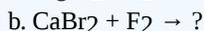
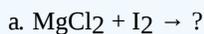


✓ Example 1.7.4.1

Will a single-replacement reaction occur? If so, identify the products.



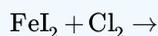
Solution

a. Because iodine is below chlorine on the periodic table, a single-replacement reaction will not occur.

b. Because fluorine is above bromine on the periodic table, a single-replacement reaction will occur, and the products of the reaction will be CaF_2 and Br_2 .

? Exercise 1.7.4.1

Will a single-replacement reaction occur? If so, identify the products.



Answer

Yes; FeCl_2 and I_2

Chemical reactivity trends are easy to predict when replacing anions in simple ionic compounds—simply use their relative positions on the periodic table. However, when replacing the cations, the trends are not as straightforward. This is partly because there are so many elements that can form cations; an element in one column on the periodic table may replace another element nearby, or it may not. A list called the **activity series** does the same thing the periodic table does for halogens: it lists the elements that will replace elements below them in single-replacement reactions. A simple activity series is shown below.

? Activity Series for Cation Replacement in Single-Replacement Reactions

- Li
- K
- Ba
- Sr
- Ca
- Na
- Mg
- Al
- Mn
- Zn
- Cr
- Fe
- Ni
- Sn
- Pb
- H_2
- Cu
- Hg
- Ag
- Pd
- Pt
- Au

Using the activity series is similar to using the positions of the halogens on the periodic table. An element on top will replace an element below it in compounds undergoing a single-replacement reaction. Elements will not replace elements above them in

compounds.

✓ Example 1.7.4.2

Use the activity series to predict the products, if any, of each equation.

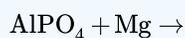
1. $\text{FeCl}_2 + \text{Zn} \rightarrow ?$
2. $\text{HNO}_3 + \text{Au} \rightarrow ?$

Solution

1. Because zinc is above iron in the activity series, it will replace iron in the compound. The products of this single-replacement reaction are ZnCl_2 and Fe.
2. Gold is below hydrogen in the activity series. As such, it will not replace hydrogen in a compound with the nitrate ion. No reaction is predicted.

? Exercise 1.7.4.2

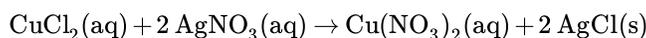
Use the activity series to predict the products, if any, of this equation.



Answer

$\text{Mg}_3(\text{PO}_4)_2$ and Al

A **double-replacement reaction** occurs when parts of two ionic compounds are exchanged, making two new compounds. A characteristic of a double-replacement equation is that there are two compounds as reactants and two different compounds as products. An example is



There are two equivalent ways of considering a double-replacement equation: either the cations are swapped, or the anions are swapped. (You cannot swap both; you would end up with the same substances you started with.) Either perspective should allow you to predict the proper products, as long as you pair a cation with an anion, and not a cation with a cation or an anion with an anion.

✓ Example 1.7.4.3

Predict the products of this double-replacement equation:



Solution

Thinking about the reaction as either switching the cations or switching the anions, we would expect the products to be BaSO_4 and NaCl.

? Exercise 1.7.4.3

Predict the products of this double-replacement equation:



Answer

KNO_3 and AgBr

Predicting whether a double-replacement reaction occurs is somewhat more difficult than predicting a single-replacement reaction. However, there is one type of double-replacement reaction that we can predict: the precipitation reaction. A **precipitation reaction** occurs when two ionic compounds are dissolved in water and form a new ionic compound that does not dissolve; this new

compound falls out of solution as a solid **precipitate**. The formation of a solid precipitate is the driving force that makes the reaction proceed.

To judge whether double-replacement reactions will occur, we need to know what kinds of ionic compounds form precipitates. For this, we use **solubility rules**, which are general statements that predict which ionic compounds dissolve (are soluble) and which do not (are not soluble, or insoluble). Table 1.7.4.1 lists some general solubility rules. We need to consider each ionic compound (both the reactants and the possible products) in light of the solubility rules. If a compound is soluble, we use the (aq) label with it, indicating that it dissolves. If a compound is not soluble, we use the (s) label with it and assume that it will precipitate out of solution. If everything is soluble, then no reaction will be expected.

