

## 14.2: Nuclear Radioactivity

### Learning Objectives

- Explain nuclear radiation.
- Explain the types of radiation—alpha rays, beta rays, and gamma rays.
- Explain the ionization of an atom by radiation.
- Define the range of radiation.

The discovery and study of nuclear radioactivity quickly revealed evidence of revolutionary new physics. In addition, uses for nuclear radiation also emerged quickly—for example, people such as Ernest Rutherford used it to determine the size of the nucleus and devices were painted with radon-doped paint to make them glow in the dark (see [Figure 14.2.1](#)). We therefore begin our study of nuclear physics with the discovery and basic features of nuclear radioactivity.



**Figure 14.2.1:** The dials of this World War II aircraft will glow in the dark, because they are painted with radium-doped phosphorescent paint. It is a poignant reminder of the dual nature of radiation. Although radium paint dials are conveniently visible day and night, they emit radon, a hazardous radioactive gas that is odorless, colorless, and tasteless. (credit: U.S. Air Force Photo)

### Discovery of Nuclear Radioactivity

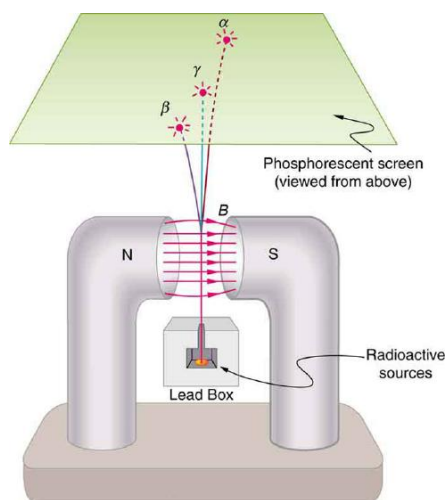
In 1896, the French physicist Antoine Henri Becquerel (1852–1908) accidentally found that a uranium-rich mineral called pitchblende emits invisible, penetrating rays that can darken a photographic plate enclosed in an opaque envelope. The rays therefore carry energy; but amazingly, the pitchblende emits them continuously without any energy input. This is an apparent violation of the law of conservation of energy, one that we now understand is due to the conversion of a small amount of mass into energy, as related in Einstein's famous equation  $E = mc^2$ . It was soon evident that Becquerel's rays originate in the nuclei of the atoms and have other unique characteristics. The emission of these rays is called **nuclear radioactivity** or simply **radioactivity**. The rays themselves are called **nuclear radiation**. A nucleus that spontaneously destroys part of its mass to emit radiation is said to **decay** (a term also used to describe the emission of radiation by atoms in excited states). A substance or object that emits nuclear radiation is said to be **radioactive**.

Two types of experimental evidence imply that Becquerel's rays originate not from the atomic structure, but something smaller inside. First, the radiation is found to be associated with certain elements, such as uranium. Radiation does not vary with chemical state—that is, uranium is radioactive whether it is in the form of an element or compound. In addition, radiation does not vary with temperature, pressure, or ionization state of the uranium atom. Since all of these factors affect electrons in an atom, the radiation cannot come from electron transitions, as atomic spectra do. The huge energy emitted during each event is the second piece of evidence that the radiation cannot be atomic in origin. Nuclear radiation has energies of the order of  $10^6$  eV per event, which is much greater than the typical atomic energies (a few eV), such as that observed in spectra and chemical reactions, and more than ten times as high as the most energetic characteristic X-rays. Becquerel did not vigorously pursue his discovery for very long. In 1898, Marie Curie (1867–1934) began her doctoral study of Becquerel's rays. She and her husband—French physicist Pierre Curie (1859–1906)—soon discovered two new radioactive elements, which she named *polonium* (after her native land) and *radium* (because it radiates). These two new elements filled holes in the periodic table and, further, displayed much higher levels of radioactivity per gram of material than uranium. Over a period of four years, working under poor conditions and spending their own funds, the Curies processed more than a ton of uranium ore to isolate a gram of radium salt. Radium became highly sought after, because it was about two million times as radioactive as uranium. Curie's radium salt glowed visibly from the radiation that took its toll on them and other unaware researchers. Shortly after completing her Ph.D., both Curies and Becquerel shared the 1903 Nobel Prize in physics for their work on radioactivity. Pierre was killed in a horse cart accident in 1906, but Marie continued her study of radioactivity for nearly 30 more years. Awarded the 1911 Nobel Prize in chemistry for her discovery of two

new elements, she remains the only person to win Nobel Prizes in physics and chemistry. Marie's radioactive fingerprints on some pages of her notebooks can still expose film, and she suffered from radiation-induced lesions. She died of leukemia likely caused by long-term exposure to radiation, but she was active in research almost until her death in 1934.

