

11.E: Quantum Mechanics and Atomic Structure (Exercises)

11.1: The Wave Theory of Light

Q11.1

What is the wavelength associated with a photon of a light with the energy is $3.6 \times 10^{-19} \text{ J}$?

S11.1

$$E = hv = \frac{hc}{\lambda}$$

$$3.6 \times 10^{-19} \text{ J} = \frac{(6.626 \times 10^{-34} \text{ J s})(3 \times 10^8 \text{ m s}^{-1})}{\lambda}$$

$$\lambda = 0.55 \times 10^{-7} \text{ m} = 550 \text{ nm}$$

Q11.2

Calculate the energy of a photon of a light with the frequency is $6.5 \times 10^{14} \text{ s}^{-1}$?

S11.2

$$E = hv$$

$$E = (6.626 \times 10^{-34} \text{ J s})(6.5 \times 10^{14} \text{ s}^{-1})$$

$$E = 4.3 \times 10^{-19} \text{ J}$$

Q11.3a

What are the wavelength of the light emitted by atomic hydrogen when the electron move from $n = 5$ to $n = 3$ orbit in the Bohr model?

S11.3a

Bohr model:

$$\frac{1}{\lambda} = R_H \left(\frac{1}{n_a^2} - \frac{1}{n_b^2} \right)$$

$$\frac{1}{\lambda} = (1.09678 \times 10^{-2} \text{ nm}^{-1}) \left(\frac{1}{9} - \frac{1}{25} \right)$$

$$\lambda = 1,282 \text{ nm}$$

Q11.3b

Calculate the different energy between $n = 3$ and $n = 4$ of atomic hydrogen based on Bohr orbits

S11.3b

$$\Delta E = -hcR_H \left(\frac{1}{3^2} - \frac{1}{4^2} \right)$$

$$\Delta E = -6.626 \times 10^{-34} \text{ J s} (2.9979 \times 10^8 \text{ m/s})(109,737 \text{ cm}^{-1}) \left(\frac{1}{3^2} - \frac{1}{4^2} \right)$$

$$\Delta E = 1.8 \times 10^{-21} \text{ J}$$

Q11.4

Calculate the wavenumber of the wavelength of the light emitted from the $n = 8$ to $n = 6$ transition.

S11.4

This is a simple application of the Rydberg equation

$$\tilde{\nu} R_H \left(\frac{1}{n_f^2} - \frac{1}{n_i^2} \right)$$

with $n_i = 6$ and $n_f = 8$.

$$\tilde{\nu} = 109,737 \text{ cm}^{-1} \left(\frac{1}{6^2} - \frac{1}{8^2} \right) = 1333.61 \text{ cm}^{-1}$$

Q11.5

Calculate the wavelength associated with a 42 g baseball with speed of 80 m/s.

S11.5

$$\lambda = \frac{h}{p} = \frac{h}{mv}$$

$$\lambda = \frac{6.626 \times 10^{-34} \text{ Js}}{(0.042 \text{ kg})(80 \text{ m/s})} = 1.97 \times 10^{-34} \text{ m}$$

This is a very small wavelength as expected since a ball behaves rather classically (i.e., non-quantum).

Q11.6

Calculate the energy of a 530-nm photon.

S11.6

$$p = \frac{E}{c}$$

$$\lambda = \frac{h}{p}$$

$$\lambda = \frac{hc}{E}$$

$$E = \frac{hc}{\lambda} = \frac{(6.626 \times 10^{-34})(3 \times 10^8)}{530 \times 10^{-9}} = 3.75 \times 10^{-19} \text{ J}$$

11.3: The Photoelectric Effect

Q11.7

Describe Planck's three experimental observations that explain the photoelectric effect.

S11.7

"When light of a certain frequency shines on a clean metal surface, electrons are ejected from the metal. Experimentally, it is found that:

1. The number of electrons ejected is proportional to the intensity of light
2. The kinetic energy of the ejected electrons is proportional to the frequency of the ejected light
3. No electrons can be ejected if the frequency of the light is lower than a certain value, called the threshold frequency (ν_0).

