

### 19.7.1 Initial pH for a Strong Acid/Strong Base Titration (Video)

This project was preformed to supply **Libretext authors** with videos on General Chemistry topics which can be used to enhance their projects. Also, these videos are meant to act as a learning resource for **all General Chemistry students**.

#### Video Topics

**This video is the first of a series of theoretical calculation for a titration of a strong acid with a strong base.**

**You start with 50.0 mL of a 0.100 M solution of HCl.**

**What is the initial pH?**

Because HCl is a strong acid we can say:

$$\{\text{HCl}\} = \{\text{H}_3\text{O}^+\} = 0.100 \text{ M}$$

So the pH can be found directly

$$\text{pH} = -\text{Log}\{0.100 \text{ M}\} = \mathbf{1.00}$$

#### Link to Video

Initial pH for a Strong Acid/Strong Base Titration: <https://youtu.be/wp3QFchYasM>



#### Attribution

- Prof. Steven Farmer (Sonoma State University)

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