

22.3.4: iv. Review Exercise Solutions

Q1

a. For non-degenerate point groups one can simply multiply the representations (since only one representation will be obtained):

$$a_1 \otimes b_1 = b_1$$

Constructing a "box" in this case is unnecessary since it would only contain a single row. Two unpaired electrons will result in a singlet ($S=0, M_S=0$), and three triplets ($S=1, M_S=1, S=1, M_S=0, S=1, M_S=-1$). The states will be: ${}^3B_1(M_S=1)$, ${}^3B_1(M_S=0)$, (${}^3B_1(M_S=-1)$), and ${}^1B_1(M_S=0)$.

b. Remember that when coupling non-equivalent linear molecule angular momenta, one simply adds the individual M_L values and vector couples the electron spin. So, in this case ($1\pi_u^1 2\pi_u^1$), we have M_L values of 1+1, 1-1, -1+1, and -1-1 (2,0,0, and -2). The term symbol Δ is used to denote the spatially doubly degenerate level ($M_L = \pm 2$) and there are two distinct spatially non-degenerate levels denote by the term symbol $\Sigma(M_L = 0)$. Again, two unpaired electrons will result in a singlet ($S = 0, M_S = 0$), and three triplets ($S = 1, M_S = 1; S = 1, M_S = 0; S = 1, M_S = -1$). The states generated are then:

$${}^1\Delta(M_L = 2); \text{ one state } (M_S = 0),$$

$${}^1\Delta(M_L = -2); \text{ one state } (M_S = 0),$$

$${}^3\Delta(M_L = 2); \text{ one state } (M_S = 1, 0, \text{ and } -1),$$

$${}^3\Delta(M_L = -2); \text{ one state } (M_S = 1, 0, \text{ and } -1),$$

$${}^1\Sigma(M_L = 0); \text{ one state } (M_S = 0),$$

$${}^1\Sigma(M_L = 0); \text{ one state } (M_S = 0),$$

$${}^3\Sigma(M_L = 0); \text{ one state } (M_S = 1, 0, \text{ and } -1), \text{ and}$$

$${}^3\Sigma(M_L = 0); \text{ one state } (M_S = 1, 0, \text{ and } -1)$$

c. Constructing the "box" for two equivalent π electrons one obtains:

M_S	M_L	2	1	0
1				$ \pi_1\alpha\pi_{-1}\alpha $
0		$ \pi_1\alpha\pi_1\beta $		$ \pi_1\alpha\pi_{-1}\beta $, $ \pi_{-1}\alpha\pi_1\beta $

From this "box" one obtains six states:

$${}^1\Delta(M_L = 2); \text{ one state } (M_S = 0),$$

$^1\Delta(M_L = -2)$; one state ($M_S = 0$),

$^1\Delta(M_L = 0)$; one state ($M_S = 0$),

$^3\Delta(M_L = 0)$; three states ($M_S = 1, 0, \text{ and } -1$),

d. It is not necessary to construct a "box" when coupling non-equivalent angular momenta since the vector coupling results in a range from the sum of the two individual angular momenta to the absolute value of their difference. In this case, $3d^14d^1$, $L=4, 3, 2, 1, 0$, and $S=1, 0$. The term symbols are: $^3G, ^1G, ^3F, ^1F, ^3D, ^1D, ^3P, ^1P, ^3S$, and 1S . The L and S angular momenta can be vector coupled to produce further splitting into levels:

$$J = L + S \dots |L - S|.$$

Denoting J as a term symbol subscript one can identify all the levels and the subsequent $(2J + 1)$ states:

3G_5 (11 states),

3G_4 (9 states),

3G_3 (7 states),

1G_4 (9 states),

3F_4 (9 states),

3F_3 (7 states),

3F_2 (5 states),

1F_3 (7 states),

3D_3 (7 states),

3D_2 (5 states),

3D_1 (3 states),

1D_2 (5 states),

3P_2 (5 states),

3P_1 (3 states),

3P_0 (1 states),

1P_1 (3 states),

3S_5 (3 states), and

1S_0 (1 states).

e. Construction of a "box" for the two equivalent d electrons generates (note the "box" has been turned side ways for convenience):

M_L	M_S	1	0
4			$ d_2\alpha d_2\beta\rangle$
3		$ d_2\alpha d_1\alpha\rangle$	$ d_2\alpha d_1\beta\rangle, d_2\beta d_1\alpha\rangle$
2		$ d_2\alpha d_0\alpha\rangle$	$ d_2\alpha d_0\beta\rangle, d_2\beta d_0\alpha\rangle, d_1\alpha d_1\beta\rangle$
1		$ d_1\alpha d_0\alpha\rangle, d_2\alpha d_{-1}\alpha\rangle$	$ d_1\alpha d_0\beta\rangle, d_1\beta d_0\alpha\rangle, d_2\alpha d_{-1}\beta\rangle, d_2\beta d_{-1}\alpha\rangle$
0		$ d_2\alpha d_{-2}\alpha\rangle, d_1\alpha d_{-1}\alpha\rangle$	$ d_2\alpha d_{-2}\beta\rangle, d_2\beta d_{-2}\alpha\rangle, d_1\alpha d_{-1}\beta\rangle, d_1\beta d_{-1}\alpha\rangle, d_0\alpha d_0\beta\rangle$

The term symbols are: 1G , 3F , 1D , 3P , and 1S . The L and S angular momenta can be vector coupled to produce further splitting into levels:

1G_4 (9 states),

3F_4 (9 states),

3F_3 (7 states),

3F_2 (5 states),

1D_2 (5 states),

3P_2 (5 states),

3P_1 (3 states),

3P_0 (1 states), and

1S_0 (1 states).

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