

## 16.3: Half-Life

### Learning Objectives

- Define *half-life*.
- Determine the amount of radioactive substance remaining after a given number of half-lives.

Whether or not a given isotope is radioactive is a characteristic of that particular isotope. Some isotopes are stable indefinitely, while others are radioactive and decay through a characteristic form of emission. As time passes, less and less of the radioactive isotope will be present, and the level of radioactivity decreases. An interesting and useful aspect of radioactive decay is **half life**, which is the amount of time it takes for one-half of a radioactive isotope to decay. The half-life of a specific radioactive isotope is constant; it is unaffected by conditions and is independent of the initial amount of that isotope.

Consider the following example. Suppose we have 100.0 g of tritium (a radioactive isotope of hydrogen). It has a half-life of 12.3 y. After 12.3 y, half of the sample will have decayed from hydrogen-3 to helium-3 by emitting a beta particle, so that only 50.0 g of the original tritium remains. After another 12.3 y—making a total of 24.6 y—another half of the remaining tritium will have decayed, leaving 25.0 g of tritium. After another 12.3 y—now a total of 36.9 y—another half of the remaining tritium will have decayed, leaving 12.5 g. This sequence of events is illustrated in Figure 16.3.1 - Radioactive Decay.

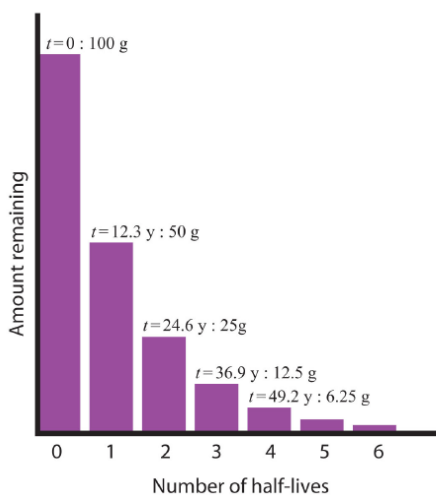


Figure 16.3.1 Radioactive Decay. During each successive half-life, half of the initial amount will radioactively decay.

We can determine the amount of a radioactive isotope remaining after a given number half-lives by using the following expression:

$$\text{amount remaining} = \text{initial amount} \times \left(\frac{1}{2}\right)^n$$

where  $n$  is the number of half-lives. This expression works even if the number of half-lives is not a whole number.

### ✓ Example 16.3.1

The half-life of fluorine-20 is 11.0 s. If a sample initially contains 5.00 g of fluorine-20, how much remains after 44.0 s?

#### Solution

If we compare the time that has passed to the isotope's half-life, we note that 44.0 s is exactly 4 half-lives, so we use the previous expression  $n = 4$ . Substituting and solving results in the following:

$$\begin{aligned}\text{amount remaining} &= 5.00 \text{ g} \times \left(\frac{1}{2}\right)^4 \\ &= 5.00 \text{ g} \times \left(\frac{1}{16}\right) \\ &= 0.313 \text{ g}\end{aligned}$$

Less than one-third of a gram of fluorine-20 remains.

### ✓ Exercise 16.3.1

The half-life of titanium-44 is 60.0 y. A sample of titanium contains 0.600 g of titanium-44. How much remains after 240.0 y?

**Answer**

0.0375 g

Half-lives of isotopes range from fractions of a microsecond to billions of years. Table 16.3.1- Half-Lives of Various Isotopes, lists the half-lives of some isotopes.

Table 16.3.1 Half-Lives of Various Isotopes

Isotope	Half-Life
$^3\text{H}$	12.3 y
$^{14}\text{C}$	5730 y
$^{40}\text{K}$	$1.26 \times 10^9$ y
$^{51}\text{Cr}$	27.70 d
$^{90}\text{Sr}$	29.1 y
$^{131}\text{I}$	8.04 d
$^{222}\text{Rn}$	3.823 d
$^{235}\text{U}$	$7.04 \times 10^8$ y
$^{238}\text{U}$	$4.47 \times 10^9$ y
$^{241}\text{Am}$	432.7 y
$^{248}\text{Bk}$	23.7 h
$^{260}\text{Sg}$	4 ms

### 📌 Chemistry is Everywhere: Radioactive Elements in the Body

You may not think of yourself as radioactive, but you are. A small portion of certain elements in the human body are radioactive and constantly undergo decay. The following table summarizes radioactivity in the normal human body.

