

11.2: Ions in Solution (Electrolytes)

In [Binary Ionic Compounds and Their Properties](#) we point out that when an ionic compound dissolves in water, the positive and negative ions originally present in the crystal lattice persist in solution. Their ability to move nearly independently through the solution permits them to carry positive or negative electrical charges from one place to another. Hence the solution conducts an electrical current.

Electrolytes

Substances whose solutions conduct electricity are called electrolytes. All soluble ionic compounds are strong electrolytes. They conduct very well because they provide a plentiful supply of ions in solution. Some polar covalent compounds are also strong electrolytes. Common examples are HCl, HBr, HI and H_2SO_4 , all of which react with H_2O to form large concentrations of ions. A solution of HCl, for example, conducts even better than one of NaCl having the same concentration.

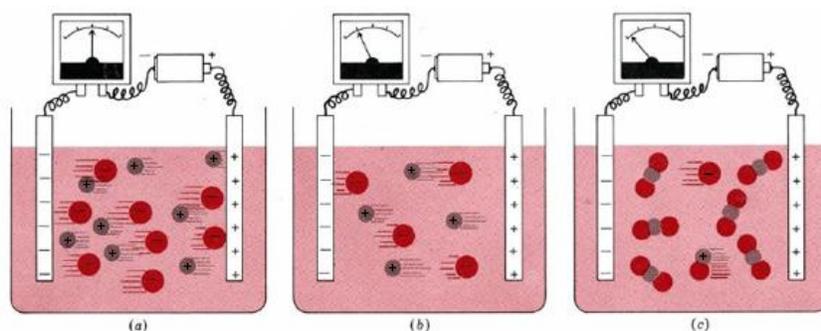


Figure 11.2.1: The conductivity of electrolyte solutions: (a) 0.1 M NaCl (b) 0.05 M NaCl (c) 0.1 M HgCl_2 . An electrolyte solution conducts electricity because of the movement of ions in the solution (see above). The larger the concentration of ions, the better the solution conducts. Weak electrolytes, such as HgCl_2 , conduct badly because they produce few ions when dissolved (low concentration of ions) and exist mainly in the form of molecules.

The effect of the concentration of ions on the electrical current flowing through a solution is illustrated in Figure 11.2.1. Part *a* of the figure shows what happens when a battery is connected through an electrical meter to two inert metal strips (**electrodes**) dipping in ethanol. Each cubic decimeter of such a solution contains 0.10 mol NaCl (that is, 0.10 mol Na^+ and 0.10 mol Cl^-). An electrical current is carried through the solution both by the Na^+ ions moving toward the negative electrode and by the Cl^- ions which are attracted toward the positive electrode. The dial on the meter indicates the quantity of current.

Figure 1*b* shows that if we replace the 0.10-*M* NaCl solution with a 0.05-*M* NaCl solution, the meter reading falls to about one-half its former value. Halving the concentration of NaCl halves the number of ions between the electrodes, and half as many ions can only carry half as much electrical charge. Therefore the current is half as great. Because it responds in such a direct way to the concentration of ions, conductivity of electrical current is a useful tool in the study of solutions.

Conductivity measurements reveal that most covalent compounds, if they dissolve in water at all, retain their original molecular structures. Neutral molecules cannot carry electrical charges through the solution, and so no current flows. A substance whose aqueous solution conducts no better than water itself is called a **nonelectrolyte**. Some examples are oxygen, O_2 , ethanol, $\text{C}_2\text{H}_5\text{OH}$, and sugar, $\text{C}_{12}\text{H}_{22}\text{O}_{11}$.

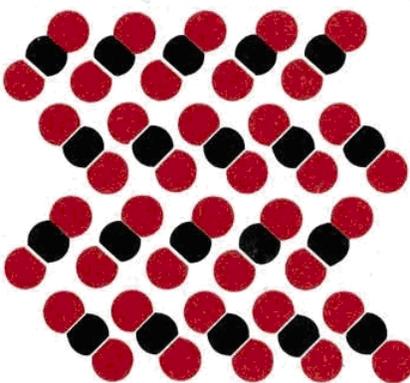
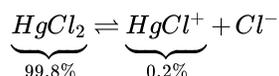


Figure 11.2.2: Mercury Bromide Crystals.

Some covalent substances behave as **weak electrolytes**—their solutions allow only a small current flow, but it is greater than that of the pure solvent. An example is mercury(II) chloride (seen in the Figure above). For a 100-*M* HgCl₂ solution the meter reading shows only about 0.2 percent as much current as for 0.10 *M* NaCl. A crystal of HgCl₂ consists of discrete molecules, like those shown for HgBr₂ in Figure 11.2.2. When the solid dissolves, most of these molecules remain intact, but a few dissociate into ions according to the equation



(The double arrows indicate that the ionization proceeds only to a limited extent and an equilibrium state is attained.) Since only 0.2 percent of the HgCl₂ forms ions, the 0.10 *M* solution can conduct only about 0.2 percent as much current as 0.10 *M* NaCl.

