CHAPTER OVERVIEW

13: Chemical Equilibrium

So far in this text, when presented a chemical reaction, it has been implicitly assumed that the reaction goes to completion. Indeed, previous stoichiometric calculations were based on this; when asked how much of a product is produced when so much of a reactant reacts, it was assumed that *all* of a reactant reacts. However, this is usually not the case; many reactions do not go to completion, and many chemists have to deal with that. In this chapter, we will study this phenomenon and see ways in which we can affect the extent of chemical reactions.