

## 9.E: Energy and Chemistry (Exercises)

### Exercises (Energy, and heat)

1. Define energy. How is work related to energy?
2. Give two units of energy and indicate which one is preferred.
3. Express the quantity of 422 J in calories.
4. Express the quantity of 3.225 kJ in calories.
5. Express the quantity 55.69 cal in joules.
6. Express the quantity 965.33 kcal in joules.
7. How does a Calorie differ from a calorie?
8. Express the quantity 965.33 Cal in joules.
9. What is the law of conservation of energy?
10. What does the word conserved mean as applied to the law of conservation of energy?

### Answers

1. Energy is the ability to do work. Work is a form of energy.
3. 101 cal
5. 233.0 J
7. A Calorie is actually a kilocalorie, or 1,000 calories.
9. The total energy of an isolated system does not increase or decrease.

### Exercises (Work)

1. Give two definitions of work.
2. What is the sign on work when a sample of gas increases its volume? Explain why work has that sign.
3. What is the work when a gas expands from 3.00 L to 12.60 L against an external pressure of 0.888 atm?
4. What is the work when a gas expands from 0.666 L to 2.334 L against an external pressure of 2.07 atm?
5. What is the work when a gas contracts from 3.45 L to 0.97 L under an external pressure of 0.985 atm?
6. What is the work when a gas contracts from 4.66 L to 1.22 L under an external pressure of 3.97 atm?
7. Like work, the sign on heat can be positive or negative. What is happening to the total energy of a system if heat is positive?
8. Like work, the sign on heat can be positive or negative. What is happening to the total energy of a system if heat is negative?
9. What is the heat when 55.6 g of Fe increase temperature from 25.6°C to 177.9°C ? The heat capacity of Fe is in Table 7.1 "Specific Heats of Various Substances".
10. What is the heat when 0.444 g of Au increases temperature from 17.8°C to 222.5°C ? The heat capacity of Au is in Table 7.1 "Specific Heats of Various Substances".
11. What is the heat when 245 g of H<sub>2</sub>O cool from 355 K to 298 K ? The heat capacity of H<sub>2</sub>O is in Table 7.1 "Specific Heats of Various Substances".
12. What is the heat when 100.0 g of Mg cool from 725 K to 552 K ? The heat capacity of Mg is in Table 7.1 "Specific Heats of Various Substances".
13. It takes 452 J of heat to raise the temperature of a 36.8 g sample of a metal from 22.9°C to 98.2°C. What is the heat capacity of the metal?
14. It takes 2,267 J of heat to raise the temperature of a 44.5 g sample of a metal from 33.9°C to 288.3°C. What is the heat capacity of the metal?

15. An experimenter adds 336 J of heat to a 56.2 g sample of Hg. What is its change in temperature? The heat capacity of Hg is in Table 7.1 "Specific Heats of Various Substances".
16. To a 0.444 g sample of H<sub>2</sub>O, 23.4 J of heat are added. What is its change in temperature? The heat capacity of H<sub>2</sub>O is in Table 7.1 "Specific Heats of Various Substances".
17. An unknown mass of Al absorbs 187.9 J of heat and increases its temperature from 23.5°C to 35.6°C. What is the mass of the aluminum? How many moles of aluminum is this?
18. A sample of He goes from 19.4°C to 55.9°C when 448 J of energy are added. What is the mass of the helium? How many moles of helium is this?

### Answers

1. Work is a force acting through a distance or a volume changing against some pressure.
3. -864 J
5. 248 J
7. When heat is positive, the total energy of the system is increasing.
9.  $3.80 \times 10^3$  J
11. -58,400 J
13. 0.163 J/g · °C
15. 43.0°C
17. 17.3 g; 0.640 mol

