

17.5: AROMATICITY AND HUCKEL'S RULE

In 1931, German chemist and physicist Erich Hückel proposed a theory to help determine if a planar ring molecule would have aromatic properties. His rule states that if a cyclic, planar molecule has $4n+2$ π electrons, it is considered aromatic. This rule would come to be known as Hückel's Rule.

HUCKEL'S RULE: AROMATIC, ANTIAROMATIC, AND NONAROMATIC

Huckel's Rule is a set of algorithms that combine the number of π electrons (N) and the physical structure of the ring system to determine whether the molecule is aromatic, antiaromatic, or nonaromatic.

The number of π electrons in an **aromatic** system can be determined by the following algorithm:

$$N = 4n + 2 \quad (17.5.1)$$

where n is an integer.

The number of π electrons in an **antiaromatic** system can be determined by the following algorithm:

$$N = 4n \quad (17.5.2)$$

where n is an integer.

If a compound does not have a continuous ring of conjugated p orbitals in a planar conformation, then it is nonaromatic.

Huckel's Rule is a useful first step in evaluating the potential for a ringed molecule to be aromatic. The planar requirement of the ring may require further investigation.

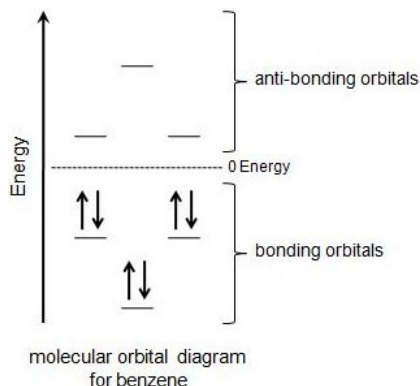
Four Criteria for Aromaticity

When deciding if a compound is aromatic, go through the following checklist. If the compound does not meet all the following criteria, it is likely not aromatic.

1. The molecule is cyclic (a ring of atoms)
2. The molecule is planar (all atoms in the molecule lie in the same plane)
3. The molecule is fully **conjugated** (p orbitals at every atom in the ring)
4. The molecule has $4n+2$ π electrons ($n=0$ or any positive integer)

WHY $4N+2$ π ELECTRONS?

According to Hückel's Molecular Orbital Theory, a compound is particularly stable if all of its bonding molecular orbitals are filled with paired electrons. This is true of aromatic compounds, meaning they are quite stable. With aromatic compounds, 2 electrons fill the lowest energy molecular orbital, and 4 electrons fill each subsequent energy level (the number of subsequent energy levels is denoted by n), leaving all bonding orbitals filled and no anti-bonding orbitals occupied. This gives a total of $4n+2$ π electrons. You can see how this works with the molecular orbital diagram for the aromatic compound, benzene, below. Benzene has 6 π electrons. Its first 2 π electrons fill the lowest energy orbital, and it has 4 π electrons remaining. These 4 fill in the orbitals of the succeeding energy level. Notice how all of its bonding orbitals are filled, but none of the anti-bonding orbitals have any electrons.



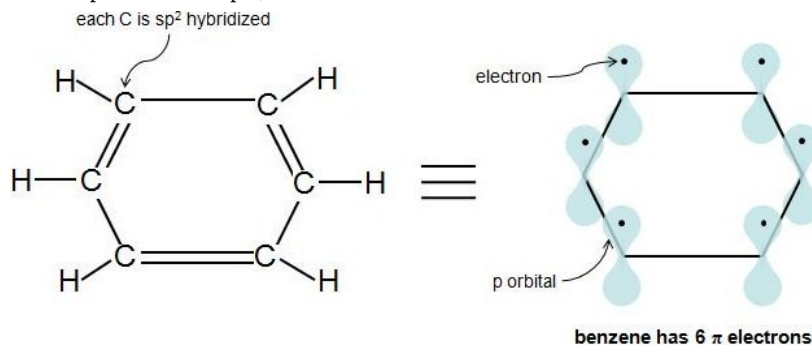
To apply the $4n+2$ rule, first count the number of π electrons in the molecule. Then, set this number equal to $4n + 2$ and solve for n . If n is 0 or any positive integer (1, 2, 3,...), the rule has been met. For example, benzene has six π electrons:

$$\begin{aligned} 4n + 2 &= 6 \\ 4n &= 4 \\ n &= 1 \end{aligned}$$

For benzene, we find that $n = 1$, which is a positive integer, so the rule is met.