Q11.8a

What is the difference between classical and quantum mechanics? Provide the equation relating the energy of emitted radiation to frequency.

S11.8a

Classical mechanics predicts that the radiant energy produced by oscillating objects is continuous. Quantum mechanics predicts that their energy can be viewed as existing in discrete levels. The equation is

$$E = h\nu$$

. Energy is emitted in discrete multiples of $h\nu$, where h = Planck's constant.

Q11.8b

Do you expect the *first* ionization of an atom to be an endothermic process or an exothermic process?

S11.8b

It will be endothermic because energy is needed to ionize the atom.

11.2: Planck's Quantum Theory

Q11.9a

Calculate the wavelength of light with an energy of $5.22 \times 10^{-19} \text{ J}$.

S11.9a

Rearrange the equation

$$E = \frac{hc}{\lambda}$$

so that the unknown value is the wavelength of light

$$\lambda = \frac{hc}{E}$$

$$\lambda = \frac{(6.626 \times 10^{-34} \text{ Js})(3.0 \times 10^8 \text{ m/s})}{5.22 \times 10^{-19} \text{ J}} = 3.81 \times 10^{-7} \text{ m} = 381 \text{ nm}$$

Q11.9b

You realize after class that you forgot to write down the final wavelength of your proton. You do, however, remember that the potential difference was 114 V. From this information, figure out the final wavelength:

S11.9b

$$\frac{1}{2}mv^2 = eV$$

$e = 1.6 \times 10^{-19} \text{ C}$ This is the charge of the electron

$m = 9.11 \times 10^{-31} \text{ kg}$

$$v = \sqrt{\frac{2eV}{m}}$$

$$\sqrt{\frac{2(1.6 \cdot 10^{-19})(114)}{1.67 \cdot 10^{-31}}}$$

$$= 2.0 \times 10^9 \text{ m/s}$$

$$\lambda = \frac{h}{mv}$$

$$\lambda = \frac{6.626 \cdot 10^{-34}}{(9.11 \cdot 10^{-31})(2 \cdot 10^9)}$$

$$\lambda = 3.64 \cdot 10^{-13} \text{ m}$$

Q11.10a

given the uncertainty for position of an electron circling an atom is 0.5 \AA . Find uncertainty in its velocity?

S11.10a

$$\Delta p \geq h / 4\pi\Delta x = 6.626 \times 10^{-34} \text{ Js} / 4\pi(0.5 \times 10^{-10} \text{ m}) = 1.31105 \times 10^{-24} \text{ kg m s}^{-1}$$

$$\Delta v = \Delta p / m \geq 1.3 \times 10^{-24} \text{ kg m s}^{-1} / 9.109 \times 10^{-31} \text{ kg} = 1 \times 10^6 \text{ m s}^{-1}$$

Q11.10

If the uncertainty of measuring the position of an electron is 2.0 \AA , what is the uncertainty of simultaneously measuring its velocity?

Hint: What formula deals with uncertainty of measurements?

11.5: de Broglie's Postulate

11.6: The Heisenberg Uncertainty Principle

11.7: The Schrödinger Wave Equation

Q11.11

A typical mass for a horse is 510 kg, and a typical galloping speed is 22 kilometers per hour. Use these values to answer the following questions.

- What is the momentum of a galloping horse? What is its wavelength?
- If a galloping horse's velocity and position are simultaneously measured, and the velocity is measured to within $\pm 1.0\%$, what is the uncertainty of its position?
- Suppose Planck's constant was actually 0.01 J s. How would that change your answers to (a) and (b)? Which values would be unchanged?