### Exercises (Enthalpy)

1. Under what circumstances are  $q$  and  $\Delta H$  the same?
2. Under what circumstances are  $q$  and  $\Delta H$  different?
3. Hydrogen gas and chlorine gas react to make hydrogen chloride gas with an accompanying enthalpy change of -184 kJ. Write a properly balanced thermochemical equation for this process.
4. Propane (C<sub>3</sub>H<sub>8</sub>) reacts with elemental oxygen gas to produce carbon dioxide and liquid water with an accompanying enthalpy change of -2,220 kJ. Write a properly balanced thermochemical equation for this process.
5. Nitrogen gas reacts with oxygen gas to make NO(g) while absorbing 180 kJ. Write a properly balanced thermochemical equation for this process.
6. Solid sodium reacts with chlorine gas to make solid sodium chloride while giving off 772 kJ. Write a properly balanced thermochemical equation for this process.
7. Hydrogen gas and chlorine gas react to make hydrogen chloride gas with an accompanying enthalpy change of -184 kJ. Is this process endothermic or exothermic?
8. Propane (C<sub>3</sub>H<sub>8</sub>) reacts with elemental oxygen gas to produce carbon dioxide while giving off 2,220 kJ of energy. Is this process endothermic or exothermic?
9. Nitrogen gas reacts with oxygen gas to make NO(g) while absorbing 180 kJ. Is this process exothermic or endothermic?
10. Sodium metal can react with nitrogen to make sodium azide (NaN<sub>3</sub>) with a  $\Delta H$  of 21.72 kJ. Is this process exothermic or endothermic?
11. Draw an energy level diagram for the chemical reaction in Exercise 8.
12. Draw an energy level diagram for the chemical reaction in Exercise 9. (See Figure 7.3 "Reaction Energy" for an example.)
13. In a 250 mL solution, 0.25 mol of KOH(aq) and 0.25 mol of HNO<sub>3</sub>(aq) are combined. The temperature of the solution increases from 22.5°C to 35.9°C. Assume the solution has the same density and heat capacity of water. What is the heat of the reaction, and what is the  $\Delta H$  of the reaction on a molar basis?

14. In a 600 mL solution, 0.50 mol of  $\text{Ca}(\text{OH})_2(\text{aq})$  and 0.50 mol of  $\text{H}_2\text{SO}_4(\text{aq})$  are combined. The temperature of the solution increases by  $22.3^\circ\text{C}$ . What is the heat of the reaction, and what is the  $\Delta H$  of the reaction on a molar basis? Assume the solution has the same density and heat capacity of water.

15. To warm 400.0 g of  $\text{H}_2\text{O}$ , 0.050 mol of ethanol ( $\text{C}_2\text{H}_5\text{OH}$ ) is burned. The water warms from  $24.6^\circ\text{C}$  to  $65.6^\circ\text{C}$ . What is the heat of the reaction, and what is the  $\Delta H$  of the reaction on a molar basis?

16. To warm 100.0 g of  $\text{H}_2\text{O}$ , 0.066 mol beeswax is burned. The water warms from  $21.4^\circ\text{C}$  to  $25.5^\circ\text{C}$ .

What is the heat of the reaction, and what is the  $\Delta H$  of the reaction on a molar basis?

### Answers

1. under conditions of constant pressure

3.  $\text{H}_2(\text{g}) + \text{Cl}_2(\text{g}) \rightarrow 2\text{HCl}(\text{g}) \Delta H = -184 \text{ kJ}$

5.  $\text{N}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2\text{NO}(\text{g}) \Delta H = 180 \text{ kJ}$

7. exothermic

9. endothermic

13. heat of reaction =  $-14.0 \text{ kJ}$ ;  $\Delta H = -56.0 \text{ kJ/mol}$  of reactants

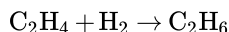
15. heat of reaction =  $-68.6 \text{ kJ}$ ;  $\Delta H = -1,370 \text{ kJ/mole}$  of ethanol

### Exercises (Hess's Law)

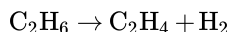
1. Define Hess's law.

2. What does Hess's law require us to do to the  $\Delta H$  of a thermochemical equation if we reverse the equation?

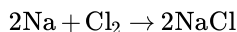
3. If the  $\Delta H$  for



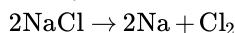
is  $-65.6 \text{ kJ}$ , what is the  $\Delta H$  for this reaction?



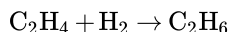
4. If the  $\Delta H$  for



is  $-772 \text{ kJ}$ , what is the  $\Delta H$  for this reaction:



5. If the  $\Delta H$  for



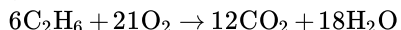
is  $-65.6 \text{ kJ}$ , what is the  $\Delta H$  for this reaction?



