

## 15.6: POLYCYCLIC AROMATIC COMPOUNDS

### OBJECTIVES

After completing this section, you should be able to draw the resonance contributors for polycyclic aromatic compounds, such as naphthalene, anthracene, etc.

### KEY TERMS

Make certain that you can define, and use in context, the key term below.

- polycyclic aromatic compounds

### STUDY NOTES

As their name indicates, *polycyclic aromatic hydrocarbons* are aromatic hydrocarbons which contain more than one benzenoid (i.e., benzene-like) ring. This section deals only with those compounds in which the benzenoid rings are fused together; in other words, compounds in which at least one carbon-carbon bond is common to two aromatic rings. Another type of polycyclic aromatic hydrocarbon contains two or more benzenoid rings joined by a carbon-carbon single bond. The simplest compound of this type is biphenyl, the compound from which PCBs (polychlorinated biphenyls) are derived.

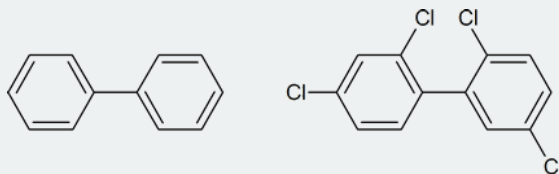


Figure 15.3: Structures of biphenyl and a typical PCB

## POLYCYCLIC AROMATIC COMPOUNDS

Hückel's  $4n + 2$  rule for aromaticity does not only apply to mono-cyclic compounds. Benzene rings may be joined together (fused) to give larger polycyclic aromatic compounds. A few examples are drawn below, together with the approved numbering scheme for substituted derivatives. The peripheral carbon atoms (numbered in all but the last three examples) are all bonded to hydrogen atoms. Unlike benzene, all the C-C bond lengths in these fused ring aromatics are not the same, and there is some localization of the pi-electrons. The six benzene rings in coronene are fused in a planar ring; whereas the six rings in hexahelicene are not joined in a larger ring, but assume a helical turn, due to the crowding together of the terminal ring atoms (in the structure below, note that the top right and center right rings are not attached to one another). This helical configuration renders the hexahelicene molecule chiral, and it has been resolved into stable enantiomers.

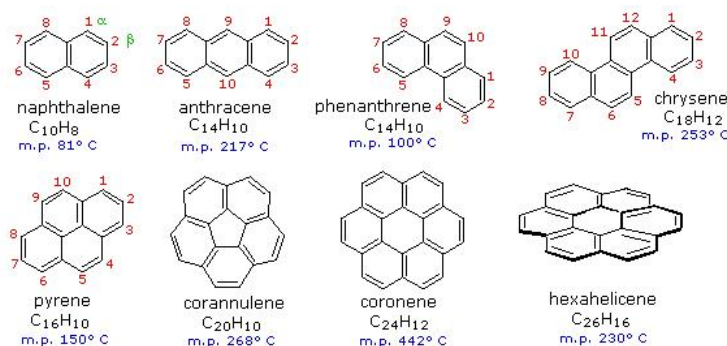
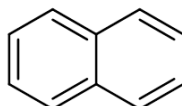
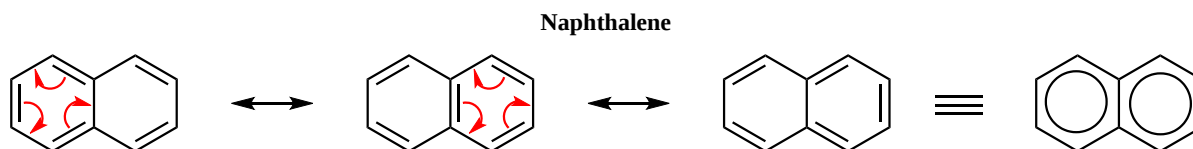


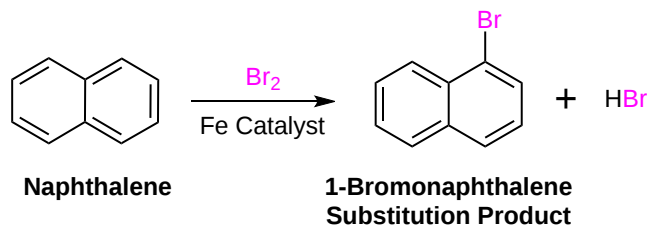
Figure 2: Examples of Polycyclic Aromatic Hydrocarbons (PAHs).