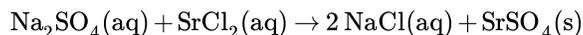
Table 1.7.4.1: Some Useful Solubility Rules (soluble)

These compounds generally dissolve in water (are soluble):	Exceptions:
All compounds of Li^+ , Na^+ , K^+ , Rb^+ , Cs^+ , and NH_4^+	None
All compounds of NO_3^- and $\text{C}_2\text{H}_3\text{O}_2^-$	None
Compounds of Cl^- , Br^- , I^-	Ag^+ , Hg_2^{2+} , Pb^{2+}
Compounds of SO_4^{2-}	Hg_2^{2+} , Pb^{2+} , Sr^{2+} , Ba^{2+}

Table 1.7.4.2: Some Useful Solubility Rules (insoluble)

These compounds generally do not dissolve in water (are insoluble):	Exceptions:
Compounds of CO_3^{2-} and PO_4^{3-}	Compounds of Li^+ , Na^+ , K^+ , Rb^+ , Cs^+ , and NH_4^+
Compounds of OH^-	Compounds of Li^+ , Na^+ , K^+ , Rb^+ , Cs^+ , NH_4^+ , Sr^{2+} , and Ba^{2+}

For example, consider the possible double-replacement reaction between Na_2SO_4 and SrCl_2 . The solubility rules say that all ionic sodium compounds are soluble and all ionic chloride compounds are soluble, except for Ag^+ , Hg_2^{2+} , and Pb^{2+} , which are not being considered here. Therefore, Na_2SO_4 and SrCl_2 are both soluble. The possible double-replacement reaction products are NaCl and SrSO_4 . Are these soluble? NaCl is (by the same rule we just quoted), but what about SrSO_4 ? Compounds of the sulfate ion are generally soluble, but Sr^{2+} is an exception: we expect it to be insoluble—a precipitate. Therefore, we expect a reaction to occur, and the balanced chemical equation would be:



You would expect to see a visual change corresponding to SrSO_4 precipitating out of solution (Figure 1.7.4.2).



Figure 1.7.4.2: Double-Replacement Reactions. Some double-replacement reactions are obvious because you can see a solid precipitate coming out of solution. Source: Photo courtesy of Choij, http://commons.wikimedia.org/wiki/File:Copper_solution.jpg (opens in new window).

✓ Example 1.7.4.4

Will a double-replacement reaction occur? If so, identify the products.

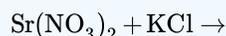
1. $\text{Ca}(\text{NO}_3)_2 + \text{KBr} \rightarrow ?$
2. $\text{NaOH} + \text{FeCl}_2 \rightarrow ?$

Solution

1. According to the solubility rules, both $\text{Ca}(\text{NO}_3)_2$ and KBr are soluble. Now we consider what the double-replacement products would be by switching the cations (or the anions)—namely, CaBr_2 and KNO_3 . However, the solubility rules predict that these two substances would also be soluble, so no precipitate would form. Thus, we predict no reaction in this case.
2. According to the solubility rules, both NaOH and FeCl_2 are expected to be soluble. If we assume that a double-replacement reaction may occur, we need to consider the possible products, which would be NaCl and $\text{Fe}(\text{OH})_2$. NaCl is soluble, but, according to the solubility rules, $\text{Fe}(\text{OH})_2$ is not. Therefore, a reaction would occur, and $\text{Fe}(\text{OH})_2(\text{s})$ would precipitate out of solution. The balanced chemical equation is



? Exercise 1.7.4.4

**Answer**

No reaction; all possible products are soluble.

Key Takeaways

- A single-replacement reaction replaces one element for another in a compound.
- The periodic table or an activity series can help predict whether single-replacement reactions occur.
- A double-replacement reaction exchanges the cations (or the anions) of two ionic compounds.
- A precipitation reaction is a double-replacement reaction in which one product is a solid precipitate.
- Solubility rules are used to predict whether some double-replacement reactions will occur.

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