

Detailed Licensing

Overview

Title: Intermediate Physical Organic (Morsch)

Webpages: 115

Applicable Restrictions: Noncommercial

All licenses found:

- **Undeclared:** 74.8% (86 pages)
- **CC BY-NC-SA 4.0:** 11.3% (13 pages)
- **CC BY-NC 3.0:** 5.2% (6 pages)
- **CC BY-SA 4.0:** 3.5% (4 pages)
- **CC BY-NC 4.0:** 2.6% (3 pages)
- **CC BY 4.0:** 1.7% (2 pages)
- **CC BY-NC-SA 3.0:** 0.9% (1 page)

By Page

- Intermediate Physical Organic (Morsch) - *Undeclared*
 - Front Matter - *Undeclared*
 - [TitlePage](#) - *Undeclared*
 - [InfoPage](#) - *Undeclared*
 - [Table of Contents](#) - *Undeclared*
 - [Licensing](#) - *Undeclared*
 - 1: Models of Chemical Bonding - *Undeclared*
 - 1.1: Atomic and Molecular Properties - *Undeclared*
 - 1.1.1: Atomic and Ionic Radius - *CC BY-NC 4.0*
 - 1.1.2: Bond Lengths and Double-Bond Character - *Undeclared*
 - 1.1.3: Dipole Moments - *CC BY 4.0*
 - 1.1.4: Electronegativity - *CC BY-NC 4.0*
 - 1.1.5: Molecular Polarity - *Undeclared*
 - 1.2: Intermolecular Forces - Introduction and London Dispersion - *Undeclared*
 - 1.2.1: Specific Interactions - *Undeclared*
 - 1.2.2: Hydrogen Bonding - *Undeclared*
 - 1.2.3: Dipole-Dipole Interactions - *Undeclared*
 - 1.2.4: Polarizability - *CC BY 4.0*
 - 1.3: Overview of Valence Bond Theory - *Undeclared*
 - 1.3.1: The Born-Oppenheimer Approximation Simplifies the Schrödinger Equation for Molecules - *Undeclared*
 - 1.3.2: Covalent Bonding and Orbital Overlap - *Undeclared*
 - 1.3.3: Hybrid Orbitals - *Undeclared*
 - 1.3.4: Delocalization of Electrons - *Undeclared*
 - 1.3.5: Lennard-Jones Potential - *Undeclared*
 - 1.4: Molecular orbital theory- conjugation and aromaticity - *CC BY-NC-SA 4.0*
 - 1.4.1: The H_2^+ Prototypical Species - *Undeclared*
 - 1.4.2: Bonding and Antibonding Orbitals - *Undeclared*
 - 1.4.3: A Simple Molecular-Orbital Treatment of H_2 Places Both Electrons in a Bonding Orbital - *Undeclared*
 - 1.4.4: Molecular Orbitals Can Be Ordered According to Their Energies - *Undeclared*
 - 1.4.5: Molecular-Orbital Theory Also Applies to Heteronuclear Diatomic Molecules - *CC BY-NC-SA 3.0*
 - 1.4.6: Butadiene is Stabilized by a Delocalization Energy - *CC BY-SA 4.0*
 - 1.4.6.1: Application of the MO Method to 1,3-Butadiene - *Undeclared*
 - 1.4.7: Molecular Orbitals of Conjugated Pi Systems - *Undeclared*
 - 1.4.8: Benzene and Aromaticity - *Undeclared*
 - 1.4.8.1: Hückel's $4n + 2$ Rule - *Undeclared*
 - 1.5: Comparison of the Resonance and Molecular-Orbital Methods - *Undeclared*
 - 1.5.1: Which Is Better- MO or VB? - *Undeclared*
- 2: Reaction Kinetics - *Undeclared*
 - 2.1: Introduction to Reaction Kinetics - *CC BY-NC 3.0*
 - 2.1.1: First-Order Reactions - *Undeclared*
 - 2.1.2: Half-lives - *Undeclared*
 - 2.1.3: Reaction Rate - *Undeclared*
 - 2.1.3.1: The "Speed" of a Chemical Reaction - *Undeclared*
 - 2.1.3.2: The Rate of a Chemical Reaction - *Undeclared*

- 8.1.4: A4. Intramolecular Catalysis - *CC BY-NC-SA 4.0*
 - 8.1.5: A5. Transition State Stabilization - *CC BY-NC-SA 4.0*
 - 8.1.6: A6. Links and References - *CC BY-NC-SA 4.0*
- Back Matter - *Undeclared*
 - Index - *Undeclared*
 - Glossary - *Undeclared*
 - Detailed Licensing - *Undeclared*