

## 15.7 Colligative Properties in Solutions (Video)

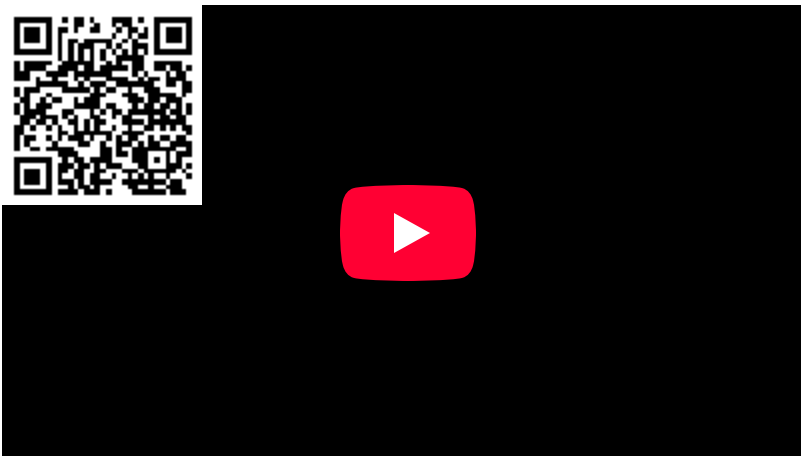
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### Video Topics

Properties dependent on the concentration of solute particles are called colligative properties. The Van't Hoff factor ( $i$ ) compensates for the fact that some compounds dissolve to more than one particle.  $i = 1$  for non-electrolytes. Sugar, alcohol For strong electrolytes (ionic compounds) (Metal bonded to a Non-metal).  $i$  = number to ions created when the electrolyte is dissolved

### Link to Video

**Colligative Properties in Solutions:** <https://youtu.be/mkHyVNswGog>



### Attribution

- Prof. Steven Farmer ([Sonoma State University](#))

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