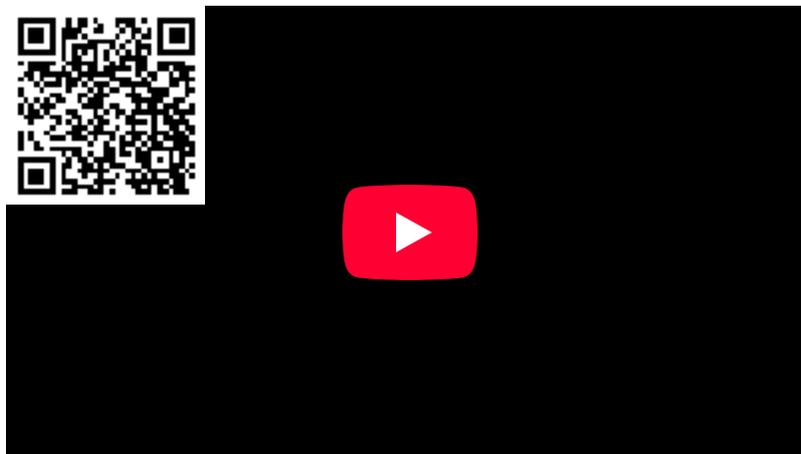


13.9: Le Chatelier's Principle

Often it is useful to predict *qualitatively* (without doing calculations such as those just described) what will happen to a [system at equilibrium](#) when conditions such as temperature or volume change or when a reactant or product is added or removed from the reaction mixture. Fortunately a simple rule, **Le Chatelier's principle**, enables us to make such qualitative predictions. This rule states that *if a system is in equilibrium and some factor in the equilibrium conditions is altered, then the system will (if possible) adjust to a new equilibrium state so as to counteract this alteration to some degree.*



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