

6.10: ABSOLUTE AND RELATIVE CONFIGURATION - THE DISTINCTION

Learning Objective

- compare and contrast absolute configuration with relative configuration

Absolute Configuration

The absolute configuration at a chiral center in a molecule is a time-independent and unambiguous symbolic description of the spatial arrangement of ligands (groups) bonded to the chiral center.



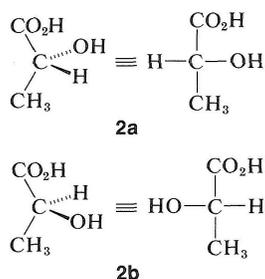
The chiral centers in 1 and 2 bear the same ligands: a,b,d, and e. However, 1 and 2 are not superimposable on each other, meaning that the arrangement of ligands around the chiral center in 1 and in 2 is different. 1 and 2 are mirror images of each other, meaning that the arrangement of ligands around the chiral center in 1 is the exact opposite of that in 2. Chiral centers in 1 and 2 are said to have opposite absolute configurations.

According to R,S convention, if the absolute configuration at the chiral center in 1 is R, that at the chiral center in 2 is S or vice versa.

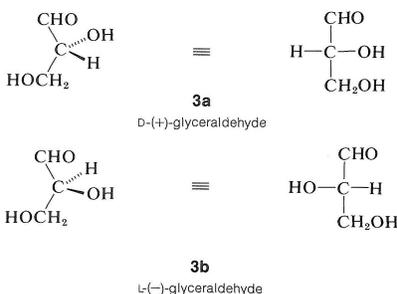
RELATIVE CONFIGURATION

The relative configuration is the experimentally determined relationship between two enantiomers even though we may not know the absolute configuration. The sign of rotation of plane-polarized light by an enantiomer is not easily related to its configuration. This is true even for substances with very similar structures. Thus, given lactic acid, $\text{CH}_3\text{CHOHCO}_2\text{H}$, with a specific rotation $+3.82^\circ$, and methyl lactate, $\text{CH}_3\text{CHOHCO}_2\text{CH}_3$, with a specific rotation -8.25° , we cannot tell from the rotation alone whether the acid and ester have the same or a different arrangement of groups about the chiral center. Their relative configurations have to be obtained by other means.

If we convert (+)-lactic acid into its methyl ester, we can be reasonably certain that the ester will be related in configuration to the acid, because esterification should not affect the configuration about the chiral carbon atom. It happens that the methyl ester so obtained is levorotatory, so we know that (+)-lactic acid and (–)-methyl lactate have the same relative configuration at the asymmetric carbon, even if they possess opposite signs of optical rotation. However, we still do not know the absolute configuration; that is, we are unable to tell which of the two possible configurations of lactic acid, *2a* or *2b*, corresponds to the dextro or (+)-acid and which to the levo or (–)-acid:



Until 1956, the absolute configuration of no optically active compound was known. Instead, configurations were assigned relative to a standard, *glyceraldehyde*, which originally was chosen by E. Fischer (around 1885) for the purpose of correlating the configuration of carbohydrates. Fischer arbitrarily assigned the configuration *3a* to dextrorotatory glyceraldehyde, which was known as *D*-(+)-glyceraldehyde. The levorotatory enantiomer, *3b*, is designated as *L*-(–)-glyceraldehyde. (If you are unsure of the terminology *D* and *L*, or of the rules for writing Fischer projection formulas, review [Sections 5-3C](#) and [5-4](#).)



The configurations of many compounds besides sugars now have been related to glyceraldehyde, including α -amino acids, terpenes, steroids, and other biochemically important substances. Compounds whose configurations are related to *D*-(+)-glyceraldehyde are said to belong to the *D* series, and those related to *L*-(-)-glyceraldehyde belong to the *L* series.

