

14.4 The Clausius-Clapeyron Equation (Video)

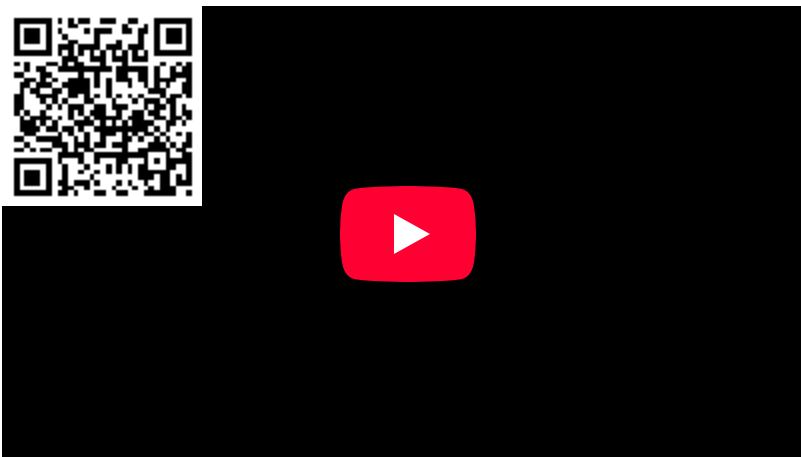
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Video Topics

As the temperature of a liquid increases the vapor pressure also increases. This relationship is shown by the Clausius-Clapeyron Equation. This equation allows for a vapor pressure of a liquid to be calculated at a new temperature. $\ln(P_2/P_1) = -\Delta H_{\text{vap}}/R(1/T_2 - 1/T_1)$ Temperatures are in Kelvin. R is the gas law constant 8.3145 J/Mol K. ΔH_{vap} is the heat of vaporization of the liquid in J/Mol. The units for ΔH_{vap} must be in J to match the units in R. P is the vapor pressure of the liquid. P can be in any unit of pressure.

Link to Video

Determining the Products for Precipitation Reactions: <https://youtu.be/r0kYeZVuTAM>



Attribution

- Prof. Steven Farmer (Sonoma State University)

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