

13.E: Work, Energy, and Energy Resources (Exercise)

Conceptual Questions

13.1: Temperature

1. What does it mean to say that two systems are in thermal equilibrium?
2. Give an example of a physical property that varies with temperature and describe how it is used to measure temperature.
3. When a cold alcohol thermometer is placed in a hot liquid, the column of alcohol goes down slightly before going up. Explain why.
4. If you add boiling water to a cup at room temperature, what would you expect the final equilibrium temperature of the unit to be? You will need to include the surroundings as part of the system. Consider the zeroth law of thermodynamics.

13.2: Thermal Expansion of Solids and Liquids

5. Thermal stresses caused by uneven cooling can easily break glass cookware. Explain why Pyrex®, a glass with a small coefficient of linear expansion, is less susceptible.
6. Water expands significantly when it freezes: a volume increase of about 9% occurs. As a result of this expansion and because of the formation and growth of crystals as water freezes, anywhere from 10% to 30% of biological cells are burst when animal or plant material is frozen. Discuss the implications of this cell damage for the prospect of preserving human bodies by freezing so that they can be thawed at some future date when it is hoped that all diseases are curable.
7. One method of getting a tight fit, say of a metal peg in a hole in a metal block, is to manufacture the peg slightly larger than the hole. The peg is then inserted when at a different temperature than the block. Should the block be hotter or colder than the peg during insertion? Explain your answer.
8. Does it really help to run hot water over a tight metal lid on a glass jar before trying to open it? Explain your answer.
9. Liquids and solids expand with increasing temperature, because the kinetic energy of a body's atoms and molecules increases. Explain why some materials *shrink* with increasing temperature.

13.3: The Ideal Gas Law

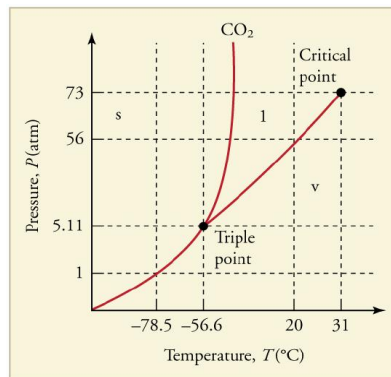
10. Find out the human population of Earth. Is there a mole of people inhabiting Earth? If the average mass of a person is 60 kg, calculate the mass of a mole of people. How does the mass of a mole of people compare with the mass of Earth?
11. Under what circumstances would you expect a gas to behave significantly differently than predicted by the ideal gas law?
12. A constant-volume gas thermometer contains a fixed amount of gas. What property of the gas is measured to indicate its temperature?

13.4: Kinetic Theory: Atomic and Molecular Explanation of Pressure and Temperature

13. How is momentum related to the pressure exerted by a gas? Explain on the atomic and molecular level, considering the behavior of atoms and molecules.

13.5: Phase Changes

14. A pressure cooker contains water and steam in equilibrium at a pressure greater than atmospheric pressure. How does this greater pressure increase cooking speed?
15. Why does condensation form most rapidly on the coldest object in a room—for example, on a glass of ice water?
16. What is the vapor pressure of solid carbon dioxide (dry ice) at -78.5°C ?



The phase diagram for carbon dioxide. The axes are nonlinear, and the graph is not to scale. Dry ice is solid carbon dioxide and has a sublimation temperature of -78.5°C .

17. Can carbon dioxide be liquefied at room temperature (20°C)? If so, how? If not, why not? (See Figure.)
18. Oxygen cannot be liquefied at room temperature by placing it under a large enough pressure to force its molecules together. Explain why this is.
19. What is the distinction between gas and vapor?

13.6: Humidity, Evaporation, and Boiling

20. Because humidity depends only on water's vapor pressure and temperature, are the saturation vapor densities listed in Table valid in an atmosphere of helium at a pressure of $1.01 \times 10^5 \text{ N/m}^2$, rather than air? Are those values affected by altitude on Earth?
21. Why does a beaker of 40.0°C water placed in a vacuum chamber start to boil as the chamber is evacuated (air is pumped out of the chamber)? At what pressure does the boiling begin? Would food cook any faster in such a beaker?
22. Why does rubbing alcohol evaporate much more rapidly than water at STP (standard temperature and pressure)?

