton.

Bases increases $\{OH^-\}$ when dissolved in water. OH^- is called the hydroxide ion

This theory has problems because it neglects NH_3 as a base and does not consider the solvent H_2O .

Brøsted-Lowry theory of acids and bases

Acid: Proton donor

Base: Proton acceptor

Lewis Definition of acids and bases

Lewis Acid = Accepts an electron pair to form a bond

Lewis Base = Donates an electron pair to form a bond

Link to Video

Definitions of Acids and Bases: <https://youtu.be/r8reN0CSIHw>



Attribution

- Prof. Steven Farmer (Sonoma State University)

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