

CHAPTER OVERVIEW

7: Chemical Equilibrium

- 7.1: Equilibrium Results when Gibbs Energy is Minimized
- 7.2: Reaction Quotient and Equilibrium Constant
- 7.3: An Equilibrium Constant is a Function of Temperature Only
- 7.4: Pressure Dependence of K_p - Le Châtelier's Principle
- 7.5: Degree of Dissociation
- 7.6: The Dumas Bulb Method for Measuring Decomposition Equilibrium
- 7.7: Gibbs Energy of a Reaction vs. Extent of Reaction is a Minimum at Equilibrium
- 7.8: The Sign of ΔG and not ΔG° Determines the Direction of Reaction Spontaneity
- 7.9: Reaction Quotient and Equilibrium Constant Ratio Determines Reaction Direction
- 7.10: The van 't Hoff Equation

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