

9.2 Principal quantum number (n) and Orbital angular momentum (l): The Orbital Subshell (Video)

This project was performed to supply **Libretext Authors** with videos on General Chemistry topics which can be used to enhance their projects. Also, these videos are meant to act as a learning resource for **all General Chemistry students**.

Video Topics

Principal quantum number (n). This number can be any positive, nonzero integral value. As (n) increases the energy of the orbital decreases and the distance from the nucleus increases. This is the same n value that we saw with the Bohr atom. Orbital angular momentum (l). Defines the type of orbital subshell. An orbital subshell can contain multiple orbitals of the same type. Possible values for l are: (l) can be zero or any positive integer but not larger than (n-1). $l = 0, 1, 2, 3, 4, \dots, (n-1)$. This video contains a sample problem, which involves these concepts.

Link to Video

Principal quantum number (n) & Orbital angular momentum (l): The Orbital Subshell: <https://youtu.be/ms7WR149fAY>



Attribution

- Prof. Steven Farmer ([Sonoma State University](#))

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