HOW CAN YOU TELL WHICH ELECTRONS ARE π ELECTRONS?

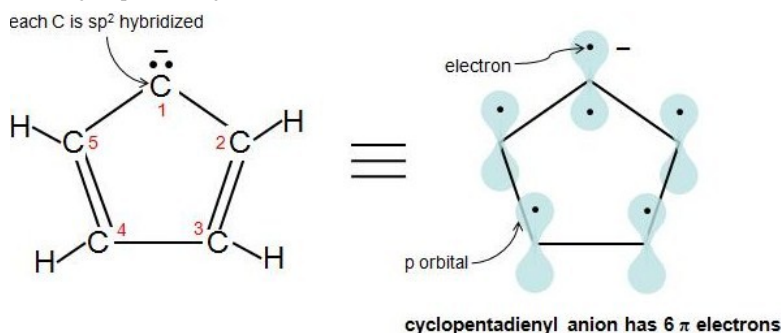
Perhaps the toughest part of Hückel's Rule is figuring out which electrons in the compound are actually π electrons. Once this is figured out, the rule is quite straightforward. π electrons lie in p orbitals and sp^2 hybridized atoms have 1 p orbital each. So if every carbon atom in the cyclic compound is sp^2 hybridized, this means the molecule is fully conjugated (has 1 p orbital at each atom), and the electrons in these p orbitals are the π electrons. A simple way to know if an atom is sp^2 hybridized is to see if it has 3 attached atoms and no lone pairs of electrons. This [video](#) provides a very nice tutorial on how to determine an atom's hybridization. In a cyclic hydrocarbon compound with alternating single and double bonds, each carbon is attached to 1 hydrogen and 2 other carbons. Therefore, each carbon is sp^2 hybridized and has a p orbital. Let's look at our previous example, benzene:



Each double bond (π bond) always contributes 2 π electrons. Benzene has 3 double bonds, so it has 6 π electrons.

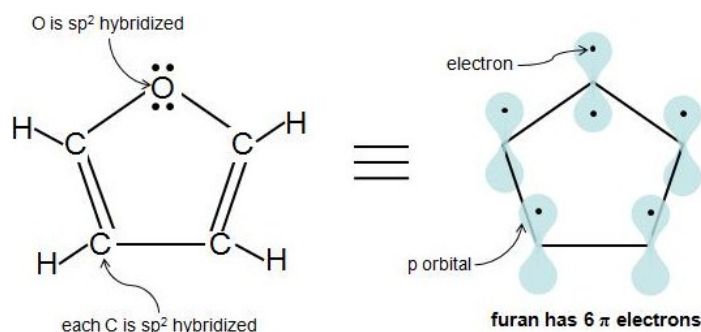
AROMATIC IONS

Hückel's Rule also applies to ions. As long as a compound has $4n+2$ π electrons, it does not matter if the molecule is neutral or has a charge. For example, cyclopentadienyl anion is an aromatic ion. How do we know that it is fully conjugated? That is, how do we know that each atom in this molecule has 1 p orbital? Let's look at the following figure. Carbons 2-5 are sp^2 hybridized because they have 3 attached atoms and have no lone electron pairs. What about carbon 1? Another simple rule to determine if an atom is sp^2 hybridized is if an atom has 1 or more lone pairs and is attached to an sp^2 hybridized atom, then that atom is sp^2 hybridized also. This [video](#) explains the rule very clearly. Therefore, carbon 1 has a p orbital. Cyclopentadienyl anion has 6 π electrons and fulfills the $4n+2$ rule.