Problems & Exercises

13.1: Temperature

23. What is the Fahrenheit temperature of a person with a 39.0°C fever?

Solution

102°F

24. Frost damage to most plants occurs at temperatures of 28.0°F or lower. What is this temperature on the Kelvin scale?
25. To conserve energy, room temperatures are kept at 68.0°F in the winter and 78.0°F in the summer. What are these temperatures on the Celsius scale?

Solution

20.0°C and 25.6°C

26. A tungsten light bulb filament may operate at 2900 K. What is its Fahrenheit temperature? What is this on the Celsius scale?
27. The surface temperature of the Sun is about 5750 K. What is this temperature on the Fahrenheit scale?

Solution

9890°F

28. One of the hottest temperatures ever recorded on the surface of Earth was 134°F in Death Valley, CA. What is this temperature in Celsius degrees? What is this temperature in Kelvin?
29. (a) Suppose a cold front blows into your locale and drops the temperature by 40.0 Fahrenheit degrees. How many degrees Celsius does the temperature decrease when there is a 40.0°F decrease in temperature?

(b) Show that any change in temperature in Fahrenheit degrees is nine-fifths the change in Celsius degrees.

Solution

(a) 22.2°C

$$\begin{aligned} \text{(b)} \quad \Delta T(^{\circ}\text{F}) &= T_2(^{\circ}\text{F}) - T_1(^{\circ}\text{F}) \\ &= \frac{9}{5}T_2(^{\circ}\text{C}) + 32.0^{\circ} - \left(\frac{9}{5}T_1(^{\circ}\text{C}) + 32.0^{\circ}\right) \\ &= \frac{9}{5}(T_2(^{\circ}\text{C}) - T_1(^{\circ}\text{C})) = \frac{9}{5}\Delta T(^{\circ}\text{C}) \end{aligned}$$

30. (a) At what temperature do the Fahrenheit and Celsius scales have the same numerical value?

(b) At what temperature do the Fahrenheit and Kelvin scales have the same numerical value?

13.2: Thermal Expansion of Solids and Liquids

31. The height of the Washington Monument is measured to be 170 m on a day when the temperature is 35.0°C . What will its height be on a day when the temperature falls to -10.0°C ? Although the monument is made of limestone, assume that its thermal coefficient of expansion is the same as marble's.

Solution

169.95 m

32. How much taller does the Eiffel Tower become at the end of a day when the temperature has increased by 15°C ? Its original height is 321 m and you can assume it is made of steel.

33. What is the change in length of a 3.00-cm-long column of mercury if its temperature changes from 37.0°C to 40.0°C , assuming the mercury is unconstrained?

Solution

$5.4 \times 10^{-6} \text{ m}$

34. How large an expansion gap should be left between steel railroad rails if they may reach a maximum temperature 35.0°C greater than when they were laid? Their original length is 10.0 m.

35. You are looking to purchase a small piece of land in Hong Kong. The price is "only" \$60,000 per square meter! The land title says the dimensions are $20 \text{ m} \times 30 \text{ m}$. By how much would the total price change if you measured the parcel with a steel tape measure on a day when the temperature was 20°C above normal?

Solution

Because the area gets smaller, the price of the land DECREASES by \$17,000.

36. Global warming will produce rising sea levels partly due to melting ice caps but also due to the expansion of water as average ocean temperatures rise. To get some idea of the size of this effect, calculate the change in length of a column of water 1.00 km high for a temperature increase of 1.00°C . Note that this calculation is only approximate because ocean warming is not uniform with depth.

37. Show that 60.0 L of gasoline originally at 15.0°C size $12\{ "15" \text{ " " } 0^{\circ}\text{C} \} \{ \}$ will expand to 61.1 L when it warms to 35.0°C , size $12\{ "35" \text{ " " } 0^{\circ}\text{C}, \}$ as claimed in Example.

