

22.3.4: iv. Review Exercise Solutions

Q1

a. For non-degenerate point groups one can simply multiply the representations (since only one representation will be obtained):

$$a_1 \otimes b_1 = b_1$$

Constructing a "box" in this case is unnecessary since it would only contain a single row. Two unpaired electrons will result in a singlet ($S=0$, $M_S=0$), and three triplets ($S=1$, $M_S=1$, $S=1$, $M_S=0$, $S=1$, $M_S=-1$). The states will be: $^3B_1(M_S=1)$, $^3B_1(M_S=0)$, ($^3B_1(M_S=-1)$), and $^1B_1(M_S=0)$.

b. Remember that when coupling non-equivalent linear molecule angular momenta, one simply adds the individual L_z values and vector couples the electron spin. So, in this case ($1\pi_u^1 2\pi_u^1$), we have M_L values of 1+1, 1-1, -1+1, and -1-1 (2,0,0, and -2). The term symbol Δ is used to denote the spatially doubly degenerate level ($M_L = \pm 2$) and there are two distinct spatially non-degenerate levels denote by the term symbol $\Sigma(M_L = 0)$. Again, two unpaired electrons will result in a singlet ($S = 0$, $M_S = 0$), and three triplets ($S = 1$, $M_S = 1$; $S = 1$, $M_S = 0$; $S = 1$, $M_S = -1$). The states generated are then:

$$^1\Delta(M_L = 2); \text{ one state } (M_S = 0),$$

$$^1\Delta(M_L = -2); \text{ one state } (M_S = 0),$$

$$^3\Delta(M_L = 2); \text{ one state } (M_S = 1, 0, \text{ and } -1),$$

$$^3\Delta(M_L = -2); \text{ one state } (M_S = 1, 0, \text{ and } -1),$$

$$^1\Sigma(M_L = 0); \text{ one state } (M_S = 0),$$

$$^1\Sigma(M_L = 0); \text{ one state } (M_S = 0),$$

$$^3\Sigma(M_L = 0); \text{ one state } (M_S = 1, 0, \text{ and } -1), \text{ and}$$

$$^3\Sigma(M_L = 0); \text{ one state } (M_S = 1, 0, \text{ and } -1)$$

c. Constructing the "box" for two equivalent π electrons one obtains:

M_L	2	1	0
M_S			
1			$ \pi_1\alpha\pi_{-1}\alpha $
0	$ \pi_1\alpha\pi_1\beta $		$ \pi_1\alpha\pi_{-1}\beta $, $ \pi_{-1}\alpha\pi_1\beta $

From this "box" one obtains six states:

$$^1\Delta(M_L = 2); \text{ one state } (M_S = 0),$$

$$^1\Delta(M_L = -2); \text{ one state } (M_S = 0),$$

$$^1\Delta(M_L = 0); \text{ one state } (M_S = 0),$$

$$^3\Delta(M_L = 0); \text{ three states } (M_S = 1, 0, \text{ and } -1),$$

d. It is not necessary to construct a "box" when coupling non-equivalent angular momenta since the vector coupling results in a range from the sum of the two individual angular momenta to the absolute value of their difference. In this case, $3d^1 4d^1$, $L=4, 3, 2, 1, 0$, and $S=1, 0$. The term symbols are: $^3G, ^1G, ^3F, ^1F, ^3D, ^1D, ^3P, ^1P, ^3S$, and 1S . The L and S angular momenta can be vector coupled to produce further splitting into levels:

$$J = L + S \dots |L - S|.$$

Denoting J as a term symbol subscript one can identify all the levels and the subsequent $(2J + 1)$ states:

$$^3G_5 \text{ (11 states),}$$

$$^3G_4 \text{ (9 states),}$$

$$^3G_3 \text{ (7 states),}$$

$$^1G_4 \text{ (9 states),}$$

$$^3F_4 \text{ (9 states),}$$

$$^3F_3 \text{ (7 states),}$$

$$^3F_2 \text{ (5 states),}$$

$$^1F_3 \text{ (7 states),}$$

$$^3D_3 \text{ (7 states),}$$

$$^3D_2 \text{ (5 states),}$$

$$^3D_1 \text{ (3 states),}$$

1D_2 (5 states),

3P_2 (5 states),

3P_1 (3 states),

3P_0 (1 states),

1P_1 (3 states),

3S_1 (3 states), and

1S_0 (1 states).

e. Construction of a "box" for the two equivalent d electrons generates (note the "box" has been turned side ways for convenience):

M_L	M_S	1	0
4			$ d_2\alpha d_2\beta $
3		$ d_2\alpha d_1\alpha $	$ d_2\alpha d_1\beta , d_2\beta d_1\alpha $
2		$ d_2\alpha d_0\alpha $	$ d_2\alpha d_0\beta , d_1\alpha d_1\beta , d_2\beta d_0\alpha $
1		$ d_1\alpha d_0\alpha , d_2\alpha d_{-1}\alpha $	$ d_1\alpha d_0\beta , d_1\beta d_0\alpha , d_2\alpha d_{-1}\beta , d_2\beta d_{-1}\alpha $
0		$ d_2\alpha d_{-2}\alpha , d_1\alpha d_{-1}\alpha $	$ d_2\alpha d_{-2}\beta , d_2\beta d_{-2}\alpha , d_1\alpha d_{-1}\beta , d_1\beta d_{-1}\alpha , d_0\alpha d_0\beta $

The term symbols are: 1G , 3F , 1D , 3P , and 1S . The L and S angular momenta can be vector coupled to produce further splitting into levels:

1G_4 (9 states),

3F_4 (9 states),

3F_3 (7 states),

3F_2 (5 states),

1D_2 (5 states),

3P_2 (5 states),

3P_1 (3 states),

3P_0 (1 states), and

1S_0 (1 states).

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