

## 19.10: Relative Enthalpies Can Be Determined from Heat Capacity Data and Heats of Transition

It should be stressed that there **are no absolute enthalpies**. All that is properly defined are differences in enthalpy  $\Delta H$  and these are only defined for **processes**

When dealing with enthalpies

### Define the Process

For example for the process of:

Process 1: **heating ice from  $-20^{\circ}\text{C}$  to  $-0^{\circ}\text{C}$**

We could write  $\Delta H_1 = \int C_{p,\text{ice}} dT$  from  $T=253\text{K}$  to  $T=273\text{K} = H(273) - H(253)$ . But before moving beyond the melting point first a different process needs to take place, that of

Process 2: **melting**

This gives us  $\Delta_{\text{fus}}H = H_{\text{liquid}} - H_{\text{solid}}$  (both at  $273\text{K}$ !). When we heat the liquid water further to say  $+20^{\circ}\text{C}$  we would have to integrate over the heat capacity of the liquid.

Process 3: **heating water from  $0^{\circ}\text{C}$  to  $+20^{\circ}\text{C}$**

We could write

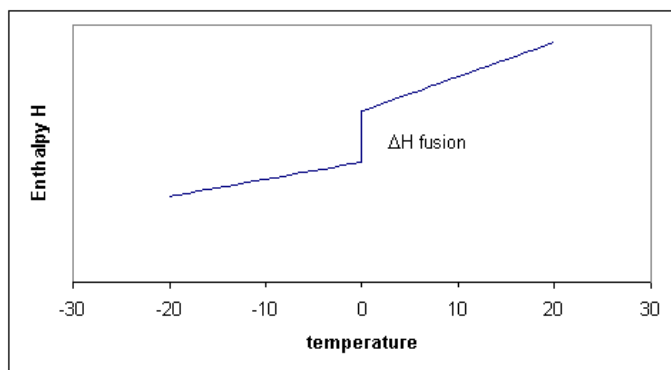
$$\Delta H_3 = \int C_{p,\text{water}} dT$$

from  $T=273\text{K}$  to  $T=293\text{K} = H(293) - H(273)$ .

The total change in enthalpy between  $-20$  and  $+20$  would be the sum of the three enthalpy changes.

$$\Delta H_{\text{total process}} = \Delta H_1 + \Delta_{\text{fus}}H + \Delta H_3$$

Of course we could consider doing the same calculation for any temperature between  $-20$  and  $+20$  and summarize all our results in a graph. The three processes can thus schematically be shown in Figure 19.10.1 .



**Figure 19.10.1** : Schematic enthalpy function showing the jump at the melting point

Notice that the slopes (i.e. the heat capacities!) before and after the melting point differ. The slope for the liquid is a little steeper because the liquid has more degrees of freedom and therefore the heat capacity of the liquid tends to be higher than of the solid. In the figure the enthalpy curves are shown as straight lines. This would be the case if the heat capacities are *constant* over the temperature interval. Although  $C_p$  is typically a 'slow' or 'weak' function of temperature it usually does change a bit, which means that the straight lines for  $H$  become curves.

Although  $C_p$  is typically a 'slow' or 'weak' function of temperature and is well approximated as a constant.

Notice that for process two, the temperature is constant, that means that  $\Delta T$  or  $dT$  is zero, but  $\Delta H$  is finite, consequently  $\Delta H/\Delta T$  is infinitely large. Taking the limit for  $\Delta T$  going to zero, we get a derivative:

$$\left( \frac{\partial H}{\partial T} \right)_p = C_p$$

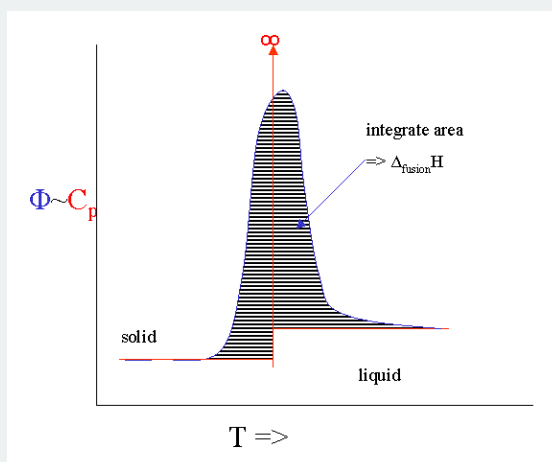
This derivative, the heat capacity must undergo a singularity: the slope is infinitely large (i.e., the  $H$  curve goes straight up). When there are more phase transitions, more discontinuities in  $H$  and singularities in  $C_p$  result (Figure 19.10.1). Note that  $H(T) - H(0)$ , not  $H(T)$  is plotted to avoid the question what the absolute enthalpy is.

### Scanning calorimetry

There is technique that allows us to measure the heat capacity as a function of temperature fairly directly. It is called Differential Scanning Calorimetry (DSC). You put a sample in a little pan and put the pan plus an empty reference pan in the calorimeter. The instrument heats up both pans with a constant heating rate. Both pans get hotter by **conduction**, but the heat capacity of the filled pan is obviously bigger. This means that the heat flow into the sample pan must be a bit bigger than into the empty one. This differential heat flow induces a tiny temperature difference  $\Delta T$  between the two pans that can be measured. This temperature difference is proportional to the heat flow difference which is proportional to the heat capacity difference.

$$\Delta T \propto \Delta C_p = C_p^{\text{sample}} - C_p^{\text{reference}} \quad (\text{if the pans cancel})$$

However, there are number of serious broadening issues with the technique. If you melt something you will never get to see the infinite singularity of the heat capacity. Instead it broadens out into a peak. If you integrate the peak you get the  $\Delta_{\text{fusion}} H$  and the onset is calibrated to give you the melting point.



The ideal heat capacity signal and its broadened DSC signal

It is even possible to heat the sample with a rate that fluctuates with a little sine wave. This "Modulated DSC" version can even give you the (small) difference in  $C_p$  before and after the melting event.

19.10: Relative Enthalpies Can Be Determined from Heat Capacity Data and Heats of Transition is shared under a CC BY-NC-SA 4.0 license and was authored, remixed, and/or curated by LibreTexts.