

CHAPTER OVERVIEW

6: Multiple Component Phase Equilibrium

- 6.1: Thermodynamics of Mixing
- 6.2: Partial Molar Volume
- 6.3: Chemical Potential
- 6.4: Non-ideality in Gases - Fugacity
- 6.5: A Mixture is a Combination of Two or More Substances
- 6.6: The Gibbs-Duhem Equation Relates Chemical Potential and Composition at Equilibrium
- 6.7: Chemical Potential of Each Component Has the Same Value in Each Phase in Which the Component Appears
- 6.8: Colligative Properties
- 6.9: Osmotic Pressure can Determine Molecular Masses
- 6.10: Raoult's Law and Phase Diagrams of Ideal Solutions
- 6.11: Fractional Distillation of Ideal Mixtures
- 6.12: Most Solutions are Not Ideal
- 6.13: Phase Diagrams of Non-Ideal Solutions
- 6.14: Fractional Distillation of Non-ideal Mixtures (Azeotropes)
- 6.15: Activity

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