### Alpha, Beta, and Gamma

Research begun by people such as New Zealander Ernest Rutherford soon after the discovery of nuclear radiation indicated that different types of rays are emitted. Eventually, three types were distinguished and named **alpha**( $\alpha$ ), **beta**( $\beta$ ), and **gamma**( $\gamma$ ), because, like X-rays, their identities were initially unknown. Figure 14.2.2 shows what happens if the rays are passed through a magnetic field. The  $\gamma$ s are unaffected, while the  $\alpha$ s and  $\beta$ s are deflected in opposite directions, indicating that the  $\alpha$ s are positive, the  $\beta$ s negative, and the  $\gamma$ s uncharged. Rutherford used both magnetic and electric fields to show that  $\alpha$ s have a positive charge twice the magnitude of an electron, or  $+2|q_e|$ . In the process, he found the  $\alpha$ s charge to mass ratio to be several thousand times smaller than the electron's. Later on, Rutherford collected  $\alpha$ s from a radioactive source and passed an electric discharge through them, obtaining the spectrum of recently discovered helium gas. Among many important discoveries made by Rutherford and his collaborators was the proof that  *$\alpha$  radiation is the emission of a helium nucleus*. Rutherford won the Nobel Prize in chemistry in 1908 for his early work. He continued to make important contributions until his death in 1934.



**Figure 14.2.2:** Alpha, beta, and gamma rays are passed through a magnetic field on the way to a phosphorescent screen. The  $\alpha$ s and  $\beta$ s bend in opposite directions, while the  $\gamma$ s are unaffected, indicating a positive charge for  $\alpha$ s, negative for  $\beta$ s, and neutral for  $\gamma$ s. Consistent results are obtained with electric fields. Collection of the radiation offers further confirmation from the direct measurement of excess charge.

Other researchers had already proved that  $\beta$ s are negative and have the same mass and same charge-to-mass ratio as the recently discovered electron. By 1902, it was recognized that  *$\beta$  radiation is the emission of an electron*. Although  $\beta$ s are electrons, they do not exist in the nucleus before it decays and are not ejected atomic electrons—the electron is created in the nucleus at the instant of decay.

Since  $\gamma$ s remain unaffected by electric and magnetic fields, it is natural to think they might be photons. Evidence for this grew, but it was not until 1914 that this was proved by Rutherford and collaborators. By scattering  $\gamma$  radiation from a crystal and observing interference, they demonstrated that  *$\gamma$  radiation is the emission of a high-energy photon by a nucleus*. In fact,  $\gamma$  radiation comes from the de-excitation of a nucleus, just as an X-ray comes from the de-excitation of an atom. The names " $\gamma$  ray" and "X-ray" identify the source of the radiation. At the same energy,  $\gamma$  rays and X-rays are otherwise identical.

**Table 14.2.1:** Properties of Nuclear Radiation

Type of Radiation	Range
$\alpha$	A sheet of paper, a few cm of air, fractions of a mm of tissue
$\beta$	A thin aluminum plate, or tens of cm of tissue
$\gamma$	Several cm of lead or meters of concrete

## Ionization and Range

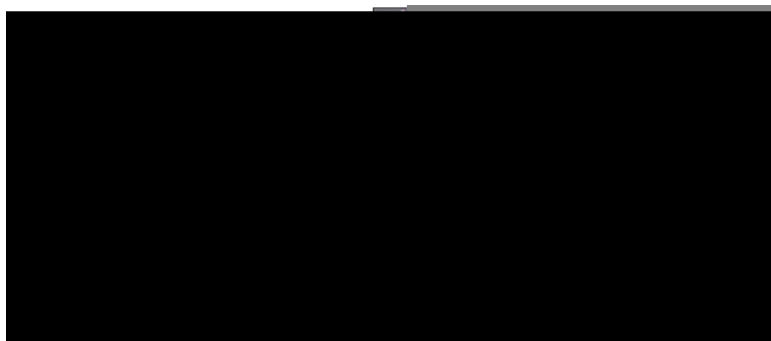
Two of the most important characteristics of  $\alpha$ ,  $\beta$ , and  $\gamma$  rays were recognized very early. All three types of nuclear radiation produce *ionization* in materials, but they penetrate different distances in materials—that is, they have different *ranges*. Let us examine why they have these characteristics and what are some of the consequences.

