

CHAPTER OVERVIEW

15: Equilibria of Other Reaction Classes

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General Chemistry

We previously learned about aqueous solutions and their importance, as well as about solubility rules. While this gives us a picture of solubility, that picture is not complete if we look at the rules alone. Solubility equilibrium, which we will explore in this chapter, is a more complex topic that allows us to determine the extent to which a slightly soluble ionic solid will dissolve, and the conditions under which precipitation.

[15.1: Precipitation and Dissolution](#)

[15.2: Lewis Acids and Bases](#)

[15.3: Coupled Equilibria](#)

[15.E: Equilibria of Other Reaction Classes \(Exercises\)](#)

Thumbnail: Lead (II) iodide precipitates when potassium iodide is mixed with lead (II) nitrate. (CC BY-SA 3.0 Unported; PRHaney via [Wikipedia](#))

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