

5.6: In-Text References

1. https://en.Wikipedia.org/wiki/Vladimir_Markovnikov ↩
 2. In fact ALL reactions are reversible in theory (this is called the principle of microscopic reversibility, https://en.Wikipedia.org/wiki/Micros..._reversibility . However, in practice it is extremely difficult to reverse some reactions in the laboratory. For example, combustion of hydrocarbons is not something you would try to reverse in the lab, since the products are gases and will be very difficult to bring back together, and the reaction is highly exergonic. However, plants can do the reverse reaction quite well using energy from sunlight. ↩
 3. While we have seen that alkenes can add water (as H^+ and OH^-) across a double bond, this is not classified as an oxidation. There is no change in oxidation state of the O or H that add to the double bonded carbons. ↩
 4. epoxide hydrolases: <http://www.annualreviews.org/doi/pdf....120403.095920> ↩
 5. Ozone is generated during the reaction by using a special generator because it is too reactive to store. It is generated in the same way that lightning generates ozone—by passing a spark of electric current through oxygen. ↩
 6. The reducing agent is present to stop “over oxidation” to the carboxylic acid. ↩
 7. Tautomers: evil twins of the bases!: <http://blc.arizona.edu/courses/181La...Tautomers.html> ↩
-

This page titled [5.6: In-Text References](#) is shared under a [CC BY-NC-SA 4.0](#) license and was authored, remixed, and/or curated by [Melanie M. Cooper & Michael W. Klymkowsky](#).