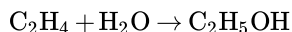
6. If the  $\Delta H$  for



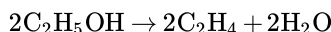
is  $-2,650 \text{ kJ}$ , what is the  $\Delta H$  for this reaction?



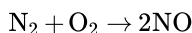
7. The  $\Delta H$  for



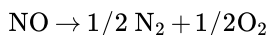
is  $-44 \text{ kJ}$ . What is the  $\Delta H$  for this reaction?



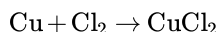
8. The  $\Delta H$  for



is  $181 \text{ kJ}$ . What is the  $\Delta H$  for this reaction?



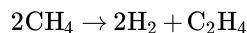
9. Determine the  $\Delta H$  for the reaction



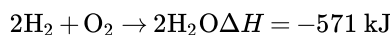
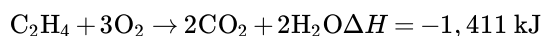
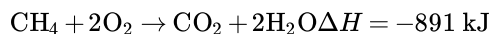
given these data:



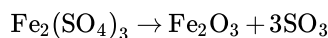
10. Determine  $\Delta H$  for the reaction



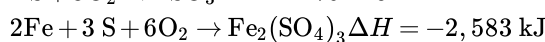
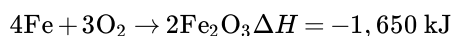
given these data:



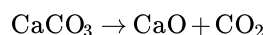
11. Determine  $\Delta H$  for the reaction



given these data:



12. Determine  $\Delta H$  for the reaction



given these data:



## Answers

1. If chemical equations are combined, their energy changes are also combined.

3.  $\Delta H = 65.6 \text{ kJ}$

5.  $\Delta H = -131.2 \text{ kJ}$

7.  $\Delta H = 88 \text{ kJ}$

9.  $\Delta H = -220 \text{ kJ}$

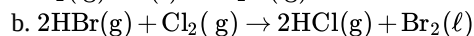
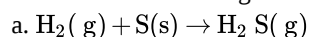
11.  $\Delta H = 570 \text{ kJ}$

## Exercises (Formation Reactions)

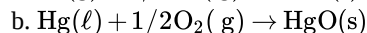
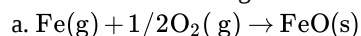
1. Define formation reaction and give an example.

2. Explain the importance of formation reactions in thermochemical equations.

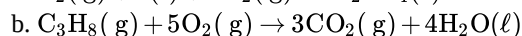
3. Which of the following reactions is a formation reaction? If it is not a formation reaction, explain why.



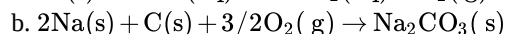
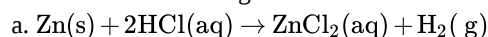
4. Which of the following reactions is a formation reaction? If it is not a formation reaction, explain why.



5. Which of the following reactions is a formation reaction? If it is not a formation reaction, explain why.



6. Which of the following reactions is a formation reaction? If it is not a formation reaction, explain why.



7. Write a proper formation reaction for each substance.

- $\text{H}_3\text{PO}_4(\text{s})$
- $\text{Na}_2\text{O}(\text{s})$
- $\text{C}_3\text{H}_7\text{OH}(\ell)$

8. Write a proper formation reaction for each substance.

- $\text{N}_2\text{O}_5(\text{g})$
- $\text{BaSO}_4(\text{s})$
- $\text{Fe}(\text{OH})_3(\text{s})$

9. Write a proper formation reaction for each substance.

- $\text{C}_{12}\text{H}_{22}\text{O}_{11}(\text{s})$
- $\text{Zn}(\text{NO}_3)_2(\text{s})$
- $\text{Al}(\text{OH})_3(\text{s})$

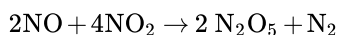
10. Write a proper formation reaction for each substance.

- $\text{Na}_2\text{O}_2(\text{s})$
- $\text{PCl}_5(\text{g})$

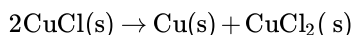
11. Write this reaction in terms of formation reactions.



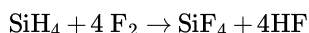
12. Write this reaction in terms of formation reactions.



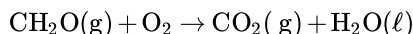
13. Write this reaction in terms of formation reactions.



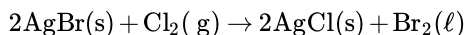
14. Write this reaction in terms of formation reactions.



