

## 6.4: Gas Laws

### Learning Objectives

- Learn what is meant by the term *gas laws*.
- Learn and apply Boyle's law.
- Learn and apply Charles's law.

When seventeenth-century scientists began studying the physical properties of gases, they noticed some simple relationships between some of the measurable properties of the gas. Take pressure ( $P$ ) and volume ( $V$ ), for example. Scientists noted that for a given amount of a gas (usually expressed in units of moles [ $n$ ]), if the temperature ( $T$ ) of the gas was kept constant, pressure and volume were related: As one increases, the other decreases. As one decreases, the other increases. We say that pressure and volume are *inversely related*.

There is more to it, however: pressure and volume of a given amount of gas at constant temperature are *numerically* related. If you take the pressure value and multiply it by the volume value, the product is a constant for a given amount of gas at a constant temperature:

$$P \times V = \text{constant at constant } n \text{ and } T$$

If either volume or pressure changes while amount and temperature stay the same, then the other property must change so that the product of the two properties still equals that same constant. That is, if the original conditions are labeled  $P_1$  and  $V_1$  and the new conditions are labeled  $P_2$  and  $V_2$ , we have

$$P_1V_1 = \text{constant} = P_2V_2$$

where the properties are assumed to be multiplied together. Leaving out the middle part, we have simply

$$P_1V_1 = P_2V_2 \text{ at constant } n \text{ and } T$$

This equation is an example of a gas law. A **gas law** is a simple mathematical formula that allows you to model, or predict, the behavior of a gas. This particular gas law is called **Boyle's law**, after the English scientist Robert Boyle, who first announced it in 1662. Figure 6.4.1 shows two representations of how Boyle's law works.

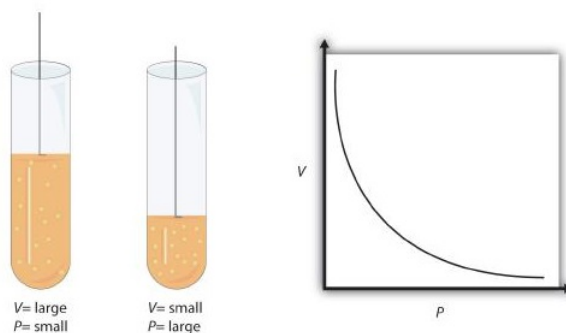


Figure 6.4.1: Boyle's Law. A piston having a certain pressure and volume (left piston) will have half the volume when its pressure is twice as much (right piston). One can also plot  $P$  versus  $V$  for a given amount of gas at a certain temperature; such a plot will look like the graph on the right.

Boyle's law is an example of a second type of mathematical problem we see in chemistry—one based on a mathematical formula. Tactics for working with mathematical formulas are different from tactics for working with conversion factors. First, most of the questions you will have to answer using formulas are word-type questions, so the first step is to identify what quantities are known and assign them to variables. Second, in most formulas, some mathematical rearrangements (i.e., algebra) must be performed to solve for an unknown variable. The rule is that to find the value of the unknown variable, you must mathematically isolate the unknown variable *by itself and in the numerator* of one side of the equation. Finally, units must be consistent. For example, in Boyle's law there are two pressure variables; they must have the same unit. There are also two volume variables; they also must have the same unit. In most cases, it won't matter *what* the unit is, but the unit must be the *same* on both sides of the equation.

## ✓ Example 6.4.1

A sample of gas has an initial pressure of 2.44 atm and an initial volume of 4.01 L. Its pressure changes to 1.93 atm. What is the new volume if temperature and amount are kept constant?

**Solution**

First, determine what quantities we are given. We are given an initial pressure and an initial volume, so let these values be  $P_1$  and  $V_1$ :

$$P_1 = 2.44 \text{ atm and } V_1 = 4.01 \text{ L}$$

We are given another quantity, final pressure of 1.93 atm, but not a final volume. This final volume is the variable we will solve for.

$$P_2 = 1.93 \text{ atm and } V_2 = ? \text{ L}$$

Substituting these values into Boyle's law, we get

$$(2.44 \text{ atm})(4.01 \text{ L}) = (1.93 \text{ atm})V_2$$

To solve for the unknown variable, we isolate it by dividing both sides of the equation by 1.93 atm—both the number *and* the unit:

$$\frac{(2.44 \text{ atm})(4.01 \text{ L})}{1.93 \text{ atm}} = \frac{(1.93 \text{ atm}) V_2}{1.93 \text{ atm}}$$

Note that, on the left side of the equation, the unit *atm* is in the numerator and the denominator of the fraction. They cancel algebraically, just as a number would. On the right side, the unit *atm* and the number 1.93 are in the numerator and the denominator, so the entire quantity cancels:

$$\frac{(2.44 \text{ atm})(4.01 \text{ L})}{1.93 \text{ atm}} = \frac{(1.93 \text{ atm}) V_2}{1.93 \text{ atm}}$$

What we have left is

$$\frac{(2.44)(4.01 \text{ L})}{1.93} = V_2$$

Now we simply multiply and divide the numbers together and combine the answer with the *L* unit, which is a unit of volume. Doing so, we get  $V_2 = 5.07 \text{ L}$

Does this answer make sense? We know that pressure and volume are inversely related; as one decreases, the other increases. Pressure is decreasing (from 2.44 atm to 1.93 atm), so volume should be increasing to compensate, and it is (from 4.01 L to 5.07 L). So the answer makes sense based on Boyle's law.

