

1.6.3.1: Practice Naming

Exercise 1.6.3.1.1

Name the compound below. Hint: Is it ionic or covalent? Is the cation predictable or variable?

**Hint**

It is ionic. It is predictable, so no Roman numerals are needed. Just name cation and name anion.

Answer

magnesium sulfate

Exercise 1.6.3.1.1

Name the compound below. Hint: Is it ionic or covalent? Is the cation predictable or variable?

**Hint**

It is ionic. It is variable, so you must determine the charge needed to make a total of zero charge. Roman numeral is needed.

Answer

copper (I) sulfide [The copper must be Cu^{1+} in order for two of them to balance the one S^{2-}]

Exercise 1.6.3.1.1

Name the compound below. Hint: Is it ionic or covalent? Is the cation predictable or variable?

**Hint**

It is ionic. It is variable, so you must determine the charge needed to make a total of zero charge. Roman numeral is needed.

Answer

iron (III) nitrate

Exercise 1.6.3.1.1

Name the compound below. Hint: Is it ionic or covalent? Is the cation predictable or variable?

**Hint**

It is ionic. It is predictable, so no Roman numerals are needed. Just name cation and name anion.

Answer

aluminum chlorate

Exercise 1.6.3.1.1

Name the compound below. (We are not dealing with peroxides. The anion is regular oxide.) Hint: Is it ionic or covalent? Is the cation predictable or variable?

PtO₂

Hint

It is ionic. It is variable, so you must determine the charge needed to make a total of zero charge. Roman numeral is needed.

Answer

platinum (IV) oxide

Exercise 1.6.3.1.1

Write the formula of the compound below. Hint: draw each ion first. Then put them together so they make a compound with zero charge.

calcium nitride

Hint

Ca²⁺ and N³⁻

Answer

Ca₃N₂

Exercise 1.6.3.1.1

Write the formula of the compound below. Hint: draw each ion first. Then put them together so they make a compound with zero charge.

lithium bisulfite

Hint

Li¹⁺ and HSO₃¹⁻

Answer

LiHSO₃

Exercise 1.6.3.1.1

Write the formula of the compound below. Hint: draw each ion first. Then put them together so they make a compound with zero charge.

nickel (III) acetate

Hint

Ni³⁺ and C₂H₃O₂¹⁻

Answer

Ni(C₂H₃O₂)₃

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