Solution

$$\begin{aligned} V &= V_0 + \Delta V = V_0(1 + \beta\Delta T) \\ &= (60.00 \text{ L})[1 + (950 \times 10^{-6} / ^{\circ}\text{C})(35.0^{\circ}\text{C} - 15.0^{\circ}\text{C})] \\ &= 61.1 \text{ L} \end{aligned}$$

38. (a) Suppose a meter stick made of steel and one made of invar (an alloy of iron and nickel) are the same length at 0°C . What is their difference in length at 22.0°C ?

(b) Repeat the calculation for two 30.0-m-long surveyor's tapes.

39. (a) If a 500-mL glass beaker is filled to the brim with ethyl alcohol at a temperature of 5.00°C , how much will overflow when its temperature reaches 22.0°C ?

(b) How much less water would overflow under the same conditions?

Solution

- (a) 9.35 mL
(b) 7.56 mL

40. Most automobiles have a coolant reservoir to catch radiator fluid that may overflow when the engine is hot. A radiator is made of copper and is filled to its 16.0-L capacity when at 10.0°C . What volume of radiator fluid will overflow when the radiator and fluid reach their 95.0°C operating temperature, given that the fluid's volume coefficient of expansion is $\beta = 400 \times 10^{-6}/^{\circ}\text{C}$? Note that this coefficient is approximate, because most car radiators have operating temperatures of greater than 95.0°C .

41. A physicist makes a cup of instant coffee and notices that, as the coffee cools, its level drops 3.00 mm in the glass cup. Show that this decrease cannot be due to thermal contraction by calculating the decrease in level if the 350cm^3 of coffee is in a 7.00-cm-diameter cup and decreases in temperature from 95.0°C to 45.0°C . (Most of the drop in level is actually due to escaping bubbles of air.)

Solution

0.832 mm

42. (a) The density of water at 0°C is very nearly $1000\text{kg}/\text{m}^3$ (it is actually $999.84\text{kg}/\text{m}^3$), whereas the density of ice at 0°C is $917\text{kg}/\text{m}^3$. Calculate the pressure necessary to keep ice from expanding when it freezes, neglecting the effect such a large pressure would have on the freezing temperature. (This problem gives you only an indication of how large the forces associated with freezing water might be.)

(b) What are the implications of this result for biological cells that are frozen?

43. Show that $\beta \approx 3\alpha$, by calculating the change in volume ΔV of a cube with sides of length L .

Solution

We know how the length changes with temperature: $\Delta L = \alpha L_0 \Delta T$. Also we know that the volume of a cube is related to its length by $V = L^3$, so the final volume is then $V = V_0 + \Delta V = (L_0 + \Delta L)^3$. Substituting for ΔL gives

$$V = (L_0 + \alpha L_0 \Delta T)^3 = L_0^3 (1 + \alpha \Delta T)^3.$$

Now, because $\alpha \Delta T$ is small, we can use the binomial expansion:

$$V \approx L_0^3 (1 + 3\alpha \Delta T) = L_0^3 + 3\alpha L_0^3 \Delta T.$$

So writing the length terms in terms of volumes gives $V = V_0 + \Delta V \approx V_0 + 3\alpha V_0 \Delta T$, and so

$$\Delta V = \beta V_0 \Delta T \approx 3\alpha V_0 \Delta T, \text{ or } \beta \approx 3\alpha.$$

13.3: The Ideal Gas Law

44. The gauge pressure in your car tires is $2.50 \times 10^5 \text{N}/\text{m}^2$ at a temperature of 35.0°C when you drive it onto a ferry boat to Alaska. What is their gauge pressure later, when their temperature has dropped to -40.0°C ?

Solution

1.623 atm

45. Convert an absolute pressure of $7.00 \times 10^5 \text{N}/\text{m}^2$ to gauge pressure in lb/in^2 . (This value was stated to be just less than $90.0\text{lb}/\text{in}^2$ in Example. Is it?)

46. Suppose a gas-filled incandescent light bulb is manufactured so that the gas inside the bulb is at atmospheric pressure when the bulb has a temperature of 20.0°C .

(a) Find the gauge pressure inside such a bulb when it is hot, assuming its average temperature is 60.0°C (an approximation) and neglecting any change in volume due to thermal expansion or gas leaks.