## ? Exercise 6.4.1

If  $P_1 = 334 \text{ torr}$ ,  $V_1 = 37.8 \text{ mL}$ , and  $P_2 = 102 \text{ torr}$ , what is  $V_2$ ?

**Answer**

124 mL

As mentioned, you can use any units for pressure or volume, but both pressures must be expressed in the same units, and both volumes must be expressed in the same units.

## ✓ Example 6.4.2

A sample of gas has an initial pressure of 722 torr and an initial volume of 88.8 mL. Its volume changes to 0.663 L. What is the new pressure?

**Solution**

We can still use Boyle's law to answer this, but now the two volume quantities have different units. It does not matter which unit we change, as long as we perform the conversion correctly. Let us change the 0.663 L to milliliters:

$$0.663 \text{ L} \times \frac{1000 \text{ ml}}{1 \text{ L}} = 663 \text{ ml}$$

Now that both volume quantities have the same units, we can substitute into Boyle's law:

$$(722 \text{ torr})(88.8 \text{ ml}) = P_2(663 \text{ ml})$$
$$\frac{(722 \text{ torr})(88.8 \text{ ml})}{(663 \text{ ml})} = P_2$$

The mL units cancel, and we multiply and divide the numbers to get  $P_2 = 96.7 \text{ torr}$

The volume is increasing, and the pressure is decreasing, which is as expected for Boyle's law.

### ? Exercise 6.4.2

If  $V_1 = 456 \text{ mL}$ ,  $P_1 = 308 \text{ torr}$ , and  $P_2 = 1.55 \text{ atm}$ , what is  $V_2$ ?

**Answer**

119 mL

There are other measurable characteristics of a gas. One of them is temperature ( $T$ ). Perhaps one can vary the temperature of a gas sample and note what effect it has on the other properties of the gas. Early scientists did just this, discovering that if the amount of a gas and its pressure are kept constant, then changing the temperature changes the volume ( $V$ ). As temperature increases, volume increases; as temperature decreases, volume decreases. We say that these two characteristics are *directly related*.

A mathematical relationship between  $V$  and  $T$  should be possible except for one thought: what temperature scale should we use? We know from Chapter 2 that science uses several possible temperature scales. Experiments show that the volume of a gas is related to its absolute temperature in Kelvin, *not its temperature in degrees Celsius*. If the temperature of a gas is expressed in kelvins, then experiments show that the *ratio* of volume to temperature is a constant:

$$\frac{V}{T} = \text{constant}$$

We can modify this equation as we modified Boyle's law: the initial conditions  $V_1$  and  $T_1$  have a certain value, and the value must be the same when the conditions of the gas are changed to some new conditions  $V_2$  and  $T_2$ , as long as pressure and the amount of the gas remain constant. Thus, we have another gas law:

$$\frac{V_1}{T_1} = \frac{V_2}{T_2} \text{ at constant } P \text{ and } n$$

This gas law is commonly referred to as **Charles's law**, after the French scientist Jacques Charles, who performed experiments on gases in the 1780s. The tactics for using this mathematical formula are similar to those for Boyle's law. To determine an unknown quantity, use algebra to isolate the unknown variable by itself and in the numerator; the units of similar variables must be the same. But we add one more tactic: all temperatures must be expressed in the absolute temperature scale (Kelvin). As a reminder, we review the conversion between the absolute temperature scale and the Celsius temperature scale:

$$\text{K} = ^\circ\text{C} + 273$$

where K represents the temperature in kelvins, and  $^\circ\text{C}$  represents the temperature in degrees Celsius.

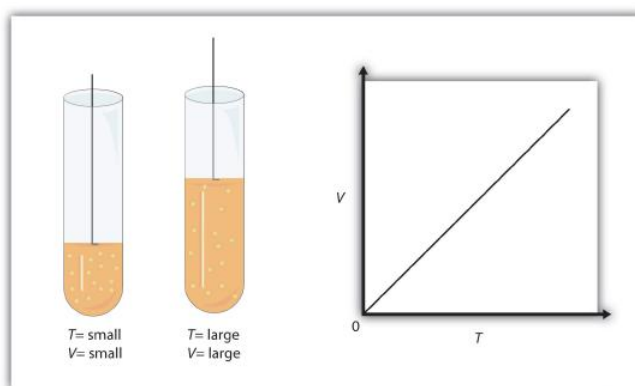


Figure 6.4.2: Charles's Law. A piston having a certain volume and temperature (left piston) will have twice the volume when its temperature is twice as much (right piston). One can also plot  $V$  versus  $T$  for a given amount of gas at a certain pressure; such a plot will look like the graph on the right.

