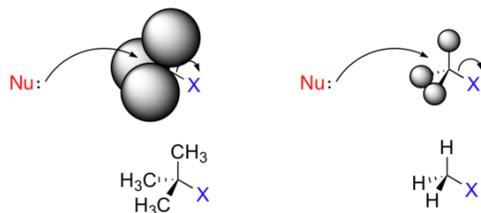


### 4.3.6: Electrophiles

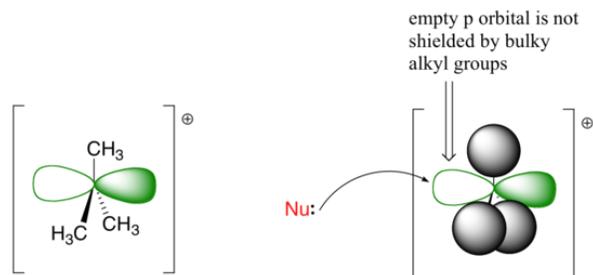
In the vast majority of the nucleophilic substitution reactions you will see in this and other organic chemistry texts, the electrophilic atom is a carbon which is bonded to an electronegative atom, usually oxygen, nitrogen, sulfur, or a halogen. The concept of electrophilicity is relatively simple: an electron-poor atom is an attractive target for something that is electron-rich, *i.e.* a nucleophile. However, we must also consider the effect of steric hindrance on electrophilicity. In addition, we must discuss how the nature of the electrophilic carbon, and more specifically the stability of a potential carbocationic intermediate, influences the  $S_N1$  vs.  $S_N2$  character of a nucleophilic substitution reaction.

Consider two hypothetical  $S_N2$  reactions: one in which the electrophile is a methyl carbon and another in which it is tertiary carbon.



Because the three substituents on the methyl carbon electrophile are tiny hydrogens, the nucleophile has a relatively clear path for backside attack. However, backside attack on the tertiary carbon is blocked by the bulkier methyl groups. Once again, steric hindrance - this time caused by bulky groups attached to the electrophile rather than to the nucleophile - hinders the progress of an associative nucleophilic ( $S_N2$ ) displacement.

The factors discussed in the above paragraph, however, do not prevent a sterically-hindered carbon from being a good electrophile - they only make it less likely to be attacked in a *concerted*  $S_N2$  reaction. Nucleophilic substitution reactions in which the electrophilic carbon is sterically hindered are more likely to occur by a two-step, dissociative ( $S_N1$ ) mechanism. This makes perfect sense from a geometric point of view: the limitations imposed by sterics are significant mainly in an  $S_N2$  displacement, when the electrophile being attacked is a  $sp^3$ -hybridized tetrahedral carbon with its relatively 'tight' angles of  $109.4^\circ$ . Remember that in an  $S_N1$  mechanism, the nucleophile attacks an  $sp^2$ -hybridized carbocation intermediate, which has trigonal planar geometry with 'open'  $120^\circ$  angles.



With this open geometry, the empty p orbital of the electrophilic carbocation is no longer significantly shielded from the approaching nucleophile by the bulky alkyl groups. A carbocation is a very potent electrophile, and the nucleophilic step occurs very rapidly compared to the first (ionization) step.

This page titled [4.3.6: Electrophiles](#) is shared under a [not declared](#) license and was authored, remixed, and/or curated by [Layne Morsch](#).

- [Electrophiles](#) is licensed [CC BY-NC-SA 4.0](#).