

CHAPTER OVERVIEW

26: Chemical Equilibrium

- 26.1: Equilibrium Results when Gibbs Energy is Minimized
- 26.2: An Equilibrium Constant is a Function of Temperature Only
- 26.3: Standard Gibbs Energies of Formation Can Be Used to Calculate Equilibrium Constants
- 26.4: Gibbs Energy of a Reaction vs. Extent of Reaction is a Minimum at Equilibrium
- 26.5: Reaction Quotient and Equilibrium Constant Ratio Determines Reaction Direction
- 26.6: The Sign of ΔG and not ΔG° Determines the Direction of Reaction Spontaneity
- 26.7: The van 't Hoff Equation
- 26.8: Equilibrium Constants in Terms of Partition Functions
- 26.9: Molecular Partition Functions and Related Thermodynamic Data Are Extensively Tabulated
- 26.10: Real Gases Are Expressed in Terms of Partial Fugacities
- 26.11: Thermodynamic Equilibrium Constants Are Expressed in Terms of Activities
- 26.12: Activities are Important for Ionic Species
- 26.13: Homework Problems

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