

CHAPTER OVERVIEW

12.6: The Kinetic Theory of Gases

Gases are literally all around us—the air that we breathe is a mixture of gases. Other gases include those that make breads and cakes soft, those that make drinks fizzy, and those that burn to heat many homes. Engines and refrigerators depend on the behaviors of gases, as we will see in later chapters. As we discussed in the preceding chapter, the study of heat and temperature is part of an area of physics known as thermodynamics, in which we require a system to be *macroscopic*, that is, to consist of a huge number (such as 10^{23}) of molecules. We begin by considering some macroscopic properties of gases: volume, pressure, and temperature. The simple model of a hypothetical “ideal gas” describes these properties of a gas very accurately under many conditions. We move from the ideal gas model to a more widely applicable approximation, called the Van der Waals model. To understand gases even better, we must also look at them on the *microscopic* scale of molecules. In gases, the molecules interact weakly, so the microscopic behavior of gases is relatively simple, and they serve as a good introduction to systems of many molecules. The molecular model of gases is called the kinetic theory of gases and is one of the classic examples of a molecular model that explains everyday behavior.

[Distribution of Molecular Speeds](#)

[Heat Capacity and Equipartition of Energy](#)

[Molecular Model of an Ideal Gas](#)

[Prelude to The Kinetic Theory of Gases](#)

[Pressure, Temperature, and RMS Speed](#)

[The Kinetic Theory of Gases \(Answer\)](#)

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[The Kinetic Theory of Gases Introduction \(Exercises\)](#)

Thumbnail: In an ordinary gas, so many molecules move so fast that they collide billions of times every second

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