Probably the most well known polycyclic aromatic compound which only contains carbon is naphthalene ( $C_{10}H_8$ ). Naphthalene shares many of the same characteristics as benzene. Naphthalene is cyclic and known to be flat. Each carbon in naphthalene is  $sp^2$  hybridized so it is completely conjugated. The true structure of naphthalene is a combination of three resonance hybrids.

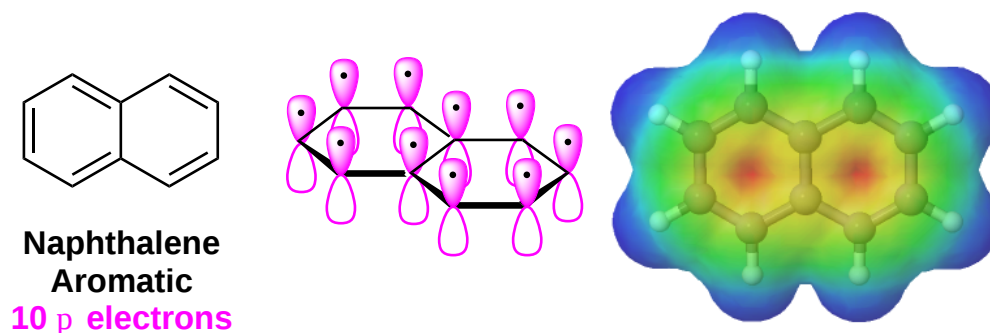




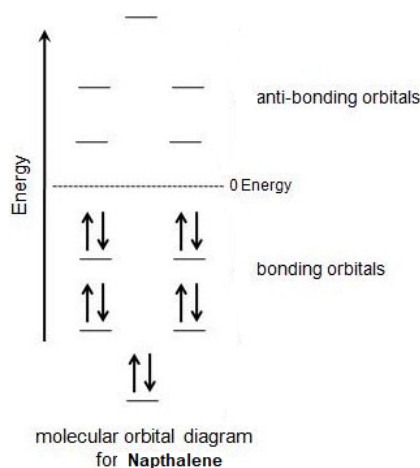
Heat of hydrogenation experiments with naphthalene shows an unusual "aromatic" stabilization energy. Also, naphthalene prefers to react with electrophiles to give substitution products rather than the typical double bond addition products.



Every carbon in naphthalene is  $sp^2$  hybridized so there is conjugated p orbital overlap over the entire ring system. The electrostatic potential map shows that pi electrons in naphthalene are evenly distributed making each carbon equivalent.

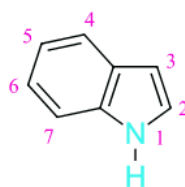


Lastly, naphthalene has 10 pi electrons which fulfills Hückel's rule. The importance of the  $4n + 2$  rule can be seen when considering the molecular orbital diagram of naphthalene. Naphthalene has 10 p orbitals which is 4 more than benzene. In the molecular orbital diagram of naphthalene, the 4 additional p orbitals become 2 bonding orbitals and two anti-bonding orbitals. The two additional bonding orbitals require 4 additional pi electrons to be filled so naphthalene needs 10 pi electrons in total to fill all of the bonding molecular orbitals.

