

1.4.6.3: Practice Trends

Exercise 1.4.6.3.1

Choose the element in each pair that has a larger atomic radius (size).

- a) S or Ar
- b) Na or K
- c) O or P

Answer a

S

Answer b

K

Answer c

P

Exercise 1.4.6.3.1

Choose the element in each pair that has a greater ionization energy.

- d) Sc or K
- e) Ca or Mg
- f) Al or P

Answer d

Sc

Answer e

Mg

Answer f

P

Exercise 1.4.6.3.1

Choose the element in each pair that has more metallic character.

- g) Si or Al
- h) Ca or Ba
- i) Ti or Fe

Answer g

Al

Answer h

Ba

Answer i

Ti

Exercise 1.4.6.3.1

C F
Ge Br

Out of the four elements listed above, choose the one that has

- j) the smallest radius
- k) the least ionization energy
- l) the most metallic character

Answer j

F

Answer k

Ge

Answer l

Ge

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