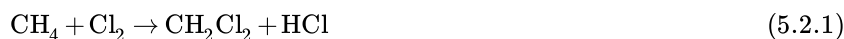


## 5.2: Balancing Chemical Equations

As noted earlier, a *balanced* chemical equation shows the same number of each kind of atom on both sides of the equation. The process of balancing chemical equations is an important exercise in chemistry and is addressed here.

Consider a simple example of balancing a chemical equation, the reaction of methane,  $\text{CH}_4$ , with elemental chlorine,  $\text{Cl}_2$ , to produce dichloromethane,  $\text{CH}_2\text{Cl}_2$ , an important laboratory solvent, and byproduct hydrogen chloride,  $\text{HCl}$ . The unbalanced chemical equation is



Inspection of this equation as it is written shows that it is not balanced because it has 4 H on the left, but just 3 on the right and 2 Cl on the left, but 3 Cl on the right. In order to balance such an equation, consider one element at a time. Carbon is already balanced, so it is best to avoid changing any of the numbers in front of the C-containing compounds. The equation can be balanced for H by putting a 2 in front of  $\text{HCl}$ :



Now everything is balanced except for Cl, of which there are 4 on the right, but just 2 on the left. Placing a 2 in front of  $\text{Cl}_2$  gives the required 4 Cls on the left:



This equation is now balanced with 1 C, 4 Hs, and 4 Cls on both the left and the right.

A crucial thing to remember in balancing a chemical equation is that *the chemical formulas must not be altered. Only the relative numbers of reactant and product species may be changed.*

Next consider the reaction of methane,  $\text{CH}_4$ , with iron oxide,  $\text{Fe}_2\text{O}_3$ , to give iron metal,  $\text{Fe}$ , carbon dioxide,  $\text{CO}_2$ , and water,  $\text{H}_2\text{O}$ . The unbalanced equation is



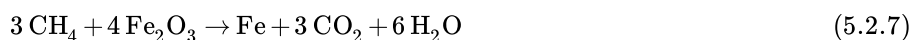
In this case it is helpful to note that  $\text{CH}_4$  is the only source of both C and H and that 4 times as many H atoms as C atoms must appear in the products. That means that for each  $\text{CO}_2$  there must be 2  $\text{H}_2\text{Os}$ . Both C and H are balanced in the following:



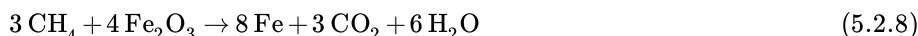
But now O is not balanced. Furthermore, the 3 Os in  $\text{Fe}_2\text{O}_3$  means that the number of O atoms must be divisible by 3, so try multiplying the three species balanced so far —  $\text{CH}_4$ ,  $\text{CO}_2$ , and  $2\text{H}_2\text{O}$  — by 3:



That gives a total of 12 O atoms on the right, 6 each in 3  $\text{CO}_2$  and 6  $\text{H}_2\text{O}$ . Taking 4 times  $\text{Fe}_2\text{O}_3$  gives 12 Os on the left:



The only species remaining to be balanced is Fe, which can be balanced by putting 8 in front of Fe on the right. The balanced equation is

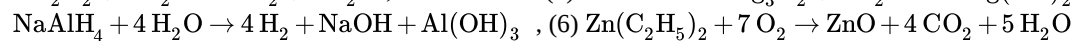
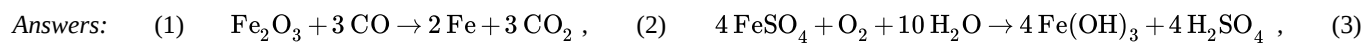
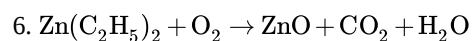


Checking the answer shows on both left and right 3 C, 8 Fe, 12 H, and 12 O demonstrating that the equation is in fact balanced.

**Exercise:** Balance the following:

1.  $\text{Fe}_2\text{O}_3 + \text{CO} \rightarrow \text{Fe} + \text{CO}_2$
2.  $\text{FeSO}_4 + \text{O}_2 + \text{H}_2\text{O} \rightarrow \text{Fe}(\text{OH})_3 + \text{H}_2\text{SO}_4$
3.  $\text{C}_2\text{H}_2 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
4.  $\text{Mg}_3\text{N}_2 + \text{M}_2\text{O} \rightarrow \text{Mg}(\text{OH})_2 + \text{NH}_3$
5.  $\text{NaAlH}_4 + \text{H}_2\text{O} \rightarrow \text{H}_2 + \text{NaOH} + \text{Al}(\text{OH})_3$





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