Conductivity measurements can tell us more than whether a substance is a strong, a weak, or a nonelectrolyte. Consider, for instance, the data in Table 11.2.1 which shows the electrical current conducted through various aqueous solutions under identical conditions. At the rather low concentration of 0.001 *M*, the strong electrolyte solutions conduct between 2500 and 10 000 times as much current as pure H₂O and about 10 times as much as the weak electrolytes HC₂H₃O₂ (acetic acid) and NH₃ (ammonia).

Closer examination of the data for strong electrolytes reveals that some compounds which contain H or OH groups [such as HCl or Ba(OH)₂] conduct unusually well. If these compounds are excluded, we find that 1:1 electrolytes (compounds which consist of equal numbers of +1 ions and -1 ions) usually conduct about half as much current as 2:2 electrolytes (+2 and -2 ions), 1:2 electrolytes (+1 and -2 ions), or 2:1 electrolytes (+2 and -1 ions).

TABLE 11.2.1: Electrical Current Conducted Through Various 0.001 *M* Aqueous Solutions at 18°C.*

Substance	Current /mA	Substance	Current /mA
Pure Water		1:2 Electrolytes	
H ₂ O	3.69 × 10 ⁻⁴	Na ₂ SO ₄	2.134
Weak Electrolytes		Na ₂ CO ₃	2.24
HC ₂ H ₃ O ₂	0.41	K ₂ CO ₃	2.660
NH ₃	0.28	2:1 Electrolytes	
1:1 Electrolytes		MgCl ₂	2.128
NaCl	1.065	CaCl ₂	2.239
NaI	1.069	SrCl ₂	2.290
KCl	1.273	BaCl ₂	2.312
KI	1.282	Ba(OH) ₂	4.14
AgNO ₃	1.131	2:2 Electrolytes	
HCl	3.77	MgSO ₄	2.00

Substance	Current /mA	Substance	Current /mA
HNO ₃	3.75	CaSO ₄	2.086
NaOH	2.08	CuSO ₄	1.97
KOH	2.34	ZnSO ₄	1.97

* All measurements refer to a cell in which the distance between the electrodes is 1.0 mm and the area of each electrode is 1.0 cm². A potential difference of 1.0 V is applied to produce the tabulated currents.

There is a simple reason for this behavior. Under similar conditions, most ions move through water at comparable speeds. This means that ions like Mg²⁺ or SO₄²⁻, which are doubly charged, will carry twice as much current through the solution as will singly charged ions like Na⁺ or Cl⁻. Consequently, a 0.001 M solution of a 2:2 electrolyte like MgSO₄ will conduct about twice as well as a 0.001 M solution of a 1:1 electrolyte like NaCl.

A similar argument applies to solutions of 1:2 and 2:1 electrolytes. A solution like 0.001 M Na₂SO₄ conducts about twice as well as 0.001 M NaCl partly because there are twice as many Na⁻ ions available to move when a battery is connected, but also because SO₄²⁻ ions carry twice as much charge as Cl⁻ ions when moving at the same speed. These differences in conductivity between different types of strong electrolytes can sometimes be very useful in deciding what ions are actually present in a given electrolyte solution as the following example makes clear.

A second, slightly more subtle, conclusion can be drawn from the data in Table 11.2.1. When an electrolyte dissolves, each type of ion makes an independent contribution to the current the solution conducts. This can be seen by comparing NaCl with KCl, and NaI with KI. In each case the compound containing K⁺ conducts about 0.2 mA more than the one containing Na⁺. If we apply this observation to Na₂CO₃ and K₂CO₃, each of which produces twice as many Na⁺ or K⁺ ions in solution, we find that the difference in current is also twice as great—about 0.4 mA.

Thus conductivity measurements confirm our statement that each ion exhibits its own characteristic properties in aqueous solutions, independent of the presence of other ions. One such characteristic property is the quantity of electrical current that a given concentration of a certain type of ion can carry.

✓ Example 11.2.1: Ions

At 18°C a 0.001-*M* aqueous solution of potassium hydrogen carbonate, KHCO₃, conducts a current of 1.10 mA in a cell of the same design as that used to obtain the data in Table 11.1. What ions are present in solution?

Solution

Referring to Table 6.2 which lists possible polyatomic ions, we can arrive at three possibilities for the ions from which KHCO₃ is made:

- K⁺ and H⁺ and C⁴⁺ and three O²⁻
- K⁺ and H⁺ and CO₃²⁻
- K⁺ and HCO₃⁻

Since the current conducted by the solution falls in the range of 1.0 to 1.3 mA characteristic of 1:1 electrolytes, possibility *c* is the only reasonable choice.

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