

3.4: In-Text References

1. Recall that this is not the case for sp^2 or sp hybridized carbon, which also form pi bonds, where the orbital overlap is destroyed by rotation. Therefore, as we will see later, a much higher input of energy is required for compounds with pi bonds to rotate the bonding carbons relative to each other. ↩
2. See Chapter 1 and Chapter 4 in CLUE text. ↩
3. https://en.Wikipedia.org/wiki/Tartaric_acid ↩
4. Wikipedia on fluoxetine: <https://en.Wikipedia.org/wiki/Fluoxetine> ↩

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