

CHAPTER OVERVIEW

13: Solutions

We explored the general properties of gases, liquids, and [solids](#). Most of the discussion focused on pure substances containing a single kind of atom, molecule, or cation–anion pair. The substances we encounter in our daily lives, however, are usually mixtures rather than pure substances. Some are heterogeneous mixtures, which consist of at least two phases that are not uniformly dispersed on a microscopic scale; others are homogeneous mixtures, consisting of a single phase in which the components are uniformly distributed. Homogeneous mixtures are also called **solutions**; they include the air we breathe, the gas we use to cook and heat our homes, the water we drink, the gasoline or diesel fuel that powers engines, and the gold and silver jewelry we wear.

[13.1: Factors Affecting Solution Formation](#)

[13.2: Solubility and Molecular Structure](#)

[13.3: Units of Concentration](#)

[13.4: Effects of Temperature and Pressure on Solubility](#)

[13.5: Colligative Properties of Solutions](#)

[13.6: Aggregate Particles](#)

[13.7: Aggregate Particles in Aqueous Solution](#)

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