

4.5: In-Text References

1. i.e. the change in concentration of reactant divided by the time $d[A]/dt$ (where A is one of the reactants). For more information, see CLUE Chapter 8. [↩](#)
 2. See Chemistry, Life, the Universe and Everything: Chapter 8 for a longer discussion. [↩](#)
 3. This hypothesis is known as Hammond's postulate. It states that for an endothermic reaction, the transition state is closer in structure to the product (in this case the carbocation), and for an exothermic reaction the transition state is closer in structure to the reactant. See https://en.Wikipedia.org/wiki/Hammond%27s_postulate [↩](#)
 4. As we will discuss shortly, this is because the most substituted alkene is the most stable. [↩](#)
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