Hints:

- de Broglie's postulate deals with the wave-like properties of particles.*
- Heisenberg's uncertainty principle deals with uncertainty of simultaneous measurements.*

S11.11

(a)

- Use the relationship between velocity and momentum to find the momentum.

$$p = mv$$

$$p = 510 \text{ kg} \times 22 \frac{\text{km}}{\text{hr}} \times \frac{1000 \text{ m}}{\text{km}} \times \frac{\text{hr}}{3600 \text{ s}}$$

$$p = 3.1 \times 10^3 \frac{\text{kg m}}{\text{s}}$$

- Find the de Broglie wavelength.

$$\lambda = \frac{h}{p}$$

$$\lambda = \frac{6.626 \times 10^{-34} \text{ J s}}{3.1 \times 10^3 \frac{\text{kg m}}{\text{s}}} \times \frac{\frac{\text{kg m}^2}{\text{s}^2}}{\text{J}}$$

$$\lambda = 2.1 \times 10^{-37} \text{ m}$$

(b)

- Find the uncertainty of momentum from the uncertainty of velocity.

$$\Delta p = m \Delta v$$

$$\Delta v = 0.01 \times v$$

$$\Delta p = 510 \text{ kg} \times 0.01 \times 22 \frac{\text{km}}{\text{hr}} \times \frac{1000 \text{ m}}{\text{km}} \times \frac{\text{hr}}{3600 \text{ s}}$$

$$\Delta p = 31 \frac{\text{kg m}}{\text{s}}$$

- Use Heisenberg's uncertainty principle to find the uncertainty of position.

$$\Delta x \Delta p \geq \frac{h}{4\pi}$$

$$\Delta x \geq \frac{h}{4\pi \Delta p}$$

$$\Delta x \geq \frac{6.626 \times 10^{-34} \text{ J s}}{4\pi \times 3.1 \times 10^3 \frac{\text{kg m}}{\text{s}}} \times \frac{\frac{\text{kg m}^2}{\text{s}^2}}{\text{J}}$$

$$\Delta x \geq 1.7 \times 10^{-36} \text{ m}$$

(c)

The calculations that involve Planck's constant (h) will change. This much larger value will make the wave-like properties of the horse more important; its wavelength and uncertainty of position will increase dramatically:

$$\lambda' = \frac{h'}{p} = \frac{0.010 \text{ J s}}{3.1 \times 10^3 \frac{\text{kg m}}{\text{s}} \times \frac{\text{kg m}^2}{\text{s}^2}}$$

$$\lambda' = 3.2 \times 10^{-6} \text{ m} = 3200 \text{ nm}$$

$$\Delta x' \geq \frac{h'}{4\pi \Delta p} = \frac{0.010 \text{ J s}}{4\pi \times 3.1 \times 10^3 \frac{\text{kg m}}{\text{s}}} \times \frac{\frac{\text{kg m}^2}{\text{s}^2}}{\text{J}}$$

$$\Delta x' \geq 2.6 \times 10^{-5} \text{ m} = 26 \mu\text{m}$$

The momentum would not change.

Q11.12a

how fast must a person weighing 42 kg move through a door 0.5 m wide in order to be diffracted?

S11.12a

$$p = h/\lambda = 6.626 \times 10^{-34} \text{ J s} / 0.5$$

$$= 3.313 \times 10^{-34} \text{ kg m s}^{-1}$$

$$V = p/m = 3.313 \times 10^{-34} / 42 \text{ kg} = 7.88 \times 10^{-36} \text{ ms}^{-1}$$

Q11.12b

When a particle passes through a slit, diffraction occurs if the particle's wavelength is on the same order of magnitude as the width of the slit. At approximately what velocity would a 1670 kg car have to move through the Lincoln Tunnel (6.6 m width) for diffraction to occur? *Hint: How can you find the wavelength of the car?*

S11.12b

1. Use de Broglie's equation to relate velocity to wavelength; solve for velocity.

$$\lambda = \frac{h}{mv}$$

$$v = \frac{h}{m\lambda}$$

2. Set the wavelength equal to the tunnel width and solve.

$$v = \frac{6.626 \times 10^{-34} \text{ J s}}{1670 \text{ kg} \times 6.6 \text{ m}} \times \frac{\frac{\text{kg m}^2}{\text{s}^2}}{\text{J}}$$

Answer: the car would have to be driving very slowly.

$$v = 6.0 \times 10^{-38} \text{ m/s}$$

11.4: Bohr's Theory of the Hydrogen Emission Spectrum

Q11.13a

The Paschen emission spectrum is the collection of spectral lines emitted by H atoms, where the final state is $n = 3$. What are the shortest and longest wavelengths (in nm) of Paschen spectral lines? *Hints: The minimum value for n_i is 4, the maximum value for n_i is infinity and the Rydberg formula describes the emission spectra of hydrogen*

Q11.13b

Spectral lines of Lyman and Balmer series do not overlap. Verify this statement.