HETEROCYCLIC AROMATIC COMPOUNDS

So far, you have encountered many carbon homocyclic rings, but compounds with elements other than carbon in the ring can also be aromatic, as long as they fulfill the criteria for aromaticity. These molecules are called heterocyclic compounds because they contain 1 or more different atoms other than carbon in the ring. A common example is furan, which contains an oxygen atom. We know that all carbons in furan are sp^2 hybridized. But is the oxygen atom sp^2 hybridized? The oxygen has at least 1 lone electron pair and is attached to an sp^2 hybridized atom, so it is sp^2 hybridized as well. Notice how oxygen has 2 lone pairs of electrons. How many of those electrons are π electrons? An sp^2 hybridized atom only has 1 p orbital, which can only hold 2 electrons, so we know that 1 electron pair is in the p orbital, while the other pair is in an sp^2 orbital. So, only 1 of oxygen's 2 lone electron pairs are π electrons. Furan has 6 π electrons and fulfills the $4n+2$ rule.



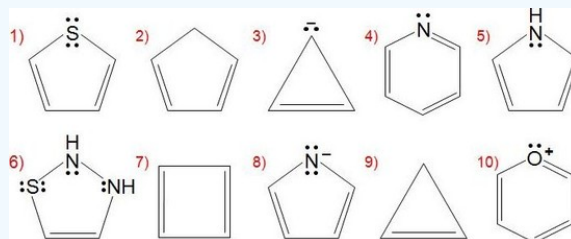
A Common Misconception

A very common misconception is that hybridization can be used to predict the geometry, or that hybridization somehow involves an energy cost associated with 'promoting' electrons into the hybrid orbitals. **This is entirely wrong.** Hybridization is always determined by geometry. You can only assign hybridization states to an atom if you already know its geometry, based on some experimental or theoretical evidence. The geometry of the oxygen in furan is trigonal planar and therefore the hybridization must be sp^2 .

The *specific* rule is that if you have an sp^2 conjugated system, the lone pair will be involved **if it makes the system more stable**. In this case, conferring Hückel $4n + 2$ aromaticity. For furan with *two* lone pairs on the oxygen atom, if we count electrons from the carbon atoms, we have 4 (one per carbon). So adding two electrons from one of the lone pairs will give $6 = 4(1) + 2$, so Hückel rule is applicable and furan is aromatic.

Exercise

3. Using the criteria for aromaticity, determine if the following molecules are aromatic:



Answer

3.

Cpd 1: Aromatic - only 1 of S's lone pairs counts as π electrons, so there are 6 π electrons, $n=1$ Not aromatic - not fully conjugated, top C is sp^3 hybridized

Cpd 2: Not aromatic - top C is sp^2 hybridized, but there are 4 π electrons, $n=1/2$

Cpd 3: Aromatic - N is using its 1 p orbital for the electrons in the double bond, so its lone pair of electrons are not π electrons, there are 6 π electrons, $n=1$

Cpd 4: Aromatic - there are 6 π electrons, $n=1$

Cpd 5: Aromatic - there are 6 π electrons, $n=1$ because the N assumes sp^2 hybridization.

Cpd 6: Not aromatic - all atoms are sp^2 hybridized, but only 1 of S's lone pairs counts as π electrons, so there 8 π electrons, $n=1.5$

Cpd 7: Not aromatic - there are 4 π electrons, $n=1/2$

Cpd 8: Aromatic - only 1 of N's lone pairs counts as π electrons, so there are 6 π electrons, $n=1$

Cpd 9: Not aromatic - not fully conjugated, top C is sp^3 hybridized

Cpd 10: Aromatic - O is using its 1 p orbital for the elections in the double bond, so its lone pair of electrons are not π electrons, there are 6 π electrons, $n=1$

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