Like X-rays, nuclear radiation in the form of  $\alpha$ s,  $\beta$ s, and  $\gamma$ s has enough energy per event to ionize atoms and molecules in any material. The energy emitted in various nuclear decays ranges from a few keV to more than 10 MeV, while only a few eV are needed to produce ionization. The effects of X-rays and nuclear radiation on biological tissues and other materials, such as solid state electronics, are directly related to the ionization they produce. All of them, for example, can damage electronics or kill cancer cells. In addition, methods for detecting X-rays and nuclear radiation are based on ionization, directly or indirectly. All of them can ionize the air between the plates of a capacitor, for example, causing it to discharge. This is the basis of inexpensive personal radiation monitors, such as pictured in Figure 14.2.3. Apart from  $\alpha$ ,  $\beta$ , and  $\gamma$ , there are other forms of nuclear radiation as well, and these also produce ionization with similar effects. We define **ionizing radiation** as any form of radiation that produces ionization whether nuclear in origin or not, since the effects and detection of the radiation are related to ionization.



**Figure 14.2.3:** These dosimeters (literally, dose meters) are personal radiation monitors that detect the amount of radiation by the discharge of a rechargeable internal capacitor. The amount of discharge is related to the amount of ionizing radiation encountered, a measurement of dose. One dosimeter is shown in the charger. Its scale is read through an eyepiece on the top. (credit: L. Chang, Wikimedia Commons)

The **range of radiation** is defined to be the distance it can travel through a material. Range is related to several factors, including the energy of the radiation, the material encountered, and the type of radiation (see Figure 14.2.4). The higher the *energy*, the greater the range, all other factors being the same. This makes good sense, since radiation loses its energy in materials primarily by producing ionization in them, and each ionization of an atom or a molecule requires energy that is removed from the radiation. The amount of ionization is, thus, directly proportional to the energy of the particle of radiation, as is its range.



**Figure 14.2.4:** The penetration or range of radiation depends on its energy, the material it encounters, and the type of radiation. (a) Greater energy means greater range. (b) Radiation has a smaller range in materials with high electron density. (c) Alphas have the smallest range, betas have a greater range, and gammas penetrate the farthest.

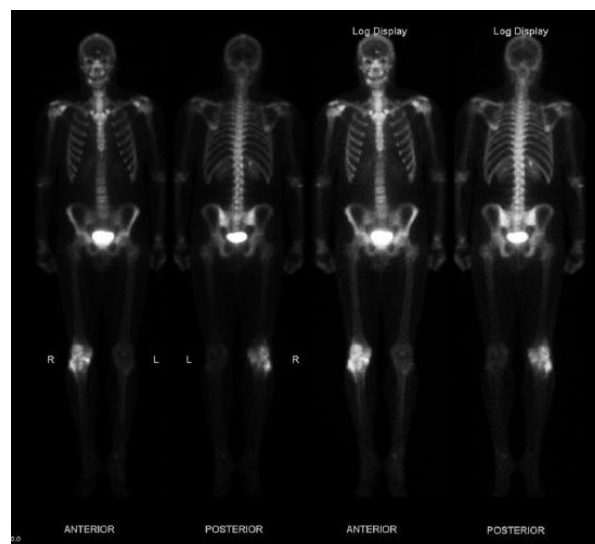
Radiation can be absorbed or shielded by materials, such as the lead aprons dentists drape on us when taking X-rays. Lead is a particularly effective shield compared with other materials, such as plastic or air. How does the range of radiation depend on *material*? Ionizing radiation interacts best with charged particles in a material. Since electrons have small masses, they most readily absorb the energy of the radiation in collisions. The greater the density of a material and, in particular, the greater the density of electrons within a material, the smaller the range of radiation.