S11.13b

A photon has a specific energy according to the energy between the excited and ground state.

$$E = (1/n_f^2 - 1/n_i^2), A = hv$$

The Lyman series is when n_f is the ground state or $n_f = 1$ and Balmer series is when $n_f = 2$

$n_f = 1$ the value, n_i from 2 to infinity with the values between 1 and $1/2$. For the Balmer series n_i is from 3 to infinity the range between $1/4$ and $5/36$. Thus, these intervals do not overlap \rightarrow the energies do not overlap \rightarrow the lines in the series can't overlap

Q11.14

Assume the Rydberg constant for He^+ is $9.72 \times 10^{-18} \text{ nm}$, calculate the wavelength of He^+ ions from $n=3$ to $n=2$.

S11.14

$$\frac{1}{\lambda} = R_H \left| \frac{1}{n_i^2} - \frac{1}{n_f^2} \right|$$

$$\frac{1}{\lambda} = 9.72 \times 10^{-18} \left| \frac{1}{3^2} - \frac{1}{2^2} \right|$$

$$\lambda = 7.41 \times 10^{17} \text{ nm}$$

Q11.15

Derive from the following equation to solve for wavelength.

$$\Delta E = hcR_H \left(\frac{1}{n_i^2} - \frac{1}{n_f^2} \right)$$

S11.15

$$\Delta E = R_H \left| \frac{1}{n_i^2} - \frac{1}{n_f^2} \right|$$

$$\frac{v}{c} = \Delta E = R_H \left| \frac{1}{n_i^2} - \frac{1}{n_f^2} \right|$$

$$\frac{1}{\lambda} = \frac{v}{c} = \Delta E = R_H \left| \frac{1}{n_i^2} - \frac{1}{n_f^2} \right|$$

$$\lambda = \frac{1}{R_H \left| \frac{1}{n_i^2} - \frac{1}{n_f^2} \right|}$$

Q11.16

Photosynthesis occurs when light excites the electrons in the chloroplasts of leaves. If you flash a light of 500 nm at the plant and assuming your flashlight emits 4.5×10^{20} numbers of photons, how much energy in joules was that one flash?

Q11.50

What is an antiatom? How is it different from a regular hydrogen? What will happen if these two collided with one another?

S11.50

An antiatom is a hydrogen atom with reversed electrical charges compared to that of a hydrogen atom. The proton is called an antiproton and it bears a negative charge rather than a positive charge. The electron, or antielectron, has a positive charge rather than a negative charge. As a result, an antiatom and regular hydrogen would obliterate one another should they collide. Subsequent energy would then be released.

Q11.51

A scientist determined the kinetic energy released electrons of cesium metal through a photoelectric experiment. Determine h and the work function for cesium graphically from the following results:

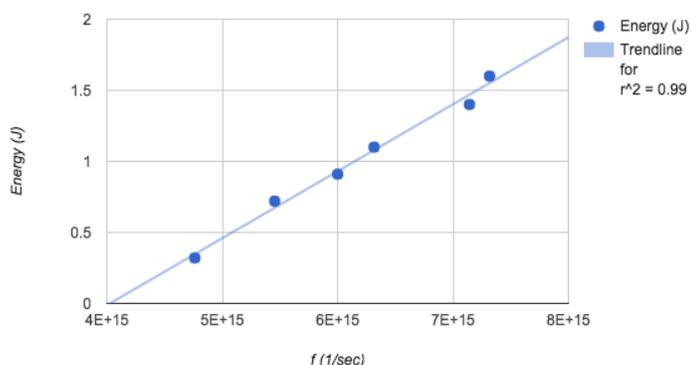
λ/nm	410	420	475	500	550	630
V/volt	1.6	1.4	1.1	0.91	0.72	0.37

S11.51

First, convert wavelength to frequency and volt to energy in J.

f (1/sec)	Energy (J)
7.31707×10^{15}	1.6
7.14286×10^{15}	1.4
6.3157×10^{15}	1.1
6×10^{15}	0.91
5.45455×10^{15}	0.72
4.7619×10^{15}	0.32

Then graph the values and determine the linear equation.