(b) The actual final pressure for the light bulb will be less than calculated in part (a) because the glass bulb will expand. What will the actual final pressure be, taking this into account? Is this a negligible difference?

Solution

(a) 0.136 atm

(b) 0.135 atm. The difference between this value and the value from part (a) is negligible.

47. Large helium-filled balloons are used to lift scientific equipment to high altitudes.

(a) What is the pressure inside such a balloon if it starts out at sea level with a temperature of 10.0°C and rises to an altitude where its volume is twenty times the original volume and its temperature is -50.0°C ?

(b) What is the gauge pressure? (Assume atmospheric pressure is constant.)

48. Confirm that the units of nRT are those of energy for each value of R :

(a) $8.31\text{ J/mol} \cdot \text{K}$,

(b) $1.99\text{ cal/mol} \cdot \text{K}$, and

(c) $0.0821\text{ L} \cdot \text{atm/mol} \cdot \text{K}$.

Solution

$$(a) nRT = (\text{mol})(\text{J/mol} \cdot \text{K})(\text{K}) = \text{J}$$

$$(b) nRT = (\text{mol})(\text{cal/mol} \cdot \text{K})(\text{K}) = \text{cal}$$

$$(c) nRT = (\text{mol})(\text{L} \cdot \text{atm/mol} \cdot \text{K})(\text{K})$$

$$= \text{L} \cdot \text{atm} = (\text{m}^3)(\text{N/m}^2)$$

$$= \text{N} \cdot \text{m} = \text{J}$$

49. In the text, it was shown that $N/V = 2.68 \times 10^{25} \text{ m}^{-3}$ for gas at STP.

(a) Show that this quantity is equivalent to $N/V = 2.68 \times 10^{19} \text{ cm}^{-3}$, as stated.

(b) About how many atoms are there in one μm^3 (a cubic micrometer) at STP?

(c) What does your answer to part (b) imply about the separation of atoms and molecules?

50. Calculate the number of moles in the 2.00-L volume of air in the lungs of the average person. Note that the air is at 37.0°C (body temperature).

Solution

$$7.86 \times 10^{-2} \text{ mol}$$

51. An airplane passenger has 100 cm^3 of air in his stomach just before the plane takes off from a sea-level airport. What volume will the air have at cruising altitude if cabin pressure drops to $7.50 \times 10^4 \text{ N/m}^2$?

52. (a) What is the volume (in km^3) of Avogadro's number of sand grains if each grain is a cube and has sides that are 1.0 mm long? (b) How many kilometers of beaches in length would this cover if the beach averages 100 m in width and 10.0 m in depth? Neglect air spaces between grains.

Solution

$$(a) 6.02 \times 10^5 \text{ km}^3$$

$$(b) 6.02 \times 10^8 \text{ km}$$

53. An expensive vacuum system can achieve a pressure as low as $1.00 \times 10^{-7} \text{ N/m}^2$ at 20°C . How many atoms are there in a cubic centimeter at this pressure and temperature?

54. The number density of gas atoms at a certain location in the space above our planet is about $1.00 \times 10^{11} \text{ m}^{-3}$, and the pressure is $2.75 \times 10^{-10} \text{ N/m}^2$ in this space. What is the temperature there?

Solution

$$-73.9^{\circ}\text{C}$$

55. A bicycle tire has a pressure of $7.00 \times 10^5 \text{ N/m}^2$ at a temperature of 18.0°C and contains 2.00 L of gas. What will its pressure be if you let out an amount of air that has a volume of 100 cm^3 at atmospheric pressure? Assume tire temperature and volume remain constant.

56. A high-pressure gas cylinder contains 50.0 L of toxic gas at a pressure of $1.40 \times 10^7 \text{ N/m}^2$ and a temperature of 25.0°C . Its valve leaks after the cylinder is dropped. The cylinder is cooled to dry ice temperature (-78.5°C) to reduce the leak rate and pressure so that it can be safely repaired.