Different *types* of radiation have different ranges when compared at the same energy and in the same material. Alphas have the shortest range, betas penetrate farther, and gammas have the greatest range. This is directly related to charge and speed of the

particle or type of radiation. At a given energy, each  $\alpha$ ,  $\beta$ , or  $\gamma$  will produce the same number of ionizations in a material (each ionization requires a certain amount of energy on average). The more readily the particle produces ionization, the more quickly it will lose its energy. The effect of *charge* is as follows: The  $\alpha$  has a charge of  $+2q_e$ , the  $\beta$  has a charge of  $-q_e$ , and the  $\gamma$  is uncharged. The electromagnetic force exerted by the  $\alpha$  is thus twice as strong as that exerted by the  $\beta$  and it is more likely to produce ionization. Although chargeless, the  $\gamma$  does interact weakly because it is an electromagnetic wave, but it is less likely to produce ionization in any encounter.

The *speed* at which they travel is the other major factor affecting the range of  $\alpha$ s,  $\beta$ s, and  $\gamma$ s. The faster they move, the less time they spend in the vicinity of an atom or a molecule, and the less likely they are to interact. Since  $\alpha$ s and  $\beta$ s are particles with mass (helium nuclei and electrons, respectively), their energy is kinetic. The mass of the  $\beta$  particle is thousands of times less than that of the  $\alpha$ s, so that  $\beta$ s must travel much faster than  $\alpha$ s to have the same energy. Since  $\beta$ s move faster, they have less time to interact than  $\alpha$ s. Gamma rays are photons, which must travel at the speed of light. They are even less likely to interact than a  $\beta$ , since they spend even less time near a given atom (and they have no charge). The range of  $\gamma$ s is thus greater than the range of  $\beta$ s.

Alpha radiation from radioactive sources has a range much less than a millimeter of biological tissues, usually not enough to even penetrate the dead layers of our skin. On the other hand, the same  $\alpha$  radiation can penetrate a few centimeters of air, so mere distance from a source prevents  $\alpha$  radiation from reaching us. This makes  $\alpha$  radiation relatively safe for our body compared to  $\beta$  and  $\gamma$  radiation. Typical  $\beta$  radiation can penetrate a few millimeters of tissue or about a meter of air. Beta radiation is thus hazardous even when not ingested. The range of  $\beta$ s in lead is about a millimeter, and so it is easy to store  $\beta$  sources in lead radiation-proof containers. Gamma rays have a much greater range than either  $\alpha$ s or  $\beta$ s. In fact, if a given thickness of material, like a lead brick, absorbs 90% of the  $\gamma$ s, then a second lead brick will only absorb 90% of what got through the first. Thus,  $\gamma$ s do not have a well-defined range; we can only cut down the amount that gets through. Typically,  $\gamma$ s can penetrate many meters of air, go right through our bodies, and are effectively shielded (that is, reduced in intensity to acceptable levels) by many centimeters of lead. One benefit of  $\gamma$ s is that they can be used as radioactive tracers (see [Figure 14.2.5](#)).



**Figure 14.2.5:** This image of the concentration of a radioactive tracer in a patient's body reveals where the most active bone cells are, an indication of bone cancer. A short-lived radioactive substance that locates itself selectively is given to the patient, and the radiation is measured with an external detector. The emitted  $\gamma$  radiation has a sufficient range to leave the body—the range of  $\alpha$ s and  $\beta$ s is too small for them to be observed outside the patient. (credit: Kieran Maher, Wikimedia Commons)

### Section Summary

- Some nuclei are radioactive—they spontaneously decay destroying some part of their mass and emitting energetic rays, a process called nuclear radioactivity.
- Nuclear radiation, like X-rays, is ionizing radiation, because energy sufficient to ionize matter is emitted in each decay.
- The range (or distance traveled in a material) of ionizing radiation is affected by the charge of the emitted particle and its energy. Greater electric charge and lower energy lead to shorter ranges; less electric charge and higher energy lead to longer ranges.
- Radiation detectors are based directly or indirectly upon the ionization created by radiation, as are the effects of radiation on living and inert materials.

## Glossary

**alpha ray**

a component of nuclear radioactivity; consists of two protons and two neutrons

**beta ray**

a component of nuclear radioactivity; consists of an electron

**gamma ray**

a component of nuclear radioactivity; consists of a high-energy photon

**ionizing radiation**

a radiation that produces ionization; examples include alpha rays, beta rays, gamma rays, and X-rays

**nuclear radiation**

radiation that originate in the nuclei of atoms; alpha, beta, and gamma rays are among most common forms; also see "radioactivity"

**radioactivity**

the emission of rays from the nuclei of atoms

**radioactive**

a substance or object that emits nuclear radiation

**range of radiation**

the distance that the radiation can travel through a material

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