

21.7 The Nernst Equation (Video)

This project was preformed to supply **Libretext Authors** with videos on General Chemistry topics which can be used to enhance their projects. Also, these videos are meant to act as a learning resource for **all General Chemistry students**.

Video Topics

E_{cell} at non-standard concentrations can be found using the the Nernst Equation

$$E_{\text{cell}} = E^{\circ}_{\text{cell}} - (0.0592 \text{ V} / N) \text{Log}Q$$

To find the value of Q you must plug the given concentrations into the equilibrium expression for the reaction.

N is the number of electrons transferred in the balanced half reaction.

Link to Video

The Nernst Equation: <https://youtu.be/x1wijHu1zXY>



Attribution

- Prof. Steven Farmer ([Sonoma State University](#))

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