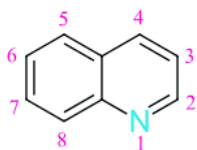


## POLYCYCLIC AROMATIC HETEROCYCLES

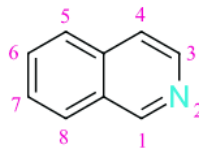
There is a wide variety of polycyclic aromatic heterocycles, many of which contain nitrogen, oxygen, or sulfur. Some of the biologically important structures, are the nitrogen containing polycyclic aromatic heterocycles indole, quinoline, isoquinoline, and purine which are all polycyclic aromatic heterocycles commonly found in nature. Notice that these compound all have 10 pi electrons.



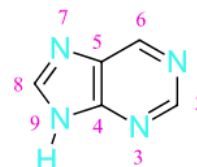
Indole



Quinoline

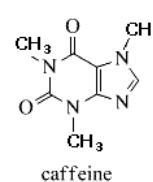
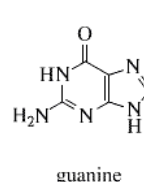
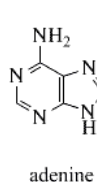
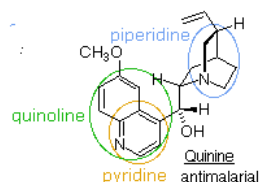
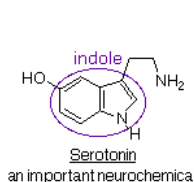


Isoquinoline



Purine

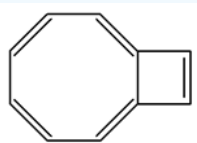
Indole, quinoline, and isoquinoline all contain a heterocyclic ring fused to benzene. Purine is made up of two heterocyclic rings, imidazole and pyrimidine, fused together. Quinoline is found in the antimalarial drug quinine. Indole is found in the neurotransmitter serotonin. The purine ring structure is found in adenine and guanine, two important parts of DNA and in the stimulant caffeine.



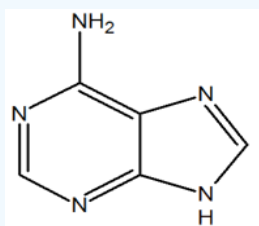
## EXERCISES

### ? EXERCISE 15.6.1

1) The following molecule is an isomer of naphthalene. Is it aromatic? Draw a resonance structure for it.

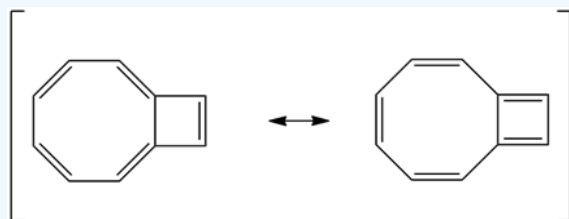


2) The following molecule is adenine. It has a purine core. Of the nitrogen in the core, how many electrons are donated into the pi system?

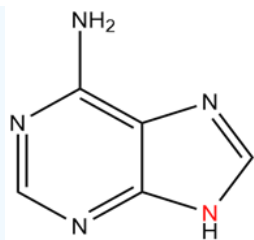


### Answer

1) It has 10 pi electrons which follows the  $4n+2$  rule making it aromatic.

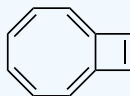


2) There is only one nitrogen of the core that contributes a set of lone pair electrons as 2 pi electrons (in red). All of the other nitrogens are  $sp^2$  hybridized and contribute 1 pi electron each. In total, the core is aromatic with 10 electrons in the pi-system.



### ? EXERCISE 15.6.2

This is an isomer of naphthalene. Is it aromatic? Draw a resonance structure for it.

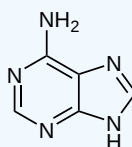


#### Answer

Yes, it is aromatic.  $4n+2$  pi-electrons.

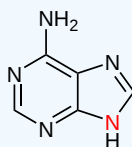
### ? EXERCISE 15.6.3

The following molecule is adenine. It has a purine core. Of the nitrogen in the core, how many electrons are donated into the pi system?



#### Answer

There is only one nitrogen of the core that contributes to the pi-system (in red). With this one lone pair the core is aromatic with 10 electrons in the pi-system.



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