The linear equation is

$$y = 4.7 \times 10^{-16} f - 1.89$$

Use the equation below:

$$hf - \phi = E_k \text{ (max kinetic energy)}$$

$$h = \text{slope}$$

$\phi =$ y-intercept

$$h = 4.7 \times 10^{-16}$$

$$\phi = 1.89$$

Q11.52

Calculate the de Broglie wavelength of a Cl_2 molecule at 300 K.

S11.52

$$v_{rms} = \sqrt{\frac{3RT}{M}}$$

$$= ((3)(8.314)(300)) / (70.8 \times 10^{-3} \text{ kg}) = 325 \text{ m/s}$$

$$\lambda = \frac{h}{mv_{rms}}$$

$$= (6.626 \times 10^{-34} \text{ J*s}) / [(70.8 \text{ amu})(1.66 \times 10^{-27} \text{ kg/amu})(325 \text{ m/s})] = 1.73 \times 10^{-11} \text{ m}$$

Q11.53

Consider a balloon with a diameter of $2.5 \times 10^{-5} \text{ m}$. What is the uncertainty of the velocity of an oxygen molecule that is trapped inside.

S11.53

Use the direct application of the uncertainty principle:

$$\Delta x \Delta p \geq \frac{h}{4\pi}$$

Let's consider the molecule has an uncertainty that is \pm radius of the balloon.

$$\Delta x = 1.3 \times 10^{-5} \text{ m}$$

The uncertainty of the momentum of the molecule can be estimated via the uncertainty principle.

$$\Delta p = \frac{h}{4\pi \Delta x} = \frac{(6.626 \times 10^{-34} \text{ J*s})}{(4\pi)(1.3 \times 10^{-5} \text{ m})} = 4.05 \times 10^{-30} \text{ kg m s}^{-1}$$

This can be converted to uncertainty in velocity via

$$p = mv$$

or

$$\Delta p = m_e \Delta v$$

with the electron mass m_e equal to $9.109 \times 10^{-31} \text{ kg}$, since the mass of the molecule is **not** uncertain.

$$\Delta v = \frac{\Delta p}{m_e} = \frac{4.05 \times 10^{-30}}{9.109 \times 10^{-31} \text{ kg}} = 4.5 \text{ m/s}$$

11.8: Particle in a One-Dimensional Box

Q11.55

What conditions must be satisfied for the Schrodinger's equation for ψ to be an acceptable wavefunction. What are some examples of unacceptable wave functions?

Q11.56

The equation for calculating the energies of the electron in a hydrogen atom or a hydrogenlike ion is given by $E_n = -(2.18 \times 10^{-18} \text{ J})Z^2(1/n^2)$, where Z is the atomic number of the element. One way to modify this equation for many-electron atoms is to replace Z with $(Z - [\sigma])$, where $[\sigma]$ is a positive dimensionless quantity called the shielding constant. Consider the lithium atom as an example. The physical significance of σ that it represents the extent of shielding that the two 1s electrons exert on each other.

Thus the quantity $(Z - [\sigma])$ is appropriated called the "effective nuclear charge." Use the first ionization energy of lithium to calculate the value of $[\sigma]$.

Q11.57

Calculate the wavelength and frequency of an emitted gamma particle that has energy of $3.11 \times 10^{12} \text{ J mol}^{-1}$.

11.9: Quantum-Mechanical Tunneling

11.10: The Schrödinger Wave Equation for the Hydrogen Atom

11.11: Many-Electron Atoms and the Periodic Table

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