- (a) What is the final pressure in the tank, assuming a negligible amount of gas leaks while being cooled and that there is no phase change?
- (b) What is the final pressure if one-tenth of the gas escapes?
- (c) To what temperature must the tank be cooled to reduce the pressure to 1.00 atm (assuming the gas does not change phase and that there is no leakage during cooling)?
- (d) Does cooling the tank appear to be a practical solution?

Solution

- (a) $9.14 \times 10^6 \text{ N/m}^2$
- (b) $8.23 \times 10^6 \text{ N/m}^2$
- (c) 2.16 K
- (d) No. The final temperature needed is much too low to be easily achieved for a large object.

57. Find the number of moles in 2.00 L of gas at 35.0°C and under $7.41 \times 10^7 \text{ N/m}^2$ of pressure.

58. Calculate the depth to which Avogadro's number of table tennis balls would cover Earth. Each ball has a diameter of 3.75 cm. Assume the space between balls adds an extra 25.0% to their volume and assume they are not crushed by their own weight.

Solution

41 km

- 59. (a) What is the gauge pressure in a 25.0°C car tire containing 3.60 mol of gas in a 30.0 L volume?
- (b) What will its gauge pressure be if you add 1.00 L of gas originally at atmospheric pressure and 25.0°C ? Assume the temperature returns to 25.0°C and the volume remains constant.
- 60. (a) In the deep space between galaxies, the density of atoms is as low as 10^6 atoms/m^3 , and the temperature is a frigid 2.7 K. What is the pressure?
- (b) What volume (in m^3) is occupied by 1 mol of gas?
- (c) If this volume is a cube, what is the length of its sides in kilometers?

Solution

- (a) $3.7 \times 10^{-17} \text{ Pa}$
- (b) $6.0 \times 10^{17} \text{ m}^3$
- (c) $8.4 \times 10^2 \text{ km}$

13.4: Kinetic Theory: Atomic and Molecular Explanation of Pressure and Temperature

61. Some incandescent light bulbs are filled with argon gas. What is v_{rms} for argon atoms near the filament, assuming their temperature is 2500 K?

Solution

$1.25 \times 10^3 \text{ m/s}$

62. Average atomic and molecular speeds (v_{rms}) are large, even at low temperatures. What is v_{rms} for helium atoms at 5.00 K, just one degree above helium's liquefaction temperature?

- 63. (a) What is the average kinetic energy in joules of hydrogen atoms on the 5500°C surface of the Sun?
- (b) What is the average kinetic energy of helium atoms in a region of the solar corona where the temperature is $6.00 \times 10^5 \text{ K}$?

Solution

- (a) $1.20 \times 10^{-19} \text{ J}$
- (b) $1.24 \times 10^{-17} \text{ J}$

64. The escape velocity of any object from Earth is 11.2 km/s.

- (a) Express this speed in m/s and km/h.

(b) At what temperature would oxygen molecules (molecular mass is equal to 32.0 g/mol) have an average velocity v_{rms} equal to Earth's escape velocity of 11.1 km/s?

65. The escape velocity from the Moon is much smaller than from Earth and is only 2.38 km/s. At what temperature would hydrogen molecules (molecular mass is equal to 2.016 g/mol) have an average velocity v_{rms} equal to the Moon's escape velocity?

Solution

458 K

66. Nuclear fusion, the energy source of the Sun, hydrogen bombs, and fusion reactors, occurs much more readily when the average kinetic energy of the atoms is high—that is, at high temperatures. Suppose you want the atoms in your fusion experiment to have average kinetic energies of $6.40 \times 10^{-14} J$. What temperature is needed?

67. Suppose that the average velocity (v_{rms}) of carbon dioxide molecules (molecular mass is equal to 44.0 g/mol) in a flame is found to be $1.05 \times 10^5 m/s$. What temperature does this represent?

Solution

$1.95 \times 10^7 K$

68. Hydrogen molecules (molecular mass is equal to 2.016 g/mol) have an average velocity v_{rms} equal to 193 m/s. What is the temperature?

69. Much of the gas near the Sun is atomic hydrogen. Its temperature would have to be $1.5 \times 10^7 K$ for the average velocity v_{rms} to equal the escape velocity from the Sun. What is that velocity?