### ✓ Example 6.4.3

A sample of gas has an initial volume of 34.8 mL and an initial temperature of 315 K. What is the new volume if the temperature is increased to 559 K? Assume constant pressure and amount for the gas.

#### Solution

First, we assign the given values to their variables. The initial volume is  $V_1$ , so  $V_1 = 34.8$  mL, and the initial temperature is  $T_1$ , so  $T_1 = 315$  K. The temperature is increased to 559 K, so the final temperature  $T_2 = 559$  K. We note that the temperatures are already given in kelvins, so we do not need to convert the temperatures. Substituting into the expression for Charles's law yields

$$\frac{34.8 \text{ mL}}{315 \text{ K}} = \frac{V_2}{559 \text{ K}}$$

We solve for  $V_2$  by algebraically isolating the  $V_2$  variable on one side of the equation. We do this by multiplying both sides of the equation by 559 K (number and unit). When we do this, the temperature unit cancels on the left side, while the entire 559 K cancels on the right side:

$$\frac{(559 \text{ K})(34.8 \text{ mL})}{315 \text{ K}} = \frac{V_2 (559 \text{ K})}{559 \text{ K}}$$

The expression simplifies to

$$\frac{(559)(34.8 \text{ mL})}{315} = V_2$$

By multiplying and dividing the numbers, we see that the only remaining unit is mL, so our final answer is

$$V_2 = 61.8 \text{ mL}$$

Does this answer make sense? We know that as temperature increases, volume increases. Here, the temperature is increasing from 315 K to 559 K, so the volume should also increase, which it does.

### ? Exercise 6.4.3

If  $V_1 = 3.77$  L and  $T_1 = 255$  K, what is  $V_2$  if  $T_2 = 123$  K?

#### Answer

1.82 L

It is more mathematically complicated if a final temperature must be calculated because the  $T$  variable is in the denominator of Charles's law. There are several mathematical ways to work this, but perhaps the simplest way is to take the reciprocal of Charles's

law. That is, rather than write it as

$$\frac{V_1}{T_1} = \frac{V_2}{T_2}$$

write the equation as

$$\frac{T_1}{V_1} = \frac{T_2}{V_2}$$

It is still an equality and a correct form of Charles's law, but now the temperature variable is in the numerator, and the algebra required to predict a final temperature is simpler.

#### ✓ Example 6.4.4

A sample of a gas has an initial volume of 34.8 L and an initial temperature of  $-67^{\circ}\text{C}$ . What must the temperature of the gas be for its volume to be 25.0 L?

##### Solution

Here, we are looking for a final temperature, so we will use the reciprocal form of Charles's law. However, the initial temperature is given in degrees Celsius, not kelvins. We must convert the initial temperature to kelvins:

$$-67^{\circ}\text{C} + 273 = 206 \text{ K}$$

In using the gas law, we must use  $T_1 = 206 \text{ K}$  as the temperature. Substituting into the reciprocal form of Charles's law, we get

$$\frac{206 \text{ K}}{34.8 \text{ L}} = \frac{T_2}{25.0 \text{ L}}$$

Bringing the 25.0 L quantity over to the other side of the equation, we get

$$\frac{(25.0 \cancel{\text{ L}})(206 \text{ K})}{34.8 \cancel{\text{ L}}} = T_2$$

The L units cancel, so our final answer is  $T_2 = 148 \text{ K}$

This is also equal to  $-125^{\circ}\text{C}$ . As temperature decreases, volume decreases, which it does in this example.

#### ? Exercise 6.4.4

If  $V_1 = 623 \text{ mL}$ ,  $T_1 = 255^{\circ}\text{C}$ , and  $V_2 = 277 \text{ mL}$ , what is  $T_2$ ?

##### Answer

235 K, or  $-38^{\circ}\text{C}$

## Summary

- The behavior of gases can be modeled with gas laws.
- Boyle's law relates a gas's pressure and volume at constant temperature and amount.
- Charles's law relates a gas's volume and temperature at constant pressure and amount.
- In gas laws, temperatures must always be expressed in kelvins.

This page titled [6.4: Gas Laws](#) is shared under a [CC BY-NC-SA 3.0](#) license and was authored, remixed, and/or curated by [Anonymous](#) via [source content](#) that was edited to the style and standards of the LibreTexts platform.