

1.3: Multi-Electron Atoms

Learning objectives for this unit are to:

- Explain how and why Z affects the size and energy of atomic orbitals
 - Explain the meaning of effective nuclear charge (Z^*) in terms of shielding
 - Rationalize electron configurations in terms of the Pauli Exclusion Principle, the Aufbau principle and Hund's rule (Pc and Pe)
 - Write electronic electron configurations for atoms, ions, and excited states
 - Use Slater's rules to calculate Z^*
 - Understand how Z^* affects the relative energies of atomic orbitals
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