

Section 9: Review and Discussion Questions

Compounds with formulae of the form C_nH_{2n+2}

are often referred to as "saturated" hydrocarbons. Using Lewis structures, explain how and in what sense these molecules are "saturated."

Molecules with formulae of the form C_nH_{2n+1}

(e.g. CH_3 , C_2H_5

) are called "radicals" and are extremely reactive. Using Lewis structures, explain the reactivity of these molecules.

State and explain the experimental evidence and reasoning which shows that multiple bonds are stronger and shorter than single bonds.

Compare N_2

to H_4N_2

. Predict which bond is stronger and explain why.

Explain why the two Lewis structures for Freon 114, shown in Figure 21Figure, are identical. Draw a Lewis structures for an isomer of Freon 114, that is, another molecule with the same molecular formula as Freon 114 but a different structural formula.

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