

### 1.4.6.3: Practice Trends

#### Exercise 1.4.6.3.1

Choose the element in each pair that has a larger atomic radius (size).

- a) S or Ar
- b) Na or K
- c) O or P

**Answer a**

S

**Answer b**

K

**Answer c**

P

#### Exercise 1.4.6.3.1

Choose the element in each pair that has a greater ionization energy.

- d) Sc or K
- e) Ca or Mg
- f) Al or P

**Answer d**

Sc

**Answer e**

Mg

**Answer f**

P

#### Exercise 1.4.6.3.1

Choose the element in each pair that has more metallic character.

- g) Si or Al
- h) Ca or Ba
- i) Ti or Fe

**Answer g**

Al

**Answer h**

Ba

**Answer i**

Ti

#### Exercise 1.4.6.3.1

C	F
Ge	Br

Out of the four elements listed above, choose the one that has

- j) the smallest radius
- k) the least ionization energy
- l) the most metallic character

**Answer j**

F

**Answer k**

Ge

**Answer l**

Ge

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