Solution

$6.09 \times 10^5 m/s$

70. There are two important isotopes of uranium— ^{235}U and ^{238}U ; these isotopes are nearly identical chemically but have different atomic masses. Only ^{235}U is very useful in nuclear reactors. One of the techniques for separating them (gas diffusion) is based on the different average velocities v_{rms} of uranium hexafluoride gas, UF_6 .

(a) The molecular masses for $^{235}U UF_6$ and $^{238}U UF_6$ are 349.0 g/mol and 352.0 g/mol, respectively. What is the ratio of their average velocities?

(b) At what temperature would their average velocities differ by 1.00 m/s?

(c) Do your answers in this problem imply that this technique may be difficult?

13.6: Humidity, Evaporation, and Boiling

71. Dry air is 78.1% nitrogen. What is the partial pressure of nitrogen when the atmospheric pressure is $1.01 \times 10^5 N/m^2$?

Solution

$7.89 \times 10^4 Pa$

72. (a) What is the vapor pressure of water at $20.0^\circ C$?

(b) What percentage of atmospheric pressure does this correspond to?

(c) What percent of $20.0^\circ C$ air is water vapor if it has 100% relative humidity? (The density of dry air at $20.0^\circ C$ is $1.20 kg/m^3$.)

73. Pressure cookers increase cooking speed by raising the boiling temperature of water above its value at atmospheric pressure.

(a) What pressure is necessary to raise the boiling point to $120.0^\circ C$?

(b) What gauge pressure does this correspond to?

Solution

(a) $1.99 \times 10^5 Pa$

(b) 0.97 atm

74. (a) At what temperature does water boil at an altitude of 1500 m (about 5000 ft) on a day when atmospheric pressure is $8.59 \times 10^4 \text{ N/m}^2$?

(b) What about at an altitude of 3000 m (about 10,000 ft) when atmospheric pressure is $7.00 \times 10^4 \text{ N/m}^2$?

75. What is the atmospheric pressure on top of Mt. Everest on a day when water boils there at a temperature of 70.0°C ?

Solution

$3.12 \times 10^4 \text{ Pa}$

76. At a spot in the high Andes, water boils at 80.0°C , greatly reducing the cooking speed of potatoes, for example. What is atmospheric pressure at this location?

77. What is the relative humidity on a 25.0°C day when the air contains 18.0 g/m^3 of water vapor?

Solution

78.3%

78. What is the density of water vapor in g/m^3 on a hot dry day in the desert when the temperature is 40.0°C and the relative humidity is 6.00%?

79. A deep-sea diver should breathe a gas mixture that has the same oxygen partial pressure as at sea level, where dry air contains 20.9% oxygen and has a total pressure of $1.01 \times 10^5 \text{ N/m}^2$.

(a) What is the partial pressure of oxygen at sea level?

(b) If the diver breathes a gas mixture at a pressure of $2.00 \times 10^6 \text{ N/m}^2$, what percent oxygen should it be to have the same oxygen partial pressure as at sea level?

Solution

(a) $2.12 \times 10^4 \text{ Pa}$

(b) 1.06

80. The vapor pressure of water at 40.0°C is $7.34 \times 10^3 \text{ N/m}^2$. Using the ideal gas law, calculate the density of water vapor in g/m^3 that creates a partial pressure equal to this vapor pressure. The result should be the same as the saturation vapor density at that temperature (51.1 g/m^3).

81. Air in human lungs has a temperature of 37.0°C and a saturation vapor density of 44.0 g/m^3 .

(a) If 2.00 L of air is exhaled and very dry air inhaled, what is the maximum loss of water vapor by the person?

(b) Calculate the partial pressure of water vapor having this density, and compare it with the vapor pressure of $6.31 \times 10^3 \text{ N/m}^2$.

Solution

(a) $8.80 \times 10^{-2} \text{ g}$

(b) $6.30 \times 10^3 \text{ Pa}$; the two values are nearly identical.

82. If the relative humidity is 90.0% on a muggy summer morning when the temperature is 20.0°C , what will it be later in the day when the temperature is 30.0°C , assuming the water vapor density remains constant?