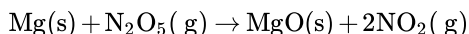
15. Determine the enthalpy change of this reaction. Data can be found in Table 7.2 "Enthalpies of Formation for Various Substances".



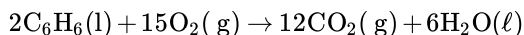
16. Determine the enthalpy change of this reaction. Data can be found in Table 7.2 "Enthalpies of Formation for Various Substances".



17. Determine the enthalpy change of this reaction. Data can be found in Table 7.2 "Enthalpies of Formation for Various Substances".



18. Determine the enthalpy change of this reaction. Data can be found in Table 7.2 "Enthalpies of Formation for Various Substances".



## Answers

1. A formation reaction is a reaction that produces one mole of a substance from its elements.

Example:  $\text{C}(\text{s}) + \text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g})$

3. a. formation reaction

b. It is not the formation of a single substance, so it is not a formation reaction.

5. a. formation reaction

b. It is not the formation of a single substance, so it is not a formation reaction.

7. a.  $3/2\text{H}_2(\text{g}) + \text{P}(\text{s}) + 2\text{O}_2(\text{g}) \rightarrow \text{H}_3\text{PO}_4(\text{s})$

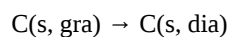
b.  $2\text{Na}(\text{s}) + 1/2\text{O}_2(\text{g}) \rightarrow \text{Na}_2\text{O}(\text{s})$

c.  $3\text{C}(\text{s}) + 1/2\text{O}_2(\text{g}) + 4\text{H}_2(\text{g}) \rightarrow \text{C}_3\text{H}_7\text{OH}(\ell)$

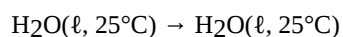
9. a.  $12\text{C(s)} + 11\text{H}_2(\text{g}) + 11/2\text{O}_2(\text{g}) \rightarrow \text{C}_{12}\text{H}_{22}\text{O}_{11}(\text{s})$   
 b.  $\text{Zn(s)} + \text{N}_2(\text{g}) + 3\text{O}_2(\text{g}) \rightarrow \text{Zn(NO}_3)_2$   
 c.  $\text{Al(s)} + 3/2\text{O}_2(\text{g}) + 3/2\text{H}_2(\text{g}) \rightarrow \text{Al(OH)}_3(\text{s})$
11. a.  $\text{MgCO}_3(\text{s}) \rightarrow \text{Mg(s)} + \text{C(s)} + 3/2\text{O}_2(\text{g})$   
 b.  $\text{Mg(s)} + 1/2\text{O}_2(\text{g}) \rightarrow \text{MgO(s)}$   
 c.  $\text{C(s)} + \text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g})$
13. a.  $2 \times [\text{CuCl(s)} \rightarrow \text{Cu(s)} + 1/2\text{Cl}_2(\text{g})]$   
 b.  $\text{Cu(s)} \rightarrow \text{Cu(s)}$   
 c.  $\text{Cu(s)} + \text{Cl}_2(\text{g}) \rightarrow \text{CuCl}_2(\text{s})$
15.  $\Delta H = -563.44 \text{ kJ}$
17.  $\Delta H = -546.7 \text{ kJ}$

