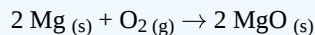


1.7.6.2: Practice Stoichiometry part 2

Exercise 1.7.6.2.1

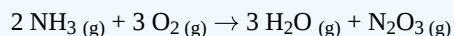


In the reaction above, how many grams of MgO would be produced from 6.87 g of Mg and 5.22 g O₂?

Answer

11.4 g MgO (Mg is limiting)

Exercise 1.7.6.2.1



In the reaction above, how many grams of N₂O₃ would be produced from 12.76 g of NH₃ and 17.44 g O₂?

Answer

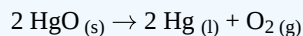
13.81 g N₂O₃ (O₂ is limiting)

In the reaction above, how many grams of N₂O₃ would be produced from 0.765 moles NH₃ and 1.206 moles O₂?

Answer

29.1 g (NH₃ is limiting)

Exercise 1.7.6.2.1

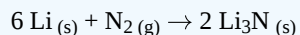


If you heat 375.0 grams of HgO and obtain 248.4 g Hg, what is your percent yield?

Answer

71.52 % (expect 347.3 g Hg product)

Exercise 1.7.6.2.1

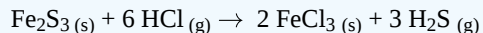


If you heat 38.44 grams of Li (with plenty of nitrogen) and obtain 44.74 g Li₃N, what is your percent yield?

Answer

69.57 % (expect 64.31 g Li₃N product)

Exercise 1.7.6.2.1



If you treat 17.82 grams of Fe₂S₃ (with plenty of HCl) and obtain 21.36 g FeCl₃, what is your percent yield?

Answer

76.83 % (expect 27.80 g FeCl₃ product)

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