83. Late on an autumn day, the relative humidity is 45.0% and the temperature is 20.0°C . What will the relative humidity be that evening when the temperature has dropped to 10.0°C , assuming constant water vapor density?

Solution

82.3%

84. Atmospheric pressure atop Mt. Everest is $.30 \times 10^4 \text{ N/m}^2$.

(a) What is the partial pressure of oxygen there if it is 20.9% of the air?

(b) What percent oxygen should a mountain climber breathe so that its partial pressure is the same as at sea level, where atmospheric pressure is $1.01 \times 10^5 \text{ N/m}^2$?

(c) One of the most severe problems for those climbing very high mountains is the extreme drying of breathing passages. Why does this drying occur?

85. What is the dew point (the temperature at which 100% relative humidity would occur) on a day when relative humidity is 39.0% at a temperature of 20.0°C ?

Solution

4.77°C

86. On a certain day, the temperature is 25.0°C and the relative humidity is 90.0%. How many grams of water must condense out of each cubic meter of air if the temperature falls to 15.0°C ? Such a drop in temperature can, thus, produce heavy dew or fog.

87. *Integrated Concepts*

The boiling point of water increases with depth because pressure increases with depth. At what depth will fresh water have a boiling point of 150°C , if the surface of the water is at sea level?

Solution

38.3m

88. *Integrated Concepts*

(a) At what depth in fresh water is the critical pressure of water reached, given that the surface is at sea level?

(b) At what temperature will this water boil?

(c) Is a significantly higher temperature needed to boil water at a greater depth?

89. *Integrated Concepts*

To get an idea of the small effect that temperature has on Archimedes' principle, calculate the fraction of a copper block's weight that is supported by the buoyant force in 0°C water and compare this fraction with the fraction supported in 95.0°C water.

Solution

$\frac{(F_B/w_{Cu})}{(F_B/w_{Cu})'} = 1.02$. The buoyant force supports nearly the exact same amount of force on the copper block in both circumstances.

90. *Integrated Concepts*

If you want to cook in water at 150°C , you need a pressure cooker that can withstand the necessary pressure.

(a) What pressure is required for the boiling point of water to be this high?

(b) If the lid of the pressure cooker is a disk 25.0 cm in diameter, what force must it be able to withstand at this pressure?

91. *Unreasonable Results*

(a) How many moles per cubic meter of an ideal gas are there at a pressure of $1.00 \times 10^{14} \text{ N/m}^2$ and at 0°C ?

(b) What is unreasonable about this result?

(c) Which premise or assumption is responsible?

Solution

(a) $4.41 \times 10^{10} \text{ mol/m}^3$

(b) It's unreasonably large.

(c) At high pressures such as these, the ideal gas law can no longer be applied. As a result, unreasonable answers come up when it is used.

92. *Unreasonable Results*

(a) An automobile mechanic claims that an aluminum rod fits loosely into its hole on an aluminum engine block because the engine is hot and the rod is cold. If the hole is 10.0% bigger in diameter than the 22.0°C rod, at what temperature will the rod be the same size as the hole?

(b) What is unreasonable about this temperature? (c) Which premise is responsible?

93. Unreasonable Results

The temperature inside a supernova explosion is said to be $2.00 \times 10^{13} K$.

- (a) What would the average velocity v_{rms} of hydrogen atoms be?
- (b) What is unreasonable about this velocity?
- (c) Which premise or assumption is responsible?

Solution

- (a) $7.03 \times 10^8 m/s$
- (b) The velocity is too high—it's greater than the speed of light
- (c) The assumption that hydrogen inside a supernova behaves as an idea gas is responsible, because of the great temperature and density in the core of a star. Furthermore, when a velocity greater than the speed of light is obtained, classical physics must be replaced by relativity, a subject not yet covered.

94. Unreasonable Results

Suppose the relative humidity is 80% on a day when the temperature is $30.0^\circ C$.

- (a) What will the relative humidity be if the air cools to $25.0^\circ C$ and the vapor density remains constant?
- (b) What is unreasonable about this result?
- (c) Which premise is responsible?

Contributors and Attributions

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