### Additional Exercises

1. What is the work when 124 mL of gas contract to 72.0 mL under an external pressure of 822 torr?
2. What is the work when 2,345 mL of gas contract to 887 mL under an external pressure of 348 torr?
3. A 3.77 L volume of gas is exposed to an external pressure of 1.67 atm. As the gas contracts, 156 J of work are added to the gas. What is the final volume of the gas?
4. A 457 mL volume of gas contracts when 773 torr of external pressure act on it. If 27.4 J of work are added to the gas, what is its final volume?
5. What is the heat when 1,744 g of Hg increase in temperature by 334°C? Express your final answer in kJ.
6. What is the heat when 13.66 kg of Fe cool by 622°C? Express your final answer in kJ.
7. What is final temperature when a 45.6 g sample of Al at 87.3°C gains 188 J of heat?
8. What is final temperature when 967 g of Au at 557°C lose 559 J of heat?
9. Plants take  $\text{CO}_2$  and  $\text{H}_2\text{O}$  and make glucose ( $\text{C}_6\text{H}_{12}\text{O}_6$ ) and  $\text{O}_2$ . Write a balanced thermochemical equation for this process. Use data in Table 9.E. 1 of Section 7.7
10. Exercise 9 described the formation of glucose in plants, which take in  $\text{CO}_2$  and  $\text{H}_2\text{O}$  and give off  $\text{O}_2$ . Is this process exothermic or endothermic? If exothermic, where does the energy go? If endothermic, where does the energy come from?
11. The basic reaction in the refining of aluminum is to take  $\text{Al}_2\text{O}_3(\text{s})$  and turn it into  $\text{Al(s)}$  and  $\text{O}_2(\text{g})$ . Write the balanced thermochemical equation for this process. Use data in Table 9.E. 1 of Section 7.7
12. Is the enthalpy change of the reaction  
 $\text{H}_2\text{O}(\ell) \rightarrow \text{H}_2\text{O}(\text{g})$   
 zero or nonzero? Use data in Table 9.E. 1 of Section 7.7
13. What mass of  $\text{H}_2\text{O}$  can be heated from 22°C to 80°C in the combustion of 1 mol of  $\text{CH}_4$ ? You will need the balanced thermochemical equation for the combustion of  $\text{CH}_4$ . Use data in Table 9.E. 1 of Section 7.7
14. What mass of  $\text{H}_2\text{O}$  can be heated from 22°C to 80°C in the combustion of 1 mol of  $\text{C}_2\text{H}_6$ ? You will need the balanced thermochemical equation for the combustion of  $\text{C}_2\text{H}_6$ . Use data in Table 9.E. 1 of Section 7.7
15. What is the enthalpy change for the unknown reaction?  
 $\text{Pb(s)} + \text{Cl}_2(\text{g}) \rightarrow \text{PbCl}_2(\text{s}) \Delta H = -359 \text{ kJ}$   
 $\text{PbCl}_2(\text{s}) + \text{Cl}_2(\text{g}) \rightarrow \text{PbCl}_4(\ell) \Delta H = ?$   
 $\text{Pb(s)} + 2\text{Cl}_2(\text{g}) \rightarrow \text{PbCl}_4(\ell) \Delta H = -329 \text{ kJ}$
16. What is the enthalpy change for the unknown reaction?  
 $\text{P(s)} + 3/2\text{Br}_2(\ell) \rightarrow \text{PBr}_3(\ell) \Delta H = -185 \text{ kJ}$   
 $\text{PI}_3(\text{s}) \rightarrow \text{P(s)} + 3/2\text{I}_2(\text{s}) \Delta H = ?$   
 $\text{PI}_3(\text{s}) + 3/2\text{Br}_2(\ell) \rightarrow \text{PBr}_3(\ell) + 3/2\text{I}_2(\text{s}) \Delta H = -139 \text{ kJ}$
17. What is the  $\Delta H$  for this reaction? The label *gra* means graphite, and the label *dia* means diamond. What does your answer mean?



18. Without consulting any tables, determine the  $\Delta H$  for this reaction. Explain your answer.



#### Answers

1. 4.69 L
3. 80.97 kJ
5.  $91.9^\circ\text{C}$
7.  $6\text{CO}_2(\text{g}) + 6\text{H}_2\text{O}(\ell) \rightarrow \text{C}_6\text{H}_{12}\text{O}_6(\text{s}) + 6\text{O}_2(\text{g}) \Delta H = 2,799 \text{ kJ}$
9.  $2\text{Al}_2\text{O}_3(\text{s}) \rightarrow 4\text{Al}(\text{s}) + 3\text{O}_2(\text{g}) \Delta H = 3351.4 \text{ kJ}$
11. 3,668 g
13.  $\Delta H = 30 \text{ kJ}$
15. 5.70 J
17.  $\Delta H = 1.897 \text{ kJ}$ ; the reaction is endothermic.

---

This page titled [9.E: Energy and Chemistry \(Exercises\)](#) is shared under a [CC BY-NC-SA 3.0](#) license and was authored, remixed, and/or curated by [Theodore Chan](#) via [source content](#) that was edited to the style and standards of the LibreTexts platform.

- [7.E: Energy and Chemistry \(Exercises\)](#) by Anonymous is licensed [CC BY-NC-SA 3.0](#). Original source: <https://2012books.lardbucket.org/books/beginning-chemistry>.