Table with four columns and eight rows. The first column on the left is labeled as Radioactive Isotope, underneath in the rows has radioactive isotopes. The second column is labeled as Half-Life (y), underneath in the rows has half-lives. The third column is labeled as Isotope Mass in the Body (g), underneath in the rows has isotope masses. The fourth column on right is labeled as Activity in the Body (decays/s), underneath in the rows has the activities in the body as the decays.

Radioactive Isotope	Half-Life (y)	Isotope Mass in the Body (g)	Activity in the Body (decays/s)
$^{40}\text{K}$	$1.26 \times 10^9$	0.0164	4,340
$^{14}\text{C}$	5,730	$1.6 \times 10^{-8}$	3,080

Radioactive Isotope	Half-Life (y)	Isotope Mass in the Body (g)	Activity in the Body (decays/s)
$^{87}\text{Rb}$	$4.9 \times 10^{10}$	0.19	600
$^{210}\text{Pb}$	22.3	$5.4 \times 10^{-10}$	15
$^3\text{H}$	12.3	$2 \times 10^{-14}$	7
$^{238}\text{U}$	$4.47 \times 10^9$	$1 \times 10^{-4}$	5
$^{228}\text{Ra}$	5.76	$4.6 \times 10^{-14}$	5
$^{226}\text{Ra}$	1,620	$3.6 \times 10^{-11}$	3

The average human body experiences about 8,000 radioactive decays.

Most of the radioactivity in the human body comes from potassium-40 and carbon-14. Potassium and carbon are two elements that we absolutely cannot live without, so unless we can remove all the radioactive isotopes of these elements, there is no way to escape at least some radioactivity. There is debate about which radioactive element is more problematic. There is more potassium-40 in the body than carbon-14, and it has a much longer half-life. Potassium-40 also decays with about 10 times more energy than carbon-14, making each decay potentially more problematic. However, carbon is the element that makes up the backbone of most living molecules, making carbon-14 more likely to be present around important molecules, such as proteins and DNA molecules. Most experts agree that while it is foolhardy to expect absolutely no exposure to radioactivity, we can and should minimize exposure to excess radioactivity.

What if the elapsed time is not an exact number of half-lives? We can still calculate the amount of material we have left, but the equation is more complicated. The equation is

$$\text{amount remaining} = (\text{amount initial}) \times e^{-0.693t/t_{1/2}}$$

where  $e$  is the base of natural logarithms (2.71828182...),  $t$  is the elapsed time, and  $t_{1/2}$  is the half-life of the radioactive isotope. The variables  $t$  and  $t_{1/2}$  should have the same units of time, and you may need to make sure you know how to evaluate natural-logarithm powers on your calculator (for many calculators, there is an "inverse logarithm" function that you can use; consult your instructor if you are not sure how to use your calculator). Although this is a more complicated formula, the length of time  $t$  need not be an exact multiple of half-lives.

#### ✓ Example 16.3.2

The half-life of fluorine-20 is 11.0 s. If a sample initially contains 5.00 g of fluorine-20, how much remains after 60.0 s?

##### Solution

Although similar to Example 3, the amount of time is not an exact multiple of a half-life. Here we identify the initial amount as 5.00 g,  $t = 60.0$  s, and  $t_{1/2} = 11.0$  s. Substituting into the equation:

$$\text{amount remaining} = (5.00 \text{ g}) \times e^{-(0.693)(60.0 \text{ s})/11.0 \text{ s}}$$

Evaluating the exponent (and noting that the s units cancel), we get

$$\text{amount remaining} = (5.00 \text{ g}) \times e^{-3.78}$$

Solving, the amount remaining is 0.114 g. (You may want to verify this answer to confirm that you are using your calculator properly.)

#### ? Exercise 16.3.2

The half-life of titanium-44 is 60.0 y. A sample of titanium contains 0.600 g of titanium-44. How much remains after 100.0 y?

##### Answer

0.189 g

## Key Takeaways

- Natural radioactive processes are characterized by a half-life, the time it takes for half of the material to decay radioactively.
- The amount of material left over after a certain number of half-lives can be easily calculated.

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