At the time the choice of absolute configuration for glyceraldehyde was made, there was no way of knowing whether the configuration of (+)-glyceraldehyde was in reality 3a or 3b. However, the choice had a 50% chance of being correct, and we now know that 3a, the *D* configuration, is in fact the correct configuration of (+)-glyceraldehyde. This was established through use of a special x-ray crystallographic technique, which permitted determination of the absolute disposition of the atoms in space of sodium rubidium (+)-tartrate. The configuration of (+)-tartaric acid (Section 5-5) previously had been shown by chemical means to be opposite to that of (+)-glyceraldehyde. Consequently the absolute configuration of any compound now is known once it has been correlated directly or indirectly with glyceraldehyde. For example,

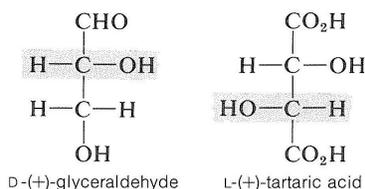
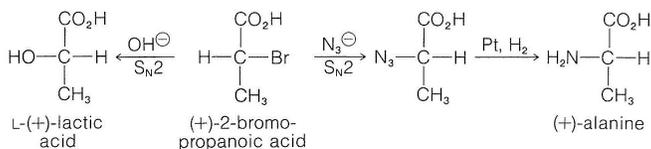
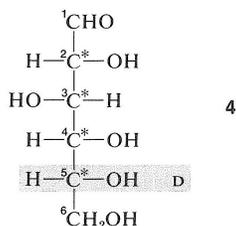


Figure below whereby the configuration of (+)-lactic acid is related to the amino acid (+)-alanine. Because (+)-lactic acid has been related to *L*-(-)-glyceraldehyde, it follows that the absolute configurations are *L*-(+)-lactic acid and *L*-(+)-alanine.



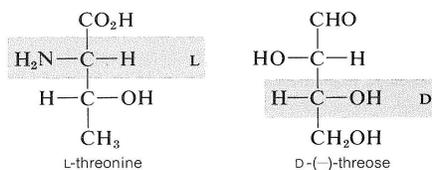
Chemical transformation showing how the configuration of natural (+)-alanine has been related to *L*-(+)-lactic acid and hence to *L*-(-)-glyceraldehyde. The transformations shown involve two $\text{S}_\text{N}2$ reactions, each of which is stereospecific and inverts the configuration (Section 8-5). Reduction of the azide group leaves the configuration unchanged.

When there are several chiral carbons in a molecule, the configuration at one center usually is related directly or indirectly to glyceraldehyde, and the configurations at the other centers are determined relative to the first. Thus in the aldehyde form of the important sugar, (+)-glucose, there are *four* chiral centers, and so there are $2^4 = 16$ possible stereoisomers. The projection formula of the isomer that corresponds to the aldehyde form of natural glucose is 4. By convention for sugars, the configuration of the *highest-numbered chiral carbon* is referred to glyceraldehyde to determine the overall configuration of the molecule. For glucose, this atom is C_5 , next to the CH_2OH group, and has the hydroxyl group on the right. Therefore, naturally occurring glucose, which has a (+) rotation, belongs to the *D* series and is properly called *D*-(+)-glucose:



However, the configurations of α -amino acids possessing more than one chiral carbon are determined by the *lowest*-numbered chiral carbon, which is the carbon *alpha* to the carboxyl group. Thus, even though the natural α -amino acid, threonine, has exactly the same kind

of arrangement of substituents as the natural sugar, threose, threonine by the amino-acid convention belongs to the *L*-series, whereas threose by the sugar convention belongs to the *D*-series:



A serious ambiguity arises for compounds such as the active tartaric acids. If the amino-acid convention is used, (+)-tartaric acid falls in the *D* series; by the sugar convention, it has the *L* configuration. One way out of this dilemma is to use the subscripts *s* and *g* to denote the amino-acid or carbohydrate conventions, respectively. Then the absolute configuration of (+)-tartaric acid can be designated as either *D_s*-(+)-tartaric acid or *L_g*-(+)-tartaric acid.

CONTRIBUTORS AND ATTRIBUTIONS

- ◦ [Gamini Gunawardena](#) from the [OChemPal](#) site ([Utah Valley University](#))
- John D. Robert and Marjorie C. Caserio (1977) *Basic Principles of Organic Chemistry, second edition*. W. A. Benjamin, Inc. , Menlo Park, CA. ISBN 0-8053-8329-8. This content is copyrighted under the following conditions, "You are granted permission for individual, educational, research and non-commercial reproduction, distribution, display and performance of this work in any format."

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