

14.2: Arrhenius Acids and Bases

Learning Objective

- Identify an Arrhenius acid and an Arrhenius base.
- Write the chemical reaction between an Arrhenius acid and an Arrhenius base.

Historically, the first chemical definition of an acid and a base was put forward by Svante Arrhenius, a Swedish chemist, in 1884. An **Arrhenius acid** is a compound that increases the H^+ ion concentration in an aqueous solution. The H^+ ion is just a bare proton, and it is rather clear that bare protons are not floating around in an aqueous solution. Instead, chemistry has defined the **hydronium ion** (H_3O^+) as the actual chemical species that represents an H^+ ion. H^+ ions and H_3O^+ ions are often considered interchangeable when writing chemical equations (although a properly balanced chemical equation should also include the additional H_2O). Classic Arrhenius acids can be considered ionic compounds in which H^+ is the cation. Table 14.2.1 lists some Arrhenius acids and their names.

Table 14.2.1 Some Arrhenius Acids

Formula	Name
$\text{HC}_2\text{H}_3\text{O}_2$ (also written CH_3COOH)	acetic acid
HClO_3	chloric acid
HCl	hydrochloric acid
HBr	hydrobromic acid
HI	hydriodic acid
HF	hydrofluoric acid
HNO_3	nitric acid
$\text{H}_2\text{C}_2\text{O}_4$	oxalic acid
HClO_4	perchloric acid
H_3PO_4	phosphoric acid
H_2SO_4	sulfuric acid
H_2SO_3	sulfurous acid

An **Arrhenius base** is a compound that increases the OH^- ion concentration in aqueous solution. Ionic compounds of the OH^- ion are classic Arrhenius bases.

✓ Example 14.2.1

Identify each compound as an Arrhenius acid, an Arrhenius base, or neither.

- HNO_3
- CH_3OH
- $\text{Mg}(\text{OH})_2$

Solution

- This compound is an ionic compound between H^+ ions and NO_3^- ions, so it is an Arrhenius acid.
- Although this formula has an OH in it, we do not recognize the remaining part of the molecule as a cation. It is neither an acid nor a base. (In fact, it is the formula for methanol, an organic compound.)
- This formula also has an OH in it, but this time we recognize that the magnesium is present as Mg^{2+} cations. As such, this is an ionic compound of the OH^- ion and is an Arrhenius base.

? Exercise 14.2.1

Identify each compound as an Arrhenius acid, an Arrhenius base, or neither.

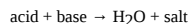
- KOH
- H_2SO_4
- C_2H_6

Answer

- Arrhenius base
- Arrhenius acid
- neither

Acids have some properties in common. They turn litmus, a plant extract, red. They react with some metals to give off H_2 gas. They react with carbonate and hydrogen carbonate salts to give off CO_2 gas. Acids that are ingested typically have a sour, sharp taste. (The name *acid* comes from the Latin word *acidus*, meaning "sour.") Bases also have some properties in common. They are slippery to the touch, turn litmus blue, and have a bitter flavor if ingested.

Acids and bases have another property: they react with each other to make water and an ionic compound called a salt. A **salt**, in chemistry, is any ionic compound made by combining an acid with a base. A reaction between an acid and a base is called a **neutralization reaction** and can be represented as:



The stoichiometry of the balanced chemical equation depends on the number of H^+ ions in the acid and the number of OH^- ions in the base.

✓ Example 14.2.2

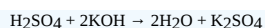
Write the balanced chemical equation for the neutralization reaction between H_2SO_4 and KOH . What is the name of the salt that is formed?

Solution

The general reaction is as follows:



Because the acid has two H^+ ions in its formula, we need two OH^- ions to react with it, making two H_2O molecules as product. The remaining ions, K^+ and SO_4^{2-} , make the salt potassium sulfate (K_2SO_4). The balanced chemical reaction is as follows:



? Exercise 14.2.2

Write the balanced chemical equation for the neutralization reaction between HCl and $\text{Mg}(\text{OH})_2$. What is the name of the salt that is formed?

Answer

$2\text{HCl} + \text{Mg}(\text{OH})_2 \rightarrow 2\text{H}_2\text{O} + \text{MgCl}_2$; magnesium chloride

Key Takeaways

- An Arrhenius acid is a compound that increases the H^+ ion concentration in aqueous solution.
- An Arrhenius base is a compound that increases the OH^- ion concentration in aqueous solution.
- The reaction between an Arrhenius acid and an Arrhenius base is called neutralization and results in the formation of water and a salt.

This page titled [14.2: Arrhenius Acids and Bases](#) is shared under a [CC BY-NC-SA 3.0](#) license and was authored, remixed, and/or curated by [Theodore Chan](#) via [source content](#) that was edited to the style and standards of the LibreTexts platform.

- [12.2: Arrhenius Acids and Bases](#) by Anonymous is licensed [CC BY-NC-SA 3.0](#). Original source: <https://2012books.lardbucket